Accurate quantum calculations of the reaction rates for H/D+CH₄

Rob van Harrevelt, a) Gunnar Nyman, b) and Uwe Manthe

Theoretische Chemie, Universität Bielefeld, Universitätsstrasse 25, D-33615 Bielefeld, Germany

(Received 11 December 2006; accepted 9 January 2007; published online 22 February 2007)

In previous work [T. Wu, H. J. Werner, and U. Manthe, Science 306, 2227 (2004)], accurate quantum reaction rate calculations of the rate constant for the H+CH₄→CH₃+H₂ reaction have been presented. Both the electronic structure calculations and the nuclear dynamics calculations are converged with respect to the basis sets employed. In this paper, the authors apply the same methodology to an isotopic variant of this reaction: D+CH₄→CH₃+HD. Accurate rate constants are presented for temperatures between 250 and 400 K. For temperatures between 400 and 800 K, they use a harmonic extrapolation to obtain approximate rate constants for H/D+CH₄. The calculations suggest that the experimentally reported rate constants for D+CH₄ are about a factor of 10–20 too high. For H+CH₄, more accurate experiments are available and agreement is much better: the difference is less than a factor of 2.6. The kinetic isotope effect for the H/D+CH₄ reactions is studied and compared with experiment and transition state theory (TST) calculations. Harmonic TST was found to provide a good description of the kinetic isotope effect. © 2007 American Institute of Physics. [DOI: 10.1063/1.2464102]

I. INTRODUCTION

The reaction H+CH₄→CH₃+H₂ is of fundamental interest as an elementary hydrocarbon-radical reaction. The reaction is also relevant in combustion and atmospheric chemistry. It has been the subject of many theoretical studies: reduced dimensionality quantum calculations,1–11 and quantum transition state theory based approaches12–18 have been published by various groups. Even accurate full-dimensional quantum dynamics calculations for thermal rate constants and cumulative reaction probabilities have been presented.19–21 In a recent study,22,23 high-level electronic structure calculations, sophisticated interpolation schemes for the potential energy surface, and full-dimensional quantum dynamics simulations have been combined to obtain an accurate rate constant from first principles. All sources of errors have been carefully studied: the convergence with the basis set in the electronic structure and quantum dynamics calculations, and the accuracy of the interpolation scheme. Only inaccuracies resulting from the coupled cluster [CCSD(T)] approach and the J-shifting approximation could not be estimated quantitatively. If the CCSD(T) method and the J-shifting approximations are reasonable, then the accuracy of the ab initio approach can compete with, or is even better than, available experiments. The estimated numerical uncertainty in the computed rate constant is about 30%.

In previous work22,23 the absolute value of the thermal rate constant was studied. Here, we will also examine the kinetic isotope effect (KIE) in the X+CH₄→CH₃+HX (X=H,D) reaction. Kurylo et al.25 found experimentally that D reacts faster with methane than H. These results are quantitatively reproduced by theoretical studies of Pu and Truhlar,15 who used quantum transition state theory approaches and a semiempirical potential energy surface. Since it is now possible to calculate accurate rate constants from first principles, we can predict kinetic isotope effects with high accuracy. These results are presumably more accurate than the experimental ones.

The main goal of the present study is to calculate accurate rate constants for D+CH₄ in the temperature range of 250–400 K. In combination with previous results for H+CH₄,22,23 we can then obtain the kinetic isotope effect. A second goal of this paper is to extend rate constant calculations for H/D+CH₄ to higher temperatures. Accurate, fully converged quantum calculations are only possible for temperatures up to about 400–500 K. At higher temperatures, the number of vibrationally excited states of the activated complex that participates in the reaction is so large that some approximation is required. In this paper, we use a harmonic progression scheme to estimate the rate constant for temperatures up to 800 K. The accuracy of this scheme is investigated in detail.

