The following full text is a publisher's version.

For additional information about this publication click this link.
http://hdl.handle.net/2066/143097

Please be advised that this information was generated on 2020-01-11 and may be subject to change.
COBALT(III) COMPLEXES WITH MIXED CHELATING LIGANDS

H.C. Brinkhoff

Department of Inorganic Chemistry, University of Nijmegen,
Driehuizerweg 200, Nijmegen, The Netherlands
(Received 12 November 1970)

As recently was reported (1) potassium bis biuretatocobaltate(III) \([\text{KCo(bi)}_2]\), which shows a striking magnetic behavior having a spin-triplet ground state, forms diamagnetic bis adducts with monodentate nitrogenous bases. The existence of successful addition reactions of this type and our interest (2) in complexes with dithioligands prompted the investigation of the reaction of KCo(bi)_2 with N,N-dialkyldithiocarbamate ion (R_2dtc^-) and O-alkylxanthate ion (Rxan^-) respectively.

It is found that, when equivalent amounts of KCo(bi)_2 and R_2dtc^- are reacted in aqueous solution, the initial suspension of KCo(bi)_2 disappears, and a blue-violet complex ion \([\text{Co(bi)}_2(R_2dtc)]^{2-}\) is formed according to the reaction equation:

\[
\text{KCo(bi)}_2 + R_2dtc^- \rightarrow [\text{Co(bi)}_2(R_2dtc)]^{2-} + \text{K}^+
\]

This reaction had to be carried out at 0°C, because otherwise the green complex ion \(\text{Co(bi)}(R\_dtc)^-\) is rapidly formed. The formation of the latter ion is favored, when two instead of one equivalent \(R_2dtc^-\) are used and when the reaction is carried out at room temperature. The reaction equation is:
It seemed not preferable to use an excess of $R_2dtc^-$ in the latter reaction because in that case the competitive formation of $Co(R_2dtc)_3$ (insoluble in aqueous solution) becomes important. Often, the isolation of the complex ions appeared to be difficult; we only succeeded in preparing $K_2Co(bi)_2(Et_2dtc)_2$ 3 $H_2O$ and $NbU_4Co(bi)(Bu_2dtc)_2$ in pure state. These complexes are diamagnetic in the solid state at room temperature, as can be expected for an approximately octahedral $Co^{3+}$ ion with a $d^6$ configuration. Conductivity studies, performed in nitrobenzene solution at $25^\circ$ on $NbU_4Co(bi)(Bu_2dtc)_2$, proved this compound to be a 1:1 electrolyte.

Using $Rxan^-$ instead of $R_2dtc^-$ in the above reactions, products are formed, whose colors (blue-violet and green respectively), suggest the existence of the complex ions $[Co(bi)_2(Rxan)]^{2-}$ and $[Co(bi)(Rxan)_2]^-$ respectively. These complex ions have so far not been isolated.

The combined results are clear evidence for the formation of the six-coordinated $Co(III)$ ions $Co(bi)_2L_2^-$ and $Co(bi)L_2^-$, when $KCo(bi)_2$ is reacted with a monoanionic chelating dithioligand $L$.

It was also possible to prepare $Co(Et_{x}an)(Et_2dtc)_2$ by reacting appropriate amounts of $Co(Et_{x}an)_3$ and $Co(Et_2dtc)_3$ in chloroform. The PMR spectrum of $Co(Et_{x}an)(Et_2dtc)_2$ proves the octahedral coordination symmetry of the cobalt showing a complicated fine structure for the N—CH$_2$ proton resonances (fig.). The latter is due to the diastereotropic
nature of the N—CH$_2$ protons. Such a phenomenon was recently reported (3) for the methylene proton resonance bands of $\text{Fe(}\text{Et}_2\text{dtc)}_2\text{S}_2\text{C}_2\text{(CF}_3)_2\text{)^2}$.

Figure. The PMR spectrum of Co($\text{Et}_2\text{dtc)}_2$ recorded in CDCl$_3$ solution at 30°.

