Hydrogenation of Cyclohexene, Catalyzed by CoH₃(PPh₃)₃; Kinetics and Possible Mechanism

J. L. HENDRIKSE AND J. W. E. COENEN

Laboratorium voor Fysische Chemie, Katholieke Universiteit Nijmegen, Toernooiveld, Nijmegen, The Netherlands

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At 40°C, the rate of hydrogenation of cyclohexene, catalyzed by CoH₃(PPh₃)₃, is described by Eq. (I).

\[ \frac{d[\text{S}]}{dt} = k[A_0][\text{S}] \frac{[\text{H}_2][\text{P}]}{1 + (k''[\text{H}_2][\text{P}]^n)} \cdot \frac{1}{(1 + k'''[\text{P}])^p} \]  

where \([A_0]\), \([\text{S}]\), \([\text{H}_2]\) and \([\text{P}]\) are the concentrations of catalyst, cyclohexene, hydrogen and triphenylphosphine, respectively, in solution. Equation (I) shows that triphenylphosphine is an inhibitor for the hydrogenation reaction.

Infrared measurements indicate that PPh₃ does not retard the hydrogen-deuterium exchange between the catalyst and the gas phase. The conclusion is that PPh₃ inhibits the addition of cyclohexene by displacing equilibrium (II),

\[ \text{CoH}_3(\text{PPh}_3)_3 \rightleftharpoons \text{CoH}_3(\text{PPh}_3)_2 + \text{PPh}_3 \]  

The structure of CoH₃(PPh₃)₃ is most likely to be octahedral with two hydride ligands in trans positions (11). This structure is supported by that of the analogous complex IrH₃(PPh₃)₃ (12), which exists in two isomeric configurations. The cis-form (one Co–H stretching vibration at 2075 cm⁻¹ in the ir spectrum) and the trans-form (two Co–H stretching modes at 2130 and 1750 cm⁻¹). It has already been shown that CoH₃(PPh₃)₃ is able to exchange hydride ligands with other molecules (e.g., N₂) (6, 7). Therefore, the main question is whether the hydrogenation of cyclohexene with CoH₃(PPh₃)₃ as a homogeneous catalyst and we give a possible mechanism. It has already been reported that CoH₃(PPh₃)₃ is a homogeneous catalyst for the hydrogenation of cyclohexene (4) and ethylene (5). However, these papers contain few kinetic data and do not deal with extensive kinetic studies of the hydrogenation of olefins with CoH₃(PPh₃)₃.

The best investigated triphenylphosphine-transition metal hydrogenation catalyst to date is RhCl(PPh₃)₃ (1, 2). Essential for this catalyst is the activation of one hydrogen molecule into two hydride ligands. The active species is considered to be the octahedral complex RhH₂Cl(PPh₃)₂S₀, where S₀ is a solvent molecule, which temporarily occupies a phosphine site. Another type of homogeneous hydrogenation catalyst is IrH(CO)(PPh₃)₂ (3). The active species of the compound, IrH(CO)(PPh₃)₂ S₀, is in equilibrium with IrH₃(CO)(PPh₃)₂, so that the rate of hydrogenation decreases with increasing hydrogen pressure and triphenylphosphine concentration. Consequently, on this catalyst, the activation of olefins occurs on a vacant phosphine site in competition with hydrogen and triphenylphosphine.

In this paper we report on the kinetics of the hydrogenation of cyclohexene with CoH₃(PPh₃)₃ as a homogeneous catalyst and we give a possible mechanism.
of our investigation has been to enquire which ligand dissociates from the coordinatively saturated complex in order to create a site for cyclohexene activation. We may ask whether it is a hydride ligand, in view of the easy exchange with N₂, CO, etc., or a phosphine ligand, by analogy with the above-mentioned hydrogenation catalysts.

