Synthesis and Supramolecular Chemistry of Novel Liquid Crystalline Crown Ether-Substituted Phthalocyanines: Toward Molecular Wires and Molecular Ionoelectronics

Cornelus F. van Nostrum,† Stephen J. Picken,† Arend-Jan Schouten,‡ and Roeland J. M. Nolte*†,

Contribution from the Department of Organic Chemistry, N.S.R. Center, University of Nijmegen, Toernooiveld, 6525 ED Nijmegen, The Netherlands, Akzo Nobel Central Research, Physical Chemistry Department, P.O. Box 9300, 6800 SB Arnhem, The Netherlands, and Laboratory of Polymer Science, University of Groningen, Nijenborgh 4, 9747 AG Groningen, The Netherlands

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Abstract: The synthesis of the metal-free and the dihydroxysilicon derivatives of tetrakis[4’,5’-bis(decoxy)benzo-18-crown-6]phthalocyanine is described. The metal-free phthalocyanine is liquid crystalline and exhibits a crystalline phase to mesophase transition at 148 °C. The structures of the crystalline phase and the mesophase are determined by X-ray measurements. The metal-free compound strongly aggregates in chloroform solution to form a gel. Electron micrographs show that this gel contains a network of fibers, each of which is built up of parallel strands of supramolecules having the thickness of one molecule and a length of several micrometers. The strands are formed by a process of self-assembly involving up to 104 molecules. They can be considered as being molecular cables, containing a central electron wire, four ion channels, and a surrounding insulating hydrocarbon mantle. The silicon derivative contains two axial hydroxy groups which prevent the molecule from aggregating. This compound is not liquid crystalline. It forms a stable monolayer at the air-water interface. In this layer, the phthalocyanine planes are oriented parallel to the water surface. The monolayers can be transferred onto glass substrates by a Y-type deposition. The resulting Langmuir—Blodgett film is built up of bilayers containing slipped face-to-face phthalocyanine dimers. The monolayer is capable of binding alkali metal ions from the subphase, as is concluded from surface area—surface pressure isotherms. The binding constant for potassium ions has been determined by analyzing the isotherms as a function of the concentration of this metal ion. The dihydroxysilicon phthalocyanine can be polymerized to form a polysiloxane.

Introduction

The study of nanometer-sized structures and molecular materials, which are formed by self-organization, is of great interest for the construction of devices that are small and fast and that can store information with high density.1 Molecular electronics and molecular ionics are new scientific fields dealing with this subject.2,3 In the latter case, the information storage and processing are achieved with the help of ions. An important step in this process is the nonlinear complexation of ions, which allows for reliable switching between on and off states.3d Furthermore, the construction of supramolecular wires and channels capable of transporting electrons and ions is a great challenge.4

Phthalocyanines (Pcs) belong to a class of compounds that has potential for application in functional nanostructures. Phthalocyanines substituted with long, flexible hydrocarbon chains can self-assemble to form columnar mesophases,5 Langmuir—Blodgett (LB) multilayers,6,7 or aggregates in solution.8

Recently, we showed that crown ether-substituted Pcs can be stacked to create ion conducting channels.8 The electron conducting properties of these structures have been extensively studied.5,6,8b,9

In this paper we describe the synthesis and self-organizing properties of the metal-free and silicon Pcs 1a,b.10 These compounds contain a central Pc core, four 18-crown-6 rings, and a total of eight decoxy chains linked via benzene units to the crown ether rings. They were designed to develop systems...
that combine the properties of electron conduction, ion conduction, and liquid crystallinity. CPK molecular models show that molecules la,b display a high degree of shape anisotropy, which is favorable for the formation of a (columnar) mesophase. Taking as the thickness of the molecules a value of 4.2 Å, the length over thickness (L/d) aspect ratio is in the range of 9.5–15, depending on the conformation of the peripheral decyano chains (see Figure 1). We will show in this paper that la indeed forms a thermotropic mesophase. Moreover, this compound self-assembles in chloroform solution to yield an extremely long supramolecular cable. Phthalocyanine lb forms stable monoalkali metal ions from the subphase and can be transferred onto solid substrates, giving a LB film. Preliminary results indicate that lb can be polymerized to form an axial polysiloxane.