The article is organized as follows. Section II discusses the theoretical approach followed in this work. It is divided into two parts. Section II A discusses accurate reaction rate calculations. We show how a harmonic extrapolation can be used to enhance convergence of the rate constant.26 Section II B presents a harmonic progression based approach to obtain approximate rate constants. Section III gives the numerical details and Sec. IV discusses the results. Section IV A presents accurate results for the temperature range of 250–500 K and Sec. IV B presents approximate rate constants for temperatures up to 800 K. The paper concludes with a summary and some final remarks.
II. THEORY

A. Accurate reaction rate calculations

The present work employs the same theoretical and computational scheme as described in detail in Ref. 22. In this scheme, the flux-correlation formalism 27–29 is used to calculate the rate constant directly, without solving the full scattering problem. A short summary of the followed approach is given below; for further details, we refer to Ref. 22. Assuming that the J-shifting 24 approximation is valid, the rate constant \( k \) is calculated from

\[
k = \frac{Q^j_{\text{rot}}}{2\pi\hbar Q_r} \int dE \exp(-E/k_BT),
\]

where \( Q^j_{\text{rot}} \) is the rotational partition function at the transition state geometry, \( Q_r \) the reactant partition function, \( E \) the total energy, and \( N(E) \) the cumulative reaction probability for total angular momentum \( J=0 \). \( N(E) \) is the quantity which is obtained from accurate quantum dynamics calculations.

The cumulative reaction probability obtained from the quantum simulations can be given as a sum of several terms \( N_j(E): N(E) = \sum N_j(E) \). The \( N_j(E) \)'s correspond to different pairs of eigenstates of the thermal flux operator. Each contribution \( N_j(E) \) can be considered as the contribution of the \( j \)th vibrational state of the activated complex to \( N(E) \). For high temperatures, the number of states required to converge the thermal rate constant increases dramatically. To enhance the convergence with respect to the number of contributions explicitly considered, a harmonic extrapolation can be employed. Assume that \( N_j(E) \) is calculated for \( i = 0, n \). The best estimate of the rate constant based on the data for the \( n+1 \) lowest vibrational states of the activated complex reads as follows: 26

\[
k_n(T) = \frac{Q^j_{\text{rot}}}{2\pi\hbar Q_r} \sum_{i=0}^{n} \exp(-E_i/k_BT) \\
\times \int dE \left\{ \sum_{i=0}^{n} N_j(E) \right\} \exp(-E/k_BT),
\]

where \( E_i \) is the energy of the \( i \)th vibrational state of the activated complex in harmonic approximation (\( i=0 \) corresponds to the ground state of the activated complex). The harmonic frequencies can be obtained from normal mode analysis at the transition state geometry. It is easy to see that \( k_n(T) \) converges to the exact \( k(T) \):

\[
\lim_{n \to \infty} k_n(T) = k(T).
\]

B. Approximate rate constants based on the harmonic progression

When only the ground state contribution \( N_0(E) \) is explicitly included in the rate constant calculation, then Eq. (2) reduces to

\[
k_0(T) = \frac{Q^j_{\text{rot}}}{2\pi\hbar Q_r} \sum_{i=0}^{\infty} \exp(-E_i/k_BT) \\
\times \int dE \exp(-E/k_BT).
\]

Since the term \( \sum_{i=0}^{\infty} \exp(-E_i/k_BT) \) is the harmonic approximation of the vibrational partition function of the transition state, \( Q^j_{\text{vib}} \), we can write

\[
k_0(T) = \frac{Q^j_{\text{rot}}Q^j_{\text{vib}}}{2\pi\hbar Q_r} \int dE \exp(-E/k_BT).
\]

Equation (5) is identical to Eq. (1) when the \( N_j(E) \)'s obey

\[
N_j(E) = N_0(E - (E_i - E_0)).
\]

This expression is the vibrational analog to the J-shifting approximation 24 for rotational motion. Here, the vibration orthogonal to the reaction path is treated similarly as the overall rotation of the complex in the J-shifting approximation. In this paper, we will show that Eq. (5) can be a very useful approximation to the rate constant.

The approximate expression for the rate constant, Eq. (5), also makes the relation with harmonic transition state theory (TST) more transparent. Harmonic TST assumes that \( N_0(E) = \Theta(E - E_b) \), where \( \Theta \) is the Heaviside step function and \( E_b \) is the vibrationally adiabatic barrier height in the harmonic approximation. Then, since \( \int dE \exp(-E/k_BT) = k_BT \exp(-E_b/k_BT) \),

\[
k_{\text{TST}}(T) = k_BT \frac{Q^j_{\text{rot}}Q^j_{\text{vib}}}{2\pi\hbar Q_r} \exp(-E_b/k_BT).
\]

It should be noted that any theory based on the harmonic approximation for the transition state should also use the harmonic approximation to describe the vibrational excitation energies in the reactant partition function \( Q_r \).