**Experimental Methods**

**Materials.** Commercial benzene, ether and cyclohexene (Merck AG) were distilled off from a sodium–lead alloy, in order to remove nitrogen, oxygen, water and, in the case of cyclohexene, hydroquinone. The hydrogen was purified by diffusion through Pd. Deuterium (Noury-Baker chemicals), cobalt(III)acetylacetonate (Merck AG) and triphenylphosphine (Merck AG) were used without purification. Tri-n-butylphosphine (Merck AG) was degassed before use.

The catalyst was prepared under hydrogen pressure, by reducing 6 g of cobalt(III) acetylacetonate and 20 g of triphenylphosphine in 150 ml dry, oxygen-free diethyl ether, with 7 ml tri-i-butylaluminum. Before the addition of the tri-i-butylaluminum the mixture had to be cooled in an ice bath. The reaction solution was allowed to stand overnight, warming up slowly from 0°C to room temperature. Yellow crystals of CoH₃(PPh₃)₃ were formed and filtered off. The collected crystals were washed with cold diethyl ether, dried and dissolved in freshly distilled, dry and nitrogen-free benzene. This catalyst solution was stored in a buret under hydrogen pressure.

**Apparatus.** The hydrogenation reactions were carried out in a glass reactor. All connections to the reactor, which might contact the used chemicals or its vapors, were of the greaseless type. The storage burets for the catalyst solution, benzene and cyclohexene were equipped with Teflon membrane valves and connected to the reactor by means of greaseless ball and socket joints. All parts of the apparatus were carefully kept at constant temperature, within 0.1°C. The entire apparatus was leakproof. The regulating manometer, an oil-filled U-tube, was equipped with a photo electric sensor, which opened the bottom valve of the oil storage vessel as soon as the pressure in the reactor system fell below the reference value, so that the reaction was carried out under constant pressure. The reference pressure of the oil manometer was adjusted by bringing the reference vessel on the desired pressure. This pressure could be read on the mercury manometer. A scheme of the hydrogenation apparatus is given in Fig. 1.

The ir measurements were carried out in a similar apparatus to the kinetic experiments. A loop containing a pump and an ir cell was placed on the reactor, such that

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**Fig. 1. Hydrogenation apparatus.**
the reactor contents were pumped through the cell continuously. So, at any desired moment, a spectrum of the reaction mixture could be taken. All parts of the pump which might contact the reaction liquid were constructed from Teflon and all necessary metal parts in the ir cell were made of stainless steel, in order to avoid corrosion by the reaction liquid.

**Experimental procedure.** The desired amount of catalyst solution (10 ml) was introduced into the reactor from a storage buret through a stainless steel connection tube (i.d. $\frac{1}{8}$ in.). This catalyst solution was diluted to the desired concentration with distilled benzene which had been stored under hydrogen. The mixture was allowed to warm up in the reactor during 15 min under stirring. Finally, the cyclohexene was added, and that was considered to be the starting point of the hydrogenation reaction. The progress of the reaction could be read on the gas buret by following the hydrogen volume decreasing at constant pressure. The reproducibility of the kinetic measurements was better than ±5%.

**Results**

The effect of the variables (catalyst concentration, cyclohexene concentration, hydrogen pressure and triphenylphosphine concentration) on the rate of hydrogenation has been measured. In the following figures, the rate of hydrogenation has been plotted as a function of these variables. The rate of hydrogenation was obtained from the tangent in the plot of hydrogen uptake against time at zero time. We have made measurements at only one temperature (40 ± 0.1°C).

Figure 2 shows that the rate of hydrogenation is proportional to the catalyst concentration. The correlation of the rate of hydrogenation with the cyclohexene concentration is also linear to a fair approximation, but extrapolation to zero cyclohexene concentration gives an almost negligible positive intercept with the rate axis (see Fig. 3).