![Figure 1. Schematic representation of the disk-like molecules showing the two limiting diameters](image)

### Experimental Section

**Materials and Methods.** All solvents were dried before use. Merck Kieselgel 60H was used for chromatography. 3,4-Dihydroxybenzaldehyde was recrystallized twice from o-xylene. K$_3$CO$_3$ was dried in an oven before use. Potassium nitrosodisulfonate (Fremy’s radical) was prepared according to the method of Zimmer et al. Catechol was recrystallized from toluene. The tetrahydroxypropyl ether of 2-(2’-chloretoxy)ethanol was synthesized according to the procedure of Kyba et al. $p$-Toluenesulfonic acid chloride was recrystallized from hexane. (N,N-Dimethylamino)ethanol and SiCl$_4$ were freshly distilled before use. Quinoline was freshly distilled from barium oxide. All other reagents were used as supplied, without further purification.

Melting points were determined with a Reichert hot-stage microscope and are uncorrected. Infrared (IR) spectra were recorded with a Perkin Elmer 28 infrared spectrophotometer or a Perkin Elmer 1720-X infrared Fourier transform spectrophotometer. UV/vis spectra were obtained with a Perkin Elmer Lambda 5 spectrophotometer. $^1$H NMR spectra were obtained with Varian EM 390, Bruker AC-100, or Bruker WM-400 instruments. Chemical shifts are reported in ppm downfield from internal TMS standard. Abbreviations used are as follows: s, singlet; d, doublet; t, triplet; m, multiplet; br, broad. Electronic ionization (EI), chemical ionization (CI), and fast atom bombardment (FAB) mass spectra (MS) were obtained with a VG-7070E apparatus. Elemental analysis were carried out on an EA 1108 Carlo Erba instrument.

Differential scanning calorimetry (DSC) data were obtained with a Mettler DSC 12E instrument. The measurements were carried out under an inert atmosphere with heating and cooling rates of 10 °C/min. Thermogravimetric analysis (TGA) was carried out on a Perkin Elmer System 4 under an inert atmosphere. For polarization microscopy, a Leitz Orthoplan polarization microscope was used, equipped with a Mettler FP80/FP82 hot stage. Small-angle X-ray scattering (SAXS) curves were recorded using a Siemens D5000 reflection diffractometer with a HTK oven and a Cu K$_\alpha$ source with a wavelength of 1.5406 Å.

Transmission electron microscopy was carried out with a Philips TEM 201 instrument. Small drops of the samples were placed on carbon-coated grids, the material was allowed to adsorb for 1 min, and then the grids were blotted dry by touching the edges with filter paper. Finally, the structures were visualized by shadowing with platinum at an angle of ~45°.

Monolayers at the air–water interface were studied by measuring pressure–area isotherms on a computer-controlled Lauda film balance (FW 2) and on a home-built film balance, in both cases using water that was purified by a Milli-Q filtration system as the subphase. A sample was dissolved in chloroform (spectroscopic quality), spread on the water surface, and isotherms were recorded at a compression speed of ~30 Å/molecule·min at 20 °C. The water surface was visualized during compression with a Brewster angle microscope (NEF BAM-1) equipped with a 10 mW He–Ne laser with a beam diameter of 0.68 mm operating at 632.8 nm. Reflections were detected using a CCD camera. Glass substrates were cleaned ultrasonically with chloroform, treated with concentrated sulfochromic acid at 80 °C for 2 h, washed several times with Milli-Q water, cleaned again ultrasonically in acetone and in chloroform, and finally stored in methanol. Before use, the substrate was rinsed with chloroform, partially hydrophobized by treatment with a boiling mixture of chloroform and hexamethyldisilazane, and finally rinsed with chloroform. Deposition was carried out by vertically dipping the substrate from the solution for 10 min.

**Syntheses.** 3,4-Bis(6-deoxybenzaldehyde) (3), 3,4-Dihydroxybenzaldehyde (10.6 g, 0.077 mol) and 1-bromodecane (35.6 g, 0.161 mol) were dissolved in DMF (100 mL). Nitrogen gas was bubbled through the solution for 10 min. K$_2$CO$_3$ (23.8 g, 0.172 mol) was then added.
and the mixture was stirred at 100 °C under a nitrogen atmosphere for 16 h. After cooling, water (200 mL) was added, and the product was extracted with chloroform (4 x 50 mL). The combined extracts were washed with water (40 mL), dried over MgSO₄, and filtered.