The AB pattern of the N—CH$_2$ proton resonances, shown in the figure, may be satisfactorily explained by assuming that the coupling between the diastereotropic protons (14 Hz) is twice as strong as that with the methyl protons (7 Hz). It is rather peculiar, however, that the diastereotropic nature of the O—CH$_2$ protons did not show in PMR. In this connection it has to be mentioned that for the complexes Co($\text{Et}_2\text{dtc)}_3$, Co($\text{Et}_2\text{dtc)}_3$ and Ni($\text{Et}_2\text{dtc)}_3^+(2)$ in CDCl$_3$ at room temperature only one quartet for the methylene proton resonances is observed, which is due to coupling with the methyl protons. The recording of temperature-dependent PMR spectra will be subject of further investigations. Upon cooling the difference in chemical shift of the diastereotropic dithiocarbamate protons in Co($\text{Et}_2\text{dtc)}_2$ increases slightly. Furthermore one might expect an extra
splitting of the N-bonded ethyl resonances (of the methylene as well as of the methyl proton resonances) due to restricted rotation around the C=N bond (3, 4, 5). At temperatures as low as +9° such a splitting was not observed, probably because at that temperature the hindering of internal rotation is too small. We have indications that the AB pattern is also found for the N—CH₂ dithiocarbamate proton resonances of NBu₄Co(bi)(Bu₂dtc)₂. However, due to the overlap with the absorption bands of the cation (N—CH₂) a well-defined interpretation was not possible so far. Moreover, the interpretation will be difficult in view of the coupling of the N—CH₂ dithiocarbamate protons with α, β and γ protons by which the fine structure is broaded considerably with respect to that of the diethyl-dithiocarbamates.

Infrared spectra

In the IR spectra of the various new Co complexes ν(C=N) is a good measure for the electron donating capacity of the ligands with respect to the dithiocarbamate ligand (table I). It is found that in the mixed biuretato-dithiocarbamate complexes no shift of any importance as compared to that determined for Co(Et₂dtc)₃ can be observed. Therefore, it may be concluded that the electron donating capacity of the biuretato ligand will be roughly equal to that of the Et₂dtc⁻ ligand. For Co(Etxan)(Et₂dtc)₂ the C=N stretching frequency is found at a higher wavenumber than for Co(Et₂dtc)₃. The C—O stretching frequency, which is characteristic of the xantnate ligand is found to be lower for Co(Etxan)(Et₂dtc)₂ than for Co(Etxan)₃. Both
findings can be ascribed to the fact that the electron releasing effect of the \( --\text{NR}_2 \) group is stronger than that of the \( --\text{OR} \) group.

**TABLE I**

<table>
<thead>
<tr>
<th>IR data of some Co(III) complexes.</th>
<th>( \nu(\text{C—N}) )</th>
<th>( \nu(\text{C—O}) ) in cm(^{-1} )</th>
</tr>
</thead>
<tbody>
<tr>
<td>( \text{K}_2\text{Co(bi)}_2(\text{Et}_2\text{dtc}) \cdot 3 \text{H}_2\text{O} )</td>
<td>1495</td>
<td></td>
</tr>
<tr>
<td>( \text{KCo(bi)}(\text{Et}_2\text{dtc})_2 )</td>
<td>1495</td>
<td></td>
</tr>
<tr>
<td>( \text{Co(} \text{Et}_2\text{dtc})_3 )</td>
<td>1490</td>
<td></td>
</tr>
<tr>
<td>( \text{Co(Etxan)}(\text{Et}_2\text{dtc})_2 )</td>
<td>1505</td>
<td>1230</td>
</tr>
<tr>
<td>( \text{Co(Etxan)}_3 )</td>
<td></td>
<td>1250</td>
</tr>
</tbody>
</table>

**TABLE II**

d-\( d \) transitions in Co(III) complexes. The band frequencies and \( \Delta \) are given in \( \text{km} \), \( B \) in \( \text{K} \). The molar extinction coefficient is given in parentheses. Unless otherwise reported, the spectra were recorded in CHCl\(_3\) solution.

| \( \text{K}_2\text{Co(bi)}_2(\text{Et}_2\text{dtc}) \cdot 3 \text{H}_2\text{O} \) | 15.9 | |
| \( \text{NBu}_4\text{Co(bi)}(\text{Bu}_2\text{dtc})_2 \) | 15.0 (500) | 20.8 ? |
| \( \text{Co(bi)}_2(\text{Etxan})^- \) | 17.5 | |

