The following full text is a publisher's version.

For additional information about this publication click this link. http://hdl.handle.net/2066/16348

Please be advised that this information was generated on 2019-04-01 and may be subject to change.
Model for Dissociative H₂ Chemisorption on Transition Metals

H. Deuss and A. van der Avoird
Institute of Theoretical Chemistry, University of Nijmegen, Nijmegen, The Netherlands
(Received 5 February 1973)

Earlier calculations have shown that hydrogen atoms can be strongly adsorbed on nickel and on copper surfaces although the character of the chemisorption bond is different. In the case of nickel, both 3d and 4s electrons are involved, whereas for copper the bond is formed exclusively by 4s electrons. On the basis of this information a model is worked out for dissociative chemisorption of H₂ molecules, which is experimentally found to occur on nickel but not on copper. In this model the activity of 3d and 4s electrons is studied separately. The 3d electrons cause an unactivated dissociative chemisorption of H₂; the 4s electrons alone do not chemisorb an H₂ molecule unless a high activation-energy barrier is surmounted.

INTRODUCTION

It is well established experimentally⁷ that some transition metals, for instance nickel and platinum, can dissociatively adsorb H₂ molecules even at very low temperatures. This dissociative chemisorption, which is very important for the catalytic activity of transition metals in many reactions,⁸ requires practically no activation energy. It is also known⁷ that copper, which is very similar to nickel except that it has a completely filled d band, does not chemisorb H₂ molecules dissociatively at room temperature. An explanation which ascribes the activity of nickel to the partial filling of the d band therefore becomes very plausible. Much work still has to be done, however, to give this hypothesis a more quantitative basis. Moreover, it must also be explained why predissociated hydrogen is adsorbed by copper almost as strongly as by nickel.⁹,¹⁰

Some time ago we proposed a model for dissociative chemisorption of H₂ on a nickel surface.⁵ We assumed that the two hydrogen atoms interact mainly with two nickel atoms. The unpaired d electrons of nickel were represented by a single effective electron in a spherical orbital on each nickel atom. The argument justifying a spherical distribution was that in the bulk of the metal the splitting between d orbitals is very small compared to the energies involved in chemisorption. The interaction energy between the four atoms was calculated by a perturbation method which takes exchange as well as van der Waals forces into account.⁶,⁷ This model indicated the possibility of unactivated H₂ dissociation which was mainly caused by the pair and three-body exchange interactions calculated in first-order perturbation theory.

Since then, we have obtained more information about the chemisorption bond by molecular-orbital calculations for hydrogen atoms adsorbed on clusters of nickel or copper atoms.⁸ We have found that, if a hydrogen atom is adsorbed on top of a nickel atom in a cluster, the main interaction takes place with the 3d₂ and the 4s orbitals of this nickel atom. This conclusion applies to chemisorption on the different surfaces: (100), (110), and (111), although, of course, the interaction is influenced by the neighboring nickel atoms. (On each surface, the 3d₂ orbital has to be understood as the d₂ orbital of which the axis points perpendicularly out of the surface.) It is the chemisorption interaction itself, therefore, rather than the bulk structure which determines the nature of the d orbitals that are mainly involved. The covalent bonding of the hydrogen atom with the 4s orbital is comparable in strength with the bond involving 3d₂. For copper we have found that the 3d orbitals do not take part in the bonding of the hydrogen atom, but that the increased bonding to the 4s orbital on the underlying copper atom causes the adsorption energy of a hydrogen atom to be of the same magnitude as for nickel.

This molecular-orbital (MO) model cannot be used to study dissociative chemisorption of H₂ molecules, since this requires interaction energies over a wide range of interatomic distances up to different dissociation limits. It is well known that the conventional MO method is not suitable for this purpose (except in special cases, such as interactions between closed-shell systems). Therefore, we used the information obtained from the MO calculation to improve our original model for dissociative chemisorption of H₂.