The resulting solution was evaporated under vacuum. The product was recrystallized from acetone at 4 °C. Yield: 27.1 g (84%). White solid, mp 119 °C. IR (KBr): 800, 825 (Ar−H), 1225, 1280 (Ar−O−C), 1510, 1605 (Ar), 3300 (OH) cm⁻¹. ¹H NMR (CDCl₃, 90 MHz): δ 0.88 (t, 6H, CH₃), 1.0–1.9 (m, 32H, CH₂), 4.0 (t, 4H, CH₂O), 6.88–7.52 (m, 3H, ArH). MS (EI): m/z 418 (M⁺). Anal. Calc'd for C₄₇H₃₆O₄: C, 76.76; H, 5.92. Found: C, 76.74; H, 5.97.

3.4-Bis(decyloxy)phenol (4). Benzenesulfonyl chloride (5.3 g, 0.074 mol) was dissolved in chloroform (100 mL). Methanol (100 mL), 35% H₂SO₄ (9.5 g, 0.098 mol), and concentrated H₂SO₄ (1 mL) were added, and the mixture was stirred at room temperature for 5 h. Water was added until two phases separated. The aqueous fraction was extracted with chloroform (2 x 50 mL). The combined organic solutions were washed with water (50 mL), dried over MgSO₄, filtered, and the solvent was evaporated under high vacuum. The resulting solid was recrystallized from acetone at 4 °C. Yield: 77.46 g (76.2 mol%). The resulting product was recrystallized from acetone at 0 °C. Yield: 108 g (93%) of a very viscous oil. The product was chromatographed on silica gel using 1:1 (v/v) ethyl acetate–hexane as the eluent, in portions that were sufficient for the next reaction step. A small amount was recrystallized from CDCl₃ to afford white needles, mp 87 °C. IR (CHCl₃): 3195 (CH₃), 2920 (CH₂), 1685, 2750 (H−C=O) cm⁻¹. ¹H NMR (CDCl₃, 100 MHz): δ 7.06 (s, 2H, ArH). MS (EI): m/z 506 (M⁺). Anal. Calc'd for C₃₆H₃₆O₄: C, 71.42; H, 9.82. Found: C, 71.45; H, 9.84.

4.5-Bis(decyloxy)-1,2-benzoquinone (5). The following solutions were prepared: tetrabutylammonium chloride (2 g, 7 mmol) in water (370 mL), KH₂PO₄ (1.5 g, 11 mmol) in water (100 mL), and phenol derivative 4 (7.5 g, 18 mmol) in THF (125 mL). Potassium nitrosodisulfonate (10 g, 44 mmol) was dissolved in chloroform (50 mL). The chloroform solution was washed with water (40 mL), dried over MgSO₄, and the solvent was removed by evaporation. The product was recrystallized from acetone at 4 °C. Yield: 6.0 g (80%). Orange solid, mp 119 °C. IR (KBr): 651 (ArBr), 1268 (ArO), 1510, 1605 (Ar), 3300 (OH) cm⁻¹. "H NMR (CDCl₃, 90 MHz): δ 0.85 (t, 6H, CH₃), 1.2–1.9 (m, 32H, CH₂), 4.0 (t, 4H, CH₂O), 6.68–7.65 (5.5 (s, 2H, ArH). MS (EI): m/z 406 (M⁺). Anal. Calc'd for C₄₄H₃₄ClO₄: C, 76.60; H, 11.40. Found: C, 76.70; H, 11.40. Compound 12 (13.5 g, 18 mmol) was added dropwise to the vigorously stirred solution (the addition can be facilitated by warming the viscous diosylate 10 with a heat gun). The reaction mixture was refluxed with vigorous stirring under nitrogen for 16 h. After the reaction was complete, the mixture was acidified with concentrated HCl, and the solvent was removed by evaporation under reduced pressure. The resulting solid was dissolved in chloroform, and the solution was washed with water (3 x), dried over MgSO₄, and concentrated. The product was recrystallized from acetone (2 x 250 mL). Yield: 8.7 g (58%). White powder, mp 129 °C. "H NMR (CDCl₃, 100 MHz): δ 0.88 (t, 6H, CH₃), 1.2–1.9 (m, 32H, CH₂), 4.0 (t, 4H, CH₂O), 6.68–7.65 (5.5 (s, 2H, ArH). MS (EI): m/z 506 (M⁺). Anal. Calc'd for C₃₆H₃₆O₄: C, 76.70; H, 11.40. Found: C, 76.60; H, 11.38.