III. NUMERICAL DETAILS

The 12-dimensional wave packet is propagated using the multiconfiguration time dependent Hartree (MCTDH) approach. 32–35 The coordinates used in the MCTDH wave packet propagation are obtained from a linear transformation of transition state normal mode coordinates. The harmonic frequencies for the transition state normal modes on the present potential energy surface are presented in Table I. The reason why a transformation of normal modes is advantageous in the MCTDH calculations is discussed in Refs. 19 and 20. Mode 1 (the mode with the imaginary frequency) is mixed with one other mode:

\[
\begin{pmatrix}
Q_1' \\
Q_2'
\end{pmatrix} = \begin{pmatrix}
\cos \gamma & -\sin \gamma \\
\sin \gamma & \cos \gamma
\end{pmatrix}
\begin{pmatrix}
Q_1 \\
Q_2
\end{pmatrix}.
\]

The two modes \( Q_1 \) and \( Q_2 \) largely correspond to the motion of the H/D atoms located on the \( C_3 \) symmetry axis. Mode \( k \) is 9 and 7 for \( \text{H}+\text{CH}_4 \) and \( \text{D}+\text{CH}_4 \), respectively. The other transition state normal mode coordinates are not transformed. In this work, we employ \( \gamma=20° \) for both systems. The dividing surface can now be defined as \( Q_1'=0 \), and the
transformed coordinates \( Q_i' \) are employed in the dynamical calculations.

The parameters used to represent the wave function, the numbers of single particle functions and the grid sizes, are based on extensive convergence tests. The procedure has been described in detail in Refs. 20–22. Table II gives the parameters for the converged D+CH4 calculations. Parameters for H+CH4 have the same values as in previous work.22,23

An important parameter in the quantum dynamics simulations is the reference temperature \( T_{\text{ref}} \).36,37 The results are formally independent of \( T_{\text{ref}} \). However, this parameter determines the energy range where the computed \( N(E) \) is reliable. This range corresponds to those energies where \( N(E) \exp(-E/k_B T_{\text{ref}}) \) makes a significant contribution to the integral in Eq. (1). To obtain the results presented below, different reference temperatures have been used to study different temperature ranges.

IV. RESULTS AND DISCUSSION

A. Accurate rate constants for D+CH4 and the kinetic isotope effect

Using a reference temperature of \( T_{\text{ref}}=300 \) K, we obtained accurate rate constants for the temperature range of 250–400 K. Rate constants are presented in Table III and plotted in Fig. 1. Figure 1 also shows the harmonic TST result and the experimental data of Kurylo et al.25 The difference between TST and accurate quantum results clearly demonstrates the role of tunneling. The experimental results are higher than the present result, probably due to side reactions such as D+CH3→CH2D+H in the experiment.25

In the present calculation, six states of the activated complex, i.e., the ground state and the first five vibrationally excited states, have been considered explicitly. We used the harmonic extrapolation, Eq. (2) to estimate the contributions of the higher excited states. The estimated contributions of higher excited states are rather small: in total 2% at 300 K and 10% at 400 K. By studying the dependence of the rate constant on the number of vibrationally excited states (\( n \)) included explicitly, we can gain insight into the accuracy of the harmonic extrapolation scheme. The differences between the rate constants obtained with \( n=0, 1, 2, 3, 4, \) and 5 are negligible when the harmonic extrapolation is used. As an example, Fig. 1 shows the results for both \( n=0 \) and \( n=5 \).

This suggests that the harmonic extrapolation yields a reliable estimate of the effect of the excited states not included explicitly. In this case, it would have been sufficient to calculate \( N_0(E) \), i.e., the contribution of the ground state of the activated complex. The contributions of all other states could be evaluated using the harmonic extrapolation [Eq. (5)]. Table IV gives the values of the rate constant obtained using

---

**TABLE I. Harmonic frequencies at the transition state geometry (in cm\(^{-1}\))**, including the imaginary frequency (mode \( Q_2 \)). Note that some modes are doubly degenerate as they correspond to the \( E \) irreducible representation of the \( C_3v \) point group.