<table>
<thead>
<tr>
<th>( \nu_1 )</th>
<th>( \nu_2 )</th>
<th>( \Delta )</th>
<th>( B )</th>
</tr>
</thead>
<tbody>
<tr>
<td>( \text{Co(} \text{Et}_2\text{dtc})_3 ) (7)</td>
<td>15.5 (600)</td>
<td>20.8 sh</td>
<td>16.3</td>
</tr>
<tr>
<td>( \text{Co(Etxan)}(\text{Et}_2\text{dtc})_2 )</td>
<td>15.7 (470)</td>
<td>21.1 sh</td>
<td>16.5</td>
</tr>
<tr>
<td>( \text{Co(Etxan)}_3 )</td>
<td>16.0 (330)</td>
<td>20.7 sh</td>
<td>16.8</td>
</tr>
</tbody>
</table>

\( \text{sh} = \text{shoulder}; \ * \ \text{reflectance spectrum (in BaSO}_4); \ * * \ \text{recorded in aqueous solution.} \)
UV-VIS spectra

For the series of complexes Co(Et$_2$dtc)$_3$, Co(Etxan)$_3$ and Co(Etxan)(Et$_2$dtc)$_2$ the crystal field splitting energy $\Delta$ and the Racah parameter $B$ were calculated from UV-VIS spectral data (according to Jørgensen (6)) assuming a low spin d$^6$ Co$^{3+}$ ion with a local $O_h$ symmetry. The tabulated values of $\Delta$ and $B$ of Co(etxan)(Et$_2$dtc)$_2$ show a good correspondence with those of Co(Et$_2$dtc)$_3$ and Co(Etxan)$_3$ (table II). Unfortunately, only one d-d transition is observed in the spectra of the Co(bi)$_2$L$^{-2}$ and Co(bi)L$^{-}$ complex ions (table II), so no calculations could be performed.

EXPERIMENTAL

The PMR spectra were recorded with a Varian HA-100 spectrometer operating at 100 MHz. As external reference HMDS was used. The IR data were obtained using a Perkin-Elmer 157 spectrophotometer. UV-VIS spectra were measured with the Cary 14 spectrometer. Magnetic measurements were performed according to the Gouy method.

For preparation of KCo(bi)$_2$ (8), Co(Et$_2$dtc)$_3$ (9) and Co(Etxan)$_3$ (10), the procedures given in the literature were used.

K$_2$Co(bi)$_2$(Et)$_2$dtc). 3 H$_2$O.

To a stirred suspension of KCo(bi)$_2$ (1 mole) in an ethanol-water mixture (1:1) a solution of K(Et$_2$dtc) (1 mole) in the same solvent was added. The reaction temperature had to be kept at 0$^\circ$. After half an hour the solution is filtered and on cooling to -5$^\circ$ blue-violet crystals slowly separated. Upon heating the product decomposed near 130$^\circ$.

Found: Co 11.0; C 19.4; H 3.9; N 17.4; S 11.6. Calcd. for K$_2$Co$_9$H$_{22}$O$_2$N$_7$S$_2$ (541.6): Co 10.9; O 20.0; H 4.1; N 18.1; S 11.8.

NBu$_4$Co(bi)(Bu$_2$dtc)$_2$. 
Two moles of Na(Bu₂dtc) were added to a suspension of KCo(bi)₂ (1 mole) in water. After the color of the solution had changed from blue-violet to green, an excess of NB₄Br was added resulting in the precipitation of a green solid substance. This precipitate was filtered off and purified by dissolving in CHCl₃ and reprecipitation with petroleum ether (40-60°C). The crystals thus obtained melted at 136-140°C.

Found: Co 7.4; C 50.9; H 9.1; N 10.3. Calc. for CoC₃₆H₇₅N₆O₂S₄ (811.2): Co 7.3; C 53.3; H 10.3.

Co(Etxan)(Et₂dtc)₂.

Two moles of Co(Et₂dtc)₃ and one mole of Co(Etxan)₃ were refluxed for 10 minutes in a chloroform solution. The solvent was evaporated and the separated green crystals melted at 131-132°C.

Found: Co 12.5; C 33.6; H 5.4; N 5.2. Calc. for CoC₁₃H₂₅N₂O₆S₆ (462.2): Co 12.7; C 33.8; H 5.5; N 6.1.

ACKNOWLEDGEMENT

The author wishes to thank Mr. P.W.A. van Poppel for his contribution in the experimental work and Mr. J.J. Bour for his practical advices. The investigations were supported in part by the Netherlands Foundation for Chemical Research (SON) with financial aid from the Netherlands Organization for the Advancement of Pure Research.

REFERENCES

