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LANTHANIDE-INDUCED CONTACT SHIFTS IN POLYGLYCOLDIMETHYLEthers

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Complexation of polyglycoldimethylethers (glymes) of the general formula
CH₃(CH₂CH₂O)ₙCH₃ with certain aromatic radical ion salts such as coronene
sodium is known to result in strong downfield nuclear magnetic resonance shifts
of the glyme protons (1,2). The shifts, amounting to as much as 12 ppm and
interpreted in terms of a Fermi contact mechanism, produce a first order spectrum
for the CH₂ protons, each peak representing a combination of two or more CH₂
groups.

The large proton nmr shifts recently reported for many organic compounds
in the presence of paramagnetic lanthanide complexes such as tris(dipivalomethanato)
europium and -praseodymium (3-9) induced us to look at mixtures of
these lanthanide shift reagents with the polyfunctional glymes. In the actual
experiments, small quantities of glyme (some of them partially deuterated) with
n ranging from 1 to 6 were added to a 0.15 M solution of a lanthanide complex
in CCl₄, and the nmr spectra recorded on a Varian A60A spectrometer. In a
number of experiments the molar ratio of lanthanide to glyme was varied from
0.2 to 4.0

The spectra of the 1:1 glyme-Eu adducts reveal a sharp peak representing
the two terminal CH₃ groups, and further downfield a series of well resolved
absorptions each representing two CH₂ groups. Addition of more glyme yields a
rapid exchange spectrum, and plots of the proton shifts versus the ratio of the
concentration of the 1:1 adduct to that of total glyme are linear and passing

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through the origin, implying that at 0.15 M of the 1:1 mixture essentially all of the glyme is bound to the Eu(dpm)₃.

It is interesting that the spectrum of a glyme 5/Eu(dpm)₃ mixture (the number following the glyme refers to the number of oxygen atoms in the glyme) clearly exhibits four triplets, each representing a combination of two CH₂ groups. For a similar mixture with glyme 4 one of the CH₂...CH₂ peaks is a singlet. Glyme 3 shows two triplets, while a glyme 6/Eu(dpm)₃ spectrum consists of four well resolved triplets and a singlet, in addition to the CH₃ group. Apparently, the two CH₂ groups representing one nmr absorption are equally far removed from the center of the molecule. Using glyme 4 as an example:

\[
\text{CH}_3\text{OCH}_2\text{CH}_2\text{OCH}_2\text{CH}_2\text{OCH}_2\text{CH}_2\text{OCH}_3
\]

\((3)(2)(1)(1)(2)(3)\)

the C₁ protons will give a singlet, while triplets are expected for the C(2) and C(3) protons. Glymes with an odd number of oxygen atoms (glyme 3 or 5) yield only triplets.

The following deuterated glymes were prepared for further peak assignments:

\((\text{CH}_3\text{OCH}_2\text{CH}_2\text{OCD}_2\text{CH}_2)_2\text{O}, (\text{CH}_3\text{OCH}_2\text{CD}_2\text{OCH}_2\text{CH}_2)_2\text{O}\) and \(\text{CH}_3\text{OCH}_2\text{CD}_2\text{OCH}_2\text{CH}_2\text{OCH}_2\text{CH}_2\text{OCH}_3\).

These compounds were sufficient to assign all glyme 4 and glyme 5 absorptions, since the CH₂ protons adjacent to a CD₂ group are changed from a triplet to a singlet. Also, some interesting isotope effects on the chemical shifts were found, with the CH₃ absorption splitting up in a doublet for the asymmetrically deuterated glyme 4 (20 cps peak separation with Pr).