<table>
<thead>
<tr>
<th>Modes</th>
<th>Frequency</th>
</tr>
</thead>
<tbody>
<tr>
<td>( Q_1 )</td>
<td>1441a</td>
</tr>
<tr>
<td>( Q_2, Q_3 )</td>
<td>534</td>
</tr>
<tr>
<td>( Q_4 )</td>
<td>1073</td>
</tr>
<tr>
<td>( Q_5, Q_6 )</td>
<td>1124</td>
</tr>
<tr>
<td>( Q_7, Q_8 )</td>
<td>1442</td>
</tr>
<tr>
<td>( Q_9 )</td>
<td>1795</td>
</tr>
<tr>
<td>( Q_{10} )</td>
<td>3076</td>
</tr>
<tr>
<td>( Q_{11}, Q_{12} )</td>
<td>3252</td>
</tr>
</tbody>
</table>

**TABLE II. Parameters for the MCTDH representation of the wave function for D+CH4.**

<table>
<thead>
<tr>
<th>Coordinate</th>
<th>Number of single-particle functions</th>
<th>Grid size</th>
<th>Grid type</th>
<th>Grid range for FFT (a.u.)</th>
</tr>
</thead>
<tbody>
<tr>
<td>( Q_i' )</td>
<td>6</td>
<td>48</td>
<td>FFT</td>
<td>–92–96</td>
</tr>
<tr>
<td>( Q_2/Q_3 )</td>
<td>4</td>
<td>32</td>
<td>Hermite DVR</td>
<td></td>
</tr>
<tr>
<td>( Q_4 )</td>
<td>3</td>
<td>48</td>
<td>FFT</td>
<td>–60–132</td>
</tr>
<tr>
<td>( Q_5/Q_6 )</td>
<td>3</td>
<td>18</td>
<td>Hermite DVR</td>
<td></td>
</tr>
<tr>
<td>( Q_7' )</td>
<td>5</td>
<td>48</td>
<td>FFT</td>
<td>–115–80</td>
</tr>
<tr>
<td>( Q_8/Q_9 )</td>
<td>2</td>
<td>8</td>
<td>Hermite DVR</td>
<td></td>
</tr>
<tr>
<td>( Q_{10} )</td>
<td>2</td>
<td>8</td>
<td>Hermite DVR</td>
<td></td>
</tr>
<tr>
<td>( Q_{11}/Q_{12} )</td>
<td>2</td>
<td>8</td>
<td>Hermite DVR</td>
<td></td>
</tr>
</tbody>
</table>

**TABLE III. Rate constants (in cm\(^{-3}\)/s) for H+CH4→CH3+H\(_2\)/D+CH4→CH3+HD: results of accurate quantum calculations (Refs. 22, 23, and 38) (H+CH\(_4\)) and the present work (D+CH\(_4\)).**

<table>
<thead>
<tr>
<th>( T (K) )</th>
<th>( H+CH_4 )</th>
<th>( D+CH_4 )</th>
</tr>
</thead>
<tbody>
<tr>
<td>250</td>
<td>( 3.6 \times 10^{-21} )</td>
<td>( 1.2 \times 10^{-20} )</td>
</tr>
<tr>
<td>300</td>
<td>( 7.8 \times 10^{-20} )</td>
<td>( 2.2 \times 10^{-19} )</td>
</tr>
<tr>
<td>350</td>
<td>( 9.5 \times 10^{-19} )</td>
<td>( 2.3 \times 10^{-18} )</td>
</tr>
<tr>
<td>400</td>
<td>( 7.3 \times 10^{-18} )</td>
<td>( 1.6 \times 10^{-17} )</td>
</tr>
</tbody>
</table>

**TABLE IV. Harmonic extrapolation, Eq. (2) to estimate the contributions of the higher excited states.**

<table>
<thead>
<tr>
<th>Number of vibrationally excited states</th>
<th>Rate constant ( a.u. )</th>
</tr>
</thead>
<tbody>
<tr>
<td>( n=0 )</td>
<td></td>
</tr>
<tr>
<td>( n=1 )</td>
<td></td>
</tr>
<tr>
<td>( n=2 )</td>
<td></td>
</tr>
<tr>
<td>( n=3 )</td>
<td></td>
</tr>
<tr>
<td>( n=4 )</td>
<td></td>
</tr>
<tr>
<td>( n=5 )</td>
<td></td>
</tr>
</tbody>
</table>

---

**FIG. 1. Rate constants for the D+CH\(_4\)→CH\(_3\)+HD reaction. \( n \) is the number of vibrationally excited states of the activated complex included explicitly in the accurate quantum calculations [see Eq. (2)]. Experimental rate constants are taken from Ref. 25.**
TABLE IV. Approximate rate constants (in cm³/s) for H+CH₄→CH₃+H₂ and D+CH₄→CH₃+HD calculated using Eq. (5). Two reference temperatures (Tref) have been used to cover different temperature ranges. The H+CH₄, Tref=300 K results are based on calculations presented in Refs. 22 and 23.