1.2-Bis(2′-4′-dihydropyrimidinyl)ethyloxy)-4,5-dibromobenzene (11). A solution of compound 9 (68.4 g, 0.154 mol) in pyridine (125 mL) was cooled to −10 °C, and a solution of p-toluene sulfonamide (70 g, 0.367 mol) in pyridine (125 mL) was added dropwise with stirring. This mixture was stirred for 2 h at −5 °C and then for one night at 0 °C. The reaction mixture was poured over ice (400 g), acidified with 6 M aqueous HCl, and extracted with dichloromethane (2 x 250 mL). The combined extracts were washed with water (2 x 250 mL), dried over MgSO₄, and concentrated under vacuum. Yield: 108 g (93%) of a very viscous oil. The product was chromatographed on silica gel using 1:1 (v/v) ethyl acetate–hexane as the eluent, in portions that were sufficient for the next reaction step. A small amount was recrystallized from CDCl₃ to afford white needles, mp 87 °C. IR (CHCl₃): 3343 (OH) cm⁻¹. "H NMR (CDCl₃, 100 MHz): δ 7.09 (s, 2H, ArH). 7.16 (2 d, 4H, ArH). 7.35 (2 d, 4H, ArH). MS (EI): m/z 572 (M⁺). Anal. Calc'd for C₃₆H₃₆O₄: C, 76.60; H, 11.40. Found: C, 76.60; H, 11.38.
clooctadeca-2,11-diene][2,3-b:2′,3′-i:2′,3′-p:2″,3″-w]phthalocyanine (1a). A mixture of dicyanocrown ether 12 (0.75 g, 1.03 mmol) and (N,N-dimethylamino)ethanol (1 mL) was stirred and refluxed under a nitrogen atmosphere for 40 h. After cooling, chloroform (5 mL) was added, and the mixture was stirred for 3 h and then allowed to stand without stirring for 1 h. The precipitate was filtered off and recrystallized three times from chloroform. Yield: 0.245 g (30%). Dark green solid, K: M 148 °C. IR (CHCl₃): 863 (Ar—H), 909, 954 (Pc), 1022 (N—H), 1101 (Pc), 1134 (C—O—C), 1280 (Ar—O—C), 1529, 1603 (Ar/Pc), 2853, 2925 (CH₃), 3294 (N—H) cm⁻¹. ¹H NMR (CDCl₃, 400 MHz, 53 °C): δ 0.89 (br s, 24H, CH₃), 1.2—1.8 (m, 128H, CH₂), 3.6—4.9 (m, 80H, CH), 6.64 (s, 8H, ArH). Anal. Calcd for C₁₀₀H₈₅N₂₀₂SiCH₂: C, 67.36; H, 8.40; N, 3.72. Found: C, 67.18; H, 8.28; N, 3.74.

(Tetrakis[4,5-bis(decyloxy)benzo[1′,2′-k]-1,4,7,10,13,16-hexaoxyclooctadeca-2,11-diene][2,3-c]-1,3-diminoisoindoline) (13). Dicyanocrown ether 12 (720 mg, 1 mmol) and sodium methoxide (100 mg, 1.9 mmol) were placed in a three-necked flask equipped with a gas inlet and reflux condenser. The flask was flushed with anhydrous ammonia, and the mixture was refluxed, cooled to room temperature, and finally filtered. The protons of the Pc core were invisible, even at elevated temperatures. At room temperature, the compound was found to be slightly soluble in chloroform and toluene only. The solubility in these solvents could be improved with heating. The compound was first converted into the 1,3-diiminoisoindoline derivative 1a, compound 1b was found to be highly soluble in chloroform. A nicely resolved 'H NMR spectrum could be recorded, including the peaks of the Pc protons. This result suggests that aggregation of 1b is prevented by the axial hydroxy substituents. The UV/vis spectrum of 1b displayed a strong Q-band at 679 nm (see Figure 2), which is in line with values found for other dihydroxysilicon Pcs.

Results and Discussion

Synthesis and Characterization. The synthesis of phthalocyanine 1a has been described in a preliminary communication (10) and is summarized in Scheme 1. Synthetic details are given in the Experimental Section.

The dihydroxysilicon derivative 1b was synthesized in two steps from dicyanide 12 (Scheme 1). The latter compound was first converted into the 1,3-diiminoisoindoline 13 by reaction with ammonia in 1:1 (v/v) THF—methanol (yield 30%). Subsequent cyclization of 13 in quinoline in the presence of SiCl₄, followed by hydrolysis of the resulting dichlorosilicon Pc, gave 1b in 17% yield. In contrast to the metal-free derivative 1a, compound 1b was found to be highly soluble in chloroform. A nicely resolved 'H NMR spectrum could be recorded, including the peaks of the Pc protons. This result suggests that aggregation of 1b is prevented by the axial hydroxy substituents. The UV/vis spectrum of 1b displayed a strong Q-band at 679 nm (see Figure 2), which is in line with values found for other dihydroxysilicon Pcs.