The respective chemical shifts are listed in Table I, all values being relative to the chemical shifts of the glyme protons in pure glyme. Results obtained with praseodymium (dpm)₃ are also included. The most striking observation concerns the Pr/Eu shift ratio for the various glyme protons. For the two 1:1 lanthanide-glyme 4 complexes the Pr/Eu shift ratio is 3.0 for the CH₃ group and 3.2 for the C(1) protons. However, the ratios are only 1.9 for the C(2) protons and 1.7 for the C(3) protons. A similar behavior is observed for glyme 5, with the respective Pr/Eu shift ratios being 2.9, 3.0 and 2.9 for the CH₃, C(1) and C(2) protons, and only 1.9 and 1.7 for the penultimate CH₂(C3) and terminal CH₂(C4) groups. These data suggest that it is not always
Table I

Proton NMR Shifts of Glyme 4 and Glyme 5 Adducts of Eu(dpm)$_3$ and Pr(dpm)$_3$ in CCl$_4$ at 40°C, relative to the Chemical Shifts of Non Complexed Glyme (both measured with respect to TMS). Concentration of Paramagnetic Complex: 0.15 M.

<table>
<thead>
<tr>
<th>Mixture</th>
<th>CH$_3$</th>
<th>C(1)</th>
<th>C(2)</th>
<th>C(3)</th>
<th>C(4)</th>
</tr>
</thead>
<tbody>
<tr>
<td>Eu/G4 = 1:0.2</td>
<td>-567</td>
<td>-521</td>
<td>-1253</td>
<td>-1215</td>
<td></td>
</tr>
<tr>
<td>Eu/G4 = 1:1</td>
<td>-253</td>
<td>-385</td>
<td>-617</td>
<td>-562</td>
<td></td>
</tr>
<tr>
<td>Eu/G4 = 1:2</td>
<td>-124</td>
<td>-212</td>
<td>-318</td>
<td>-282</td>
<td></td>
</tr>
<tr>
<td>Pr/G4 = 1:0.2</td>
<td>1341</td>
<td>2120</td>
<td>2065</td>
<td>2011</td>
<td></td>
</tr>
<tr>
<td>Pr/G4 = 1:1</td>
<td>763</td>
<td>1251</td>
<td>1200</td>
<td>977</td>
<td></td>
</tr>
<tr>
<td>Pr/G4 = 1:2</td>
<td>373</td>
<td>635</td>
<td>599</td>
<td>497</td>
<td></td>
</tr>
<tr>
<td>Eu/G5 = 1:1</td>
<td>-208</td>
<td>-256</td>
<td>-335</td>
<td>-519</td>
<td>-465</td>
</tr>
<tr>
<td>Pr/G5 = 1:1</td>
<td>604</td>
<td>767</td>
<td>965</td>
<td>965</td>
<td>809</td>
</tr>
</tbody>
</table>

justified to convert proton shifts obtained for one complex into those for another lanthanide complex by applying a constant shift ratio factor (8,9).

Our tentative conclusion is that the Eu or Pr ion is chelated predominantly to the outer two oxygen atoms of the glyme chain, bringing the two CH$_2$ groups connecting these two oxygen atoms in close proximity to the paramagnetic ion. This apparently results in a considerable Fermi contact shift on these protons, in addition to the expected dipolar shift. Both kinds of shift are downfield for the Eu-glyme adduct, but they are in opposite direction for the Pr complex. The combined effects lead to the low shift ratios found for these protons.

Supporting evidence for the simultaneous binding of two glyme oxygen atoms to Eu or Pr is found in the behavior of ortho and meta-dimethoxybenzene. The CH$_3$ protons of the ortho compound in a 1:1 mixture with Eu(dpm)$_3$ are shifted to the same extent as found for 1,2 dimethoxyethane, while the CH$_3$ shift for the meta derivative is concentration dependent and, at 0.15 M, amounts to only about 1/5 of that of the ortho compound. Also, 1,2 dimethoxyethane is much
stronger bound to Eu(dpm)$_3$ than ethers with one oxygen atom, including tetrhydrofuran.

The nmr patterns of the glyme adducts (e.g., the identical shifts of CH$_2$ groups equally far removed from the center of the glyme molecule) suggest a rapid intramolecular exchange of the lanthanide reagent between -OCH$_2$CH$_2$O- moieties on opposite ends of the chain and/or an intermolecular exchange between complexed glyme molecules. Also, when the lanthanide/glyme ratio exceeds unity, a second paramagnetic complex binds to the glyme (not observed for DME), causing an even further shift of the glyme protons (See Table 1). Additional work on these interesting chelating systems are in progress in order to compare their behavior with that found for the radical anion shift reagents.

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