DESCRIPTION OF THE MODEL

The MO calculations for hydrogen atoms on clusters of approximately 10 metal atoms showed that the main interaction takes place between a hydrogen atom and the underlying metal atom. Initially, it is justifiable, therefore, to include only two metal atoms in our model for dissociative H₂ chemisorption. On these metal atoms we place an electron either in a 3d₂ or in a 4s orbital (see Fig. 1). The
interaction energy with two hydrogen atoms which contain an electron in a 1s orbital is calculated in first-order-exchange perturbation theory:

$$\Delta E = \langle \phi_0 | (a \mathcal{U}) | \phi_0 \rangle / \langle \phi_0 | \alpha | \phi_0 \rangle.$$  

The unperturbed wave function $\phi_0$ is a product of four atomic orbitals, the two 3$d^2$ orbitals or the 4$s$ orbitals on the metal atoms and the two 1$s$ orbitals on hydrogen. The radial parts of these orbitals are approximated by linear combinations of simple Gaussian functions. The exponents and coefficients are obtained from atomic SCF calculations for nickel and copper$^9,10$ and by a fit to the exact 1$s$ orbital for hydrogen$^{11}$ (see Table I). The operator $\mathcal{U}$ consists of the interatomic interaction terms in the total four-electron Hamiltonian. $\alpha$ is an operator which takes antisymmetry of the total wave function into account as well as the fact that this wave function should be a singlet eigenfunction of the total-spin operator. We can construct two such singlet wave functions and, actually, a $2 \times 2$ secular problem had to be solved in order to compute the energy of the lowest singlet state.$^{12}$

If the unperturbed Hamiltonian $H_0 = H - \mathcal{U}$ would have the products of atomic wave functions as its exact eigenfunctions, the first-order interaction energy would equal the interaction energy calculated by the valence-bond method. The metal orbitals are not eigenfunctions of their respective parts in $H_0$, however, and, therefore, both the perturbation method and the valence-bond method yield an approximation to the exact interaction energy. The second-order energy is omitted as it showed no essential contribution to the process of dissociative chemisorption.$^3$ The first-order interaction energy is calculated as a function of $h$, the height of the $H_2$ molecule above the Ni atoms, and $d$, half the distance between the hydrogen atoms (see Fig. 1). In the case of nickel, we used both 3$d^2$ and 4$s$ orbitals and a Ni-Ni distance of 4.73$\sigma_0$ (the nearest-neighbor distance of the metal), for copper we only used 4$s$ orbitals (since the 3$d$ orbitals do not cause adsorption bonding of hydrogen atoms) and a Cu-Cu distance of 4.82$\sigma_0$.

**RESULTS AND DISCUSSION**

The results are shown in the contour maps of Figs. 2(a) and 2(b), where curves of equal interaction energy are plotted as a function of $h$ and $d$. The interaction energy of $H_2$ with the copper 4$s$ electrons is not shown as it is very similar to the one with nickel 4$s$ electrons, which is drawn in Fig. 2(b).

At large distance $h$ from the surface, the hydrogen atoms attract each other to form a stable molecule (equilibrium distance $2d = 1.58\sigma_0$, binding energy 72 kcal/mole, the valence-bond results). The maps clearly show what happens if this hydrogen molecule approaches the "metal surface." The 3$d^2$ electrons of the two nickel atoms in the "surface" [Fig. 2(a)] attract the molecule and, on coming closer to the metal atoms, it will be dissociated. This process of adsorption and dissociation requires no activation energy. The two separated hydrogen atoms will be strongly bound to the nickel atoms. The 4$s$ electrons of nickel [Fig. 2(b)] or copper cause very different behavior. The hydrogen molecule is repelled and can only be chemisorbed and dissociated if an activation barrier of approximately 50 kcal/mole for nickel or 45 kcal/mole for copper can be surmounted. The hydrogen atoms are then bound.

An analysis of these results shows that the three- and four-atom contributions to the interaction energy of $H_2$ with nickel and copper are very different. In nickel, the three- and four-atom contributions are comparable, whereas in copper the four-atom contributions are more important.