We polymerized phthalocyanine 1b using two different methods, viz. as described by van der Pol et al. (14) and by Caseri et al. (15). The first method afforded a purple insoluble polymer, whereas the second method resulted in a blue pentamer. The
The second term in the above equation, therefore, is likely to be proportional to their mass. It can be concluded from the DSC, with the peak onset in the heating run found at 148 °C (cooling run, 105 °C). The transition enthalpy amounted to 130 kJ/mol. This value is ~20 kJ/mol larger than the value reported for the crystalline phase to mesophase transition of octa(dodecoxy)phthalocyanine. This is due to the fact that the substituents of la are larger than those of the aforementioned phthalocyanine.

Solid State and Mesophase Structures. The phase behavior of compound la has been described previously. One reversible transition from the solid to the mesophase was detected by DSC, with the peak onset in the heating run found at 148 °C (cooling run, 105 °C). The transition enthalpy amounted to 130 kJ/mol. This value is ~20 kJ/mol larger than the value reported for the crystalline phase to mesophase transition of octa(dodecoxy)phthalocyanine. This is due to the fact that the substituents of la are larger than those of the aforementioned phthalocyanine.

Figure 3 shows a plot of the transition enthalpies of octa(alcohol)phthalocyanines, together with that of la, versus the total molar mass of the flexible parts of the substituents. For the calculation of this molar mass, the rigid benzene rings of la, as well as the oxygen atoms connected to these benzene rings and those connected to the Pc cores of la and the other octa(alcohol)-Pcs, were excluded. The following linear relationship was found: \( \Delta H = 44.4 + 0.05m_{\text{flex}} \), with \( m_{\text{flex}} \) being the total flexible mass of the substituents in atomic mass units. At the crystalline phase to mesophase transition, the side chains of the octa(alcohol)phthalocyanines start to melt.

The second term in the above equation, therefore, is likely to be equal to the melting enthalpy of these chains, which appears to be proportional to their mass. It can be concluded from Figure 3 that la exhibits a melting behavior similar to that of the other liquid crystalline phthalocyanines. The melting enthalpy of these chains, which appears to be the melting enthalpy of these chains, which appears to be proportional to their mass. It can be concluded from Figure 3 that la exhibits a melting behavior similar to that of the other liquid crystalline phthalocyanines.
chains become oriented perpendicular to the c-axis and are
stacked with a separating distance of 4.3 Å. The tilt angle of
the Pc units is then calculated to be 38° (3.4/4.3 = \cos 38°).
As can be noticed from Figure 5a, some empty space may be
present in the crystalline structure. This space is probably
occupied by solvent molecules, since elemental analysis reveals
that chloroform is tightly bound in the solid material (vide
supra). The density of the unit cell is calculated to be 1.293
g/cm³. This value is quite reasonable and similar to the values
found for other octa(alkoxy)-Pcs in the crystalline phase.⁵⁴

The SAXS measurements on 1a at 170 °C confirmed the
presence of a mesophase at this temperature. The scattering
curve in Figure 4 shows an intense peak corresponding to a
spacing of 35.3 Å which can be attributed to the (100) reflection
of a hexagonal columnar (D₆) mesophase. Small peaks observed
at 17.3 and 13.4 Å are then related to the (200) and (210)
reflections, respectively. These values correspond to an inter-
columnar distance of 40.5 Å. This distance suggests that the
decoxy chains are folded backward, as shown in Figure 1b, or,
more probably, are in an interlacing position, as illustrated in
Figure 6. A very small band is observed around 3.34 Å (see
Figure 4B), which can be assigned to the intracolumnar stacking
period of the Pc cores. This reflection is somewhat diffuse,
indicating that the stacking is disordered. A broad and diffuse
band is visible at 4.4 Å, which is characteristic for disordered
side chains.⁵⁴ This band is accompanied with a small reflection
at approximately twice this spacing, viz. 8.7 Å. These data
suggest that the Pc molecules are stacked cofacially in a
staggered conformation. A density of 0.89 g/cm³ can be
indicated that the stacking is disordered. A broad and diffuse
calculated for the unit cell of the mesophase of 1a, which is
very close to the density of metal-free octa(dodec oxy)-Pc in its
mesophase.⁵⁴

An interesting feature of the structure of 1a is the presence of
ion channels: in the crystalline phase, the molecules are in
an eclipsed position, with the crown ether rings superimposed.
Also, in the mesophase, ion channels exist because the next-
nearest crown ether rings are superimposed.