<table>
<thead>
<tr>
<th></th>
<th>H+CH₄</th>
<th>D+CH₄</th>
</tr>
</thead>
<tbody>
<tr>
<td></td>
<td>T (K)</td>
<td>Tref=300 K</td>
</tr>
<tr>
<td>---</td>
<td>-------</td>
<td>------------</td>
</tr>
<tr>
<td>250</td>
<td></td>
<td>3.4×10⁻²¹</td>
</tr>
<tr>
<td>300</td>
<td></td>
<td>7.8×10⁻²⁰</td>
</tr>
<tr>
<td>350</td>
<td></td>
<td>9.8×10⁻¹⁹</td>
</tr>
<tr>
<td>400</td>
<td></td>
<td>7.6×10⁻¹⁸</td>
</tr>
<tr>
<td>450</td>
<td></td>
<td></td>
</tr>
<tr>
<td>500</td>
<td></td>
<td>1.8×10⁻¹⁶</td>
</tr>
<tr>
<td>600</td>
<td></td>
<td>1.6×10⁻¹⁵</td>
</tr>
<tr>
<td>700</td>
<td></td>
<td>8.3×10⁻¹⁵</td>
</tr>
<tr>
<td>800</td>
<td></td>
<td>3.0×10⁻¹⁴</td>
</tr>
</tbody>
</table>

The success of the harmonic extrapolation scheme can also be seen in the cumulative reaction probabilities (see Fig. 2). For a detailed investigation of the energy shifting scheme, Eq. (6), Fig. 2(b) shows Nₐ(E) and the shifted contributions Nᵢ(E+(Eᵢ−E₀)) and N₂(E+(E₂−E₀)), where Eᵢ−E₀ is the energy difference in harmonic approximation. To a good approximation, the Nᵢ(E) and N₂(E) obey Nᵢ(E+(Eᵢ−E₀)) = Nᵢ(E), which is equivalent to Eq. (6).

Using the rate constant for the H+CH₄→CH₃+H₂ reaction obtained in previous work 22,23 the KIE for H/D+CH₄ can be calculated. The individual rate constants for both reactions are given in Table II. The KIE (the ratio of the H+CH₄ to the D+CH₄ rate) is presented in Fig. 3. The rate constants for H+CH₄ and D+CH₄ are obtained using the same potential energy surface and computational approach. Errors due to inaccuracies of the potential energy surface, in particular, the barrier height, should largely cancel in the calculated KIE. However, the dynamical calculations are converged with respect to basis set parameters independently. Consequently, inaccuracies of the dynamical calculations, which are expected to be in the 10%–20% range, might add up in the computed KIE. The KIE is therefore expected to be accurate to within about 20%. The numerical error could be temperature dependent. Considering the numerical uncertainty, the difference between the TST and quantum results is not significant.

Thus, the kinetic isotope effect is adequately described by harmonic TST, even in the temperature range where tunneling is important. This was also found in the quantum transition state theory study of Pu and Truhlar.15 Apparently, the tunneling enhancement k/kTST is almost the same for H+CH₄ and D+CH₄. This should not be surprising. In both reactions, a hydrogen atom is transferred. The imaginary frequency modes are almost identical for H+CH₄ and D+CH₄ (see Table I). The main reason for the higher rate constant for D+CH₄ is the lower zero-point vibrational energy at the transition state, which is 9296 and 8992 cm⁻¹ for H+CH₄ and D+CH₄, respectively.

Figure 3 includes the KIE obtained from the analysis of the experimental data of Refs. 25 and 39. The analysis is based on fits of the experimental rate constants for H+CH₄ and D+CH₄ (see Table I). The main reason for the higher rate constant for D+CH₄ is the lower zero-point vibrational energy at the transition state, which is 9296 and 8992 cm⁻¹ for H+CH₄ and D+CH₄, respectively.