**TABLE I. Atomic orbital data.**

<table>
<thead>
<tr>
<th>Atomic orbital</th>
<th>Exponents ($\alpha \phi_0$)</th>
<th>Contraction coefficients</th>
<th>Refs</th>
</tr>
</thead>
<tbody>
<tr>
<td>Ni 3d</td>
<td>6.731</td>
<td>0.40164</td>
<td>9</td>
</tr>
<tr>
<td></td>
<td>1.353</td>
<td>0.78272</td>
<td></td>
</tr>
<tr>
<td>Ni 4s</td>
<td>0.1174</td>
<td>0.40986</td>
<td></td>
</tr>
<tr>
<td></td>
<td>0.05894</td>
<td>0.53552</td>
<td>10</td>
</tr>
<tr>
<td></td>
<td>0.02428</td>
<td>0.09184</td>
<td></td>
</tr>
<tr>
<td>Cu 4s</td>
<td>0.1269</td>
<td>0.41618</td>
<td></td>
</tr>
<tr>
<td></td>
<td>0.06575</td>
<td>0.53280</td>
<td>10</td>
</tr>
<tr>
<td></td>
<td>0.02122</td>
<td>0.09329</td>
<td></td>
</tr>
<tr>
<td>H 1s</td>
<td>0.1514</td>
<td>0.64767</td>
<td>11</td>
</tr>
<tr>
<td></td>
<td>0.6813</td>
<td>0.40789</td>
<td></td>
</tr>
<tr>
<td></td>
<td>4.500</td>
<td>0.07048</td>
<td></td>
</tr>
</tbody>
</table>
energy are of essential importance: The sum of
pair-interaction energies would cause a different
picture [Fig. 3(a) and 3(b)]. In the case of nickel
3d,2 orbitals, the hydrogen molecule would be ad­
sorbed, but a small activation barrier for dissoci­
ation would remain. This barrier is erased by the
three- and four-atom contributions. With the 4s
orbitals the effect of these contributions is even
more important. The pair interactions alone would

cause an H2 molecule to be attracted instead of repelled.

This conclusion about the importance of at least
three-center contributions, which is in agreement
with the results from our earlier model,5 prevents
the use of simple pair potentials of the Lennard-
Jones or other types. It requires an approach in
which all exchange effects (also three- and four-
electron interatomic permutations) are included.
The present model is one of the simplest possibili­
ties to perform this.

One additional remark about the model must be
made. If one considers the interaction energy with nickel 3\(d^2\) electrons only, one finds that the hydrogen atoms could come very close to the nickel atoms. This is not caused by the approximate form of the 3\(d^2\) orbitals; a fit of their radial parts with three Gaussians instead of two yielded practically the same result. Neither is it owing to our use of perturbation theory with atomic functions that are not exact eigenfunctions of the unperturbed Hamiltonian. When we approximated the interaction energy using the valence-bond method we found a similar result. The reason for this effect is probably that the interactions with the 4s electrons, the other 3d electrons and the core electrons must still be included. These interactions are very likely to be repulsive if a hydrogen atom comes close to the nickel atom.

Summarizing, the present model allows the following conclusions. For nickel, where MO calculations have shown a strong adsorption bond between a hydrogen atom and the 3\(d^2\) and 4s electrons of the underlying nickel atom, we find that the effect of the 3\(d^2\) orbitals could enable an \(\text{H}_2\) molecule to become dissociatively chemisorbed without an activation energy. For copper, where only the 4s orbitals take part in the covalent chemisorption bond with a hydrogen atom, the chemisorption and dissociation of an \(\text{H}_2\) molecule would require an activation energy of approximately 45 kcal/mole. For these results the three- and four-center interactions are important. Presently, we try to extend the model in order to calculate the combined effect of 3\(d^2\) and 4s electrons for nickel and, moreover, to include the effect of the other electrons as well.

1D. A. Cadenhead and N. J. Wagner, J. Catal. 21, 312 (1971), and references therein.
9G. A. van der Velde (private communication).