Aggregation in Solution. (i) UV/Vis Spectroscopy. The
aggregation behavior of the metal-free phthalocyanine 1a in
chloroform solution was studied by UV/vis spectroscopy.
Figure 7 shows the absorption spectra of a diluted solution (11
µM) of 1a recorded at room temperature and at 50 °C. The
high-temperature spectrum displays the split Q-band that is
characteristic for nonaggregated metal-free Pcs, with maxima
appearing at 661 and 700 nm. The spectrum recorded at room
temperature is broadened, and the maximum is blue-shifted to
614 nm. According to the molecular exciton theory,¹⁷ the blue

shift points to stacking of the Pc molecules, the tilt angle being
smaller than 35.3°. The exciton splitting, 2e, which is due to
the coupling of the π-systems including long-range interactions,
is given by

\[
2e = 2[-(\Delta E_a - \Delta E_m) + \Delta D] =
-4.8 \frac{[N-1][|M|^2]}{N^2r^2} (\cos \alpha - 3 \cos^2 \theta)
\]

where \(\Delta E_a\) and \(\Delta E_m\) are the energies corresponding to the
aggregate and monomer Q-band maxima, and \(\Delta D\) is the
lowering of the difference between the ground and excited states
due to van der Waals interaction (\(\Delta D = -615 \text{ cm}^{-1}\))¹⁸
The center-to-center distance between the molecules, \(r\), is related to the angle \(\theta\) between the transition dipole moments (M)
according to \(r = r_{\text{mon}} \sin \theta\), where the minimum van der Waals
distance \(r_{\text{mon}} = 3.4 \text{ Å}\. \ [M]^2\) can be calculated from the oscillator
strength \(f\). The latter value was determined by integration of the
monomeric Q-band, resulting in \(f = 0.59 \text{ erg cm}^{-1}\), from
which a value of \([M]^2 = 2.17 \times 10^{-19} \text{ cm}^2\) could be derived.
These values are very close to the literature values found for a
substituted dihydroxysilicon phthalocyanine.¹⁸ When the
aggregation number \(N\) is considered to be large, which is justified
as shown in the next section, then \((N - 1)/N\) can be taken as
unity by good approximation. With the values of \(\Delta E_a\) and \(\Delta E_m\)
being 16.287 and 14.695 cm⁻¹, respectively, and taking an
eclipsed conformation (torsion angle \(\alpha = 0°\)), a value of \(\theta = 60°\) was calculated using formula 1. This value corresponds to
a tilt angle of 30° and a center-to-center distance \(r = 3.9 \text{ Å}\. If
a staggered conformation is assumed (\(\alpha = 45°\)), then a tilt angle
of 24° and \(r = 3.7 \text{ Å}\ are calculated. The latter structure cannot
be ruled out, but the former is more plausible since the structural
parameters are close to those of the \(\alpha\)-modification of solid
unsubstituted Pcs.¹⁹

The aggregation of 1a was found to be not only temperature
dependent but also concentration dependent. The heat of
association can be estimated from the temperature and concen-
tration dependency of the UV/vis spectrum. Figure 8a shows a
plot of the ratio of the absorbances of the aggregate and the
monomer versus the temperature for four different concentra-
tions of 1a in chloroform. The aggregation number is related to
this absorbance ratio: in other words, at a fixed ratio, the
weight fraction of the monomer, \(m_1\), is constant. For a constant
\(m_1\), the following relationship holds for linear aggregates of
uniform structure:²⁰

183, 387–402.
The existence of strong attraction between the molecules of 

a high value has never before been reported in the literature on the packing of particles in 

the reciprocal temperature (Figure 8b) afforded a straight line, from which a value of $W_1 = -125$ kJ mol$^{-1}$ was derived. To our knowledge, such a high value has never before been reported in the literature for the aggregation of PCs. For example, the heats of association of tetrakis(sulfonatophthalocyanines) in water are 2.5–5 times lower.$^{21}$ The driving forces for aggregation in this literature example are π–π interactions, van der Waals interactions, and, in particular, hydrophobic effects. The aggregation of Pc 1a in chloroform is mainly the result of π–π and van der Waals interactions. These interactions are exceptionally strong in our system because they can take place not only between the Pc cores but also between the peripheral benzene units of different eclipsed molecules, resulting in a five-points attachment. The existence of strong attraction between the molecules of 1a leads to extremely long linear aggregates, as will be shown next.