FIG. 2. Cumulative reaction probabilities for D+CH₄ as functions of the total energy. Panel (a): N(E) and several contributions Nᵢ(E). Panel (b): The ground state contribution N₀(E) and the shifted contributions N₁(E+(E₁−E₀)) and N₂(E+(E₂−E₀)).

FIG. 3. The kinetic isotope effect (kB/KB₄CH₄). Solid line: the KIE based on accurate quantum results for H+CH₄ (Refs. 22 and 23) and D+CH₄ (this work). Dashed line: harmonic TST result. Dashed-dotted line: experiment (Ref. 25).
mental KIE is probably reasonably accurate. A similar temperature dependence of the experimental and TST results can clearly be seen in Fig. 3.

**B. H/D+CH\textsubscript{4} rate constants at high temperatures**

To calculate the rate constants at high temperatures, 350 ≤ T ≤ 800 K, we use a higher reference temperature: T\textsubscript{ref} = 500 K. We further employ a harmonic progression to calculate the rate constant from the ground state contribution \(N_0(E)\) to the cumulative reaction probability. In the previous section, we have shown that the harmonic extrapolation based on the ground state contribution \(N_0(E)\) gives an accurate rate constant for 250 ≤ T ≤ 400 K. One can reasonably expect that this approach yields a reliable estimate also at higher temperatures. The results are presented in Table IV and Fig. 4. The results obtained with the two reference temperatures differ by less than 12%, in the overlapping temperature range (350–400 K), which is consistent with the estimated numerical uncertainty, 20%, in the calculations.

Figure 4 also presents experimental rate constants. Sutherland \textit{et al.}\textsuperscript{40} have recently reviewed different rate constant measurements for H+CH\textsubscript{4}. Based on a reanalysis of old results and new high-temperature results, they obtain an analytical expression for the rate constant that fits the available experimental results for temperatures between 350 and 1950 K. Their result is somewhat higher than the present theoretical result. However, with increasing temperature, the relative difference between theory and experiment diminishes. The ratio between the experimental and the calculated rate constants varies from 2.6 at 350 K to 1.3 at 800 K. The estimated uncertainty in the calculated result is about 30%. In contrast, experimental results at higher temperatures are likely more accurate than those at lower temperatures, where the rate constant is very low. It therefore seems reasonable to assume that at low temperatures the present theoretical results are more reliable than the available experimental data. More accurate experiments at low temperatures are required to be able to draw final conclusions.

For D+CH\textsubscript{4}, the experimental rate constant is about a factor of 10–20 higher than the theoretical results, even at high temperatures. The experiments for D+CH\textsubscript{4} are more difficult than for H+CH\textsubscript{4} due to side reactions such as D+CH\textsubscript{3} → CH\textsubscript{2}D,\textsuperscript{25} which might explain the large difference. Thus, additional experimental work is clearly desirable.

**V. CONCLUSION**

In this paper, we present calculations of the rate constant for the D+CH\textsubscript{4} → CH\textsubscript{3}+HD reaction. For temperatures below 400 K, we have shown that a harmonic progression of the cumulative reaction probabilities yields a reliable estimate of the rate constant. We have therefore also used this method to calculate the rate constants for temperatures up to 800 K. Using harmonic extrapolation, we have also extended previous calculations for H+CH\textsubscript{4} → CH\textsubscript{3}+H\textsubscript{2} (Refs. 22 and 23) to higher temperatures.

The same theoretical and computational scheme as used in Refs. 22 and 23 has been employed to calculate the cumulative reaction probabilities. It combines accurate quantum dynamics simulations with high-quality electronic structure calculations. As discussed in Ref. 22, we expect that the theoretical rate constants are accurate within about 30%. We therefore attribute the deviation of the experimental and theoretical rate constants to inaccuracies in the experiment. We hope that the present work and the work in Refs. 22 and 23 will challenge experimentalists to measure accurate rate constants for H/D+CH\textsubscript{4}, in particular, at low temperatures (around 300 K), where tunneling significantly enhances the rate constant.

This paper has also examined the secondary isotope effect in the H+CH\textsubscript{4} reaction. As expected, the isotope effect...
is adequately described with transition state theory, even at low temperatures were tunneling is important. Since the transferred atom is a H atom in both H+CH₄ and D+CH₄ reactions, the tunneling enhancement is almost the same.

38Minor errors found in the data analysis have been corrected.