(ii) Electron Microscopy. Interestingly, when a more concentrated solution of compound 1a in hot chloroform was cooled to room temperature, a firm gel was formed. This was observed at concentrations as low as 7 mg mL$^{-1}$. A sample having the latter concentration and a sample of 1 mg mL$^{-1}$ were studied with transmission electron microscopy (TEM). Some representative TEM pictures are shown in Figure 9. At a concentration of 7 mg mL$^{-1}$, dense networks of large fibers are observed (micrograph a). Micrograph b was taken from the same sample with the same magnification but at another position. The latter picture clearly shows that the fibers consist of bundles of single parallel strands. Some isolated single strands can also be seen. The length of the fibers is of the order of a few micrometers. A magnification of micrograph b is shown in Figure 9c. The dimensions of the uncovered and platinum-covered parts suggest that the strands are cylinders with a diameter of ~60 Å, as illustrated in Figure 9d. This value matches very well with the diameter of a molecule of 1a as estimated from CPK molecular models (see Figure 1a). At the lower concentration of 1 mg mL$^{-1}$, fibers are also observed, but they are smaller in length and no network is formed (not shown).

From these results, we may conclude that compound 1a self-assembles in chloroform solution to produce linear aggregates. The stacks that are visualized by TEM have molecular thickness and are built up from ~10 molecules. Such a large aggregation number in solution is unprecedented and reflects the existence of large attracting forces between the molecules. As discussed in the UV/Vis Spectroscopy section, the molecules are most likely stacked in an eclipsed conformation. Such a conformation was also shown to be present in the crystalline state of 1a. A (close to) eclipsed conformation places the peripheral benzene units in a favorable position for π–π interaction. This has interesting consequences. The isolated strands of molecules 1a can be considered as being molecular cables (Figure 10). Such cables contain a central electron-conducting wire of stacked Pc molecules, four ion channels formed by the crown ether rings, and a surrounding insulating hydrocarbon mantle. The electron-and-ion-conducting properties of the solid and the mesophase, as well as of the aggregates in solution, will be the subject of future studies.

Formation of Monolayers and LB Films. It was not

\[ W_1 = \frac{R \Delta \ln c}{\Delta(1/T)} \]
possible to form monomolecular layers from 1a at the air—water interface due to the strong tendency of the molecules to aggregate. The dihydroxysilicon derivative 1b does not aggregate in organic solution. This compound turned out to be suitable for the formation of LB films, as will be shown next.

A solution of 1b in chloroform was spread on a water surface at 20 °C. Figure 11, curve A, shows the surface pressure—surface area isotherm for 1b at a subphase temperature of 20 °C. Curve A, pure water; curve B, 1 M KCl. The solid circles represent the K⁺ ions that are bound in the crown ether rings. The absence of such islands in the case of Pc 1b suggests that this molecule does not aggregate into large domains. This behavior is probably related to the axial hydroxy groups in the center of the molecule. These groups not only hinder the cofacial aggregation but also increase the hydrophilicity of the molecular core of 1b, which favors the side-on orientation of the molecules on the water surface.

![Figure 10. Schematic representation of the multiwired molecular cable formed by aggregation of 1a in chloroform.](image)

We were able to transfer monolayers of 1b onto a hydrophobic glass substrate by vertically dipping the substrate through the monolayer at a surface pressure of 15 mN m⁻¹. Y-type deposition occurred with a transfer ratio of ~1.0 with each upstroke transfer. The downstroke transfer resulted in a lower ratio of ~0.5, which might be the result of an incomplete transfer or a conformational change of the molecules during the deposition process. The UV–vis spectrum of a 2 × 10-layer film is shown in Figure 2. The maximum of the Q-band is seen at 697 nm, which is red-shifted compared to the Q-band of a chloroform solution of 1b. An additional small peak is observed at 630 nm. These spectral changes can be attributed to exciton splitting due to a slipped cofacial stacking of the Pc units with a tilt angle larger than 35.3 °.

The observed energy difference between the allowed and partly forbidden transition, 26, amounts to 1526 cm⁻¹. We can calculate the aggregation number N from this exciton splitting with the help of formula (1), taking for the center-to-center distance r the value that was previously measured for the dihydroxysilicon derivative of octa(alkoxy)-Pc, i.e., 4.5 Å.¹⁶ This value corresponds to θ = 49.1° or a tilt angle of 40.9°. The numerical factor in formula 1 has to be reduced to 4.0 to exclude longer range interactions. If we take for the tilt angle between the molecules the value α = 0°, the upper limit of N can be calculated; for α = 45°, the lower limit is found. These limits amount to 2.3 and 1.4, respectively. From these numbers, we may conclude that the molecules in the LB film are present as dimers. This result is in line with the observed Y-type deposition, which results in an alternating “face-to-face” and “tail-to-tail” multilayer structure.

chloride as the subphase. It can be seen that a first rise of the surface pressure is found at an area of 440 Å²/molecule, which is much larger than the observed molecular area with pure water as the subphase. The increased area corresponds to a circle with a diameter of 23.7 Å. This value is very close to the diameter of the molecule when the Pc core as well as the four crown ether rings lie completely flat on the water surface. This situation is illustrated in Figure 12b. It is obvious that the alkali metal ions induce a conformational change of the molecules of 1b at the air—water interface. The binding of the metal ions in the crown ether rings makes these macrocycles more rigid and more hydrophilic, favoring their interaction with water. Figure 11 shows that a phase transition takes place when the monolayer area of 1b on 1 M KCl is decreased. A second rise of the surface pressure is found at an area that is close to the molecular area on pure water. Probably the decrease of the monolayer area forces the crown ether rings to lift from the water surface. It is not clear at present whether or not this results in the release of the cation.

We recorded isotherms at various subphase concentrations between 0 and 1 M KCl and measured the area per molecule related to the first surface pressure rise. Figure 13 shows a plot of the area per molecule versus the potassium ion concentration. We assume the following equilibrium and the corresponding binding constant $K$:

$$\text{Pc} + 4\text{K}^+ \leftrightarrow \text{Pc(K}^+)_4$$

$$K = \frac{[\text{Pc(K}^+)_4]}{[\text{Pc}][\text{K}^+]^4}$$

(3)

The concentrations of the complex and the free ligand can be calculated from the measured area per molecule, $A$, to be

$$[\text{Pc(K}^+)_4] = \frac{A - A_0}{A_c - A_0} [\text{Pc}]_0$$

(4)

and

$$[\text{Pc}] = \frac{A_c - A}{A_c - A_0} [\text{Pc}]_0$$

(5)

where $A_0$ is the area per molecule on pure water, $A_c$ is the area per molecule when all complexation sites are occupied, and $[\text{Pc}]_0$ is the total Pc concentration. By good approximation, the concentration of free potassium, $[\text{K}^+]$, is assumed not to be influenced by binding. This is true in our case because $[\text{K}^+] \gg [\text{Pc}]_0$. Combining eqs 3–5 results in the following relationship:

$$K = \frac{A - A_0}{A_c - A} [\text{K}^+]^{-4}$$

(6)

Computer fitting of the data points of Figure 13 to this equation provides a binding constant of $K = 48 \text{ M}^{-1}$. This value is remarkably close to the value of the binding constant of dibenzo-18-crown-6 with potassium ions in aqueous solutions, viz. 46.5 M⁻¹.²⁵

Concluding Remarks

We have described several ways to construct supramolecular structures from the novel phthalocyanine building blocks 1a, b by a process of self-organization: in the solid state, in the liquid crystalline state, in solution, and as monolayers and LB films. Two features are assumed to be mainly responsible for the spontaneous formation of the supramolecules: the high degree of shape anisotropy of the Pcs, which favors their packing into columnar assemblies, and the fact that the molecules are strongly held together by van der Waals and $\pi-\pi$ interactions.

Very unexpected is the self-organization of 1a in chloroform solution, leading to extremely long individual stacks of molecules. Aggregation of Pcs is usually observed in more polar solvents.²⁶ The very high association energy which we determined from UV/vis spectroscopy experiments indicates that large attracting forces are responsible for this unique behavior. One-dimensional electron-conducting wires and ion-conducting channels are formed in the stacks, which are surrounded by an insulating hydrocarbon mantle.

The possibility to form monolayers and LB films from phthalocyanine 1b is of interest for future applications. The ability of this compound to bind cations in its crown ether rings was demonstrated in monolayers. Interestingly, the obvious nonlinearity of the titration curve of Figure 13 is suggestive of a positive cooperative effect in the complexation of potassium ions, i.e., a measurable change is observed only above a certain critical concentration of these ions. This result may be important for application in molecular switches based on the complexation of ions.²⁶

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