The dependence of the rate on the hydrogen pressure shows a strong saturation effect (Fig. 4). This dependence is described (see Fig. 5) by:

$$\frac{d[S]}{dt} = \frac{P_{\text{H}_2}^{3/2}}{c_4 + c_6 P_{\text{H}_2}^{3/2}}$$

Added triphenylphosphine has a strong retardation effect on the hydrogenation. The dependence of the rate of hydrogenation on the added triphenylphosphine concentration is described by (Fig. 6):

$$-\frac{d[S]}{dt} = \frac{P_{\text{H}_2}^{3/2}}{c_4 + c_6 P_{\text{H}_2}^{3/2}}$$

**Fig. 2.** Rate of hydrogenation as a function of the catalyst concentration. $[S] = 1.65$ mole/liter; $P_{\text{H}_2} = 896$ mm Hg; $T = 40^\circ$C; volume of reaction mixture, 50 ml.

**Fig. 3.** Rate of hydrogenation as a function of the cyclohexene concentration. $[A_0] = 8.3 \times 10^{-2}$ mole/liter; $P_{\text{H}_2} = 896$ mm Hg; $T = 40^\circ$C; volume of reaction mixture, 60 ml.
Combining the results of the measured separate influences of the variables on the rate of hydrogenation, the following expression results:

\[- \frac{d[S]}{dt} = \frac{1}{c_{6}[P] + c_{5}}\]

which may be rewritten as:

\[- \frac{1}{[A_{0}]} \frac{d[S]}{dt} = \frac{(c_{3}c_{6}/c_{4}c_{5})[S][H_{2}]^{3/2}}{1 + (c_{3}c_{6}/c_{4}c_{5})[H_{2}]^{3/2}[P] + c_{6}/c_{5}[H_{2}]^{3/2} + c_{6}/c_{5}[P]}\]  

The \(\text{ir}\) spectra have also been measured. The spectrum of \(\text{CoH}_{3}(\text{PPh}_{3})_{3}\) shows two bands, due to \(\text{Co-H}\) stretching vibrations. In a solution of benzene, we measured the positions of these bands at 1930 and 1750 cm\(^{-1}\). The bending mode was located at 760 cm\(^{-1}\). After exposure to deuterium, all three bands disappeared and only two new bands appeared, at 1265 and 565 cm\(^{-1}\), respectively. The spectra were recorded under reaction conditions, by circulating the reaction liquid through the cell continuously.

When using \(\text{PBu}_{3}\), a simultaneous shift of both \(\text{Co-H}\) stretching modes was observed. With a sixfold excess of \(\text{PBu}_{3}\) with respect to \(\text{CoH}_{3}(\text{PPh}_{3})_{3}\), the positions of these bands were 1920 and 1720 cm\(^{-1}\), which are the band positions of the \(\text{Co-H}\) stretching modes in \(\text{CoH}_{3}(\text{PBu}_{3})_{3}\).

**DISCUSSION**

In view of the observed inhibition of the hydrogenation by free triphenylphosphine, the first step in the mechanism is most

\[\text{Fig. 4. Rate of hydrogenation as a function of the hydrogen pressure. } [A_{0}] = 8.3 \times 10^{-3} \text{ mole/liter; } [S] = 4.12 \text{ mole/liter; } T = 40^\circ\text{C; volume of reaction mixture, 60 ml.}\]

\[\text{Fig. 5. Plot of the reciprocal rate of hydrogenation against } (1/P[H_{2}])^{3/2}.\]

The spectrum of the solvent was subtracted by placing a variable \(\text{ir}\) cell containing benzene in the reference beam of the spectrophotometer. The observed change from hydride to deuteride ligands was very rapid. The catalyst was deuterated within 3 min, so that a dynamic equilibrium must exist between the hydrides in the complex and the hydrogen in the gas phase. No retardation of this \(\text{H}_{2}-\text{D}_{2}\) exchange was observed after addition of free triphenylphosphine or tri-\(n\)-butylphosphine, although these latter compounds are effective inhibitors for the catalytic hydrogenation of cyclohexene.
likely to be the dissociation of one PPh₃ ligand. The kinetic data suggest that the addition of cyclohexene to the catalyst is the rate-limiting step under the experimental conditions of this investigation. Therefore, the following mechanism is proposed.

1. $A + \frac{1}{2}H_2 \rightarrow AH_3$,  \( A = \text{Co(PPh}_3\text{)}_2 \) or its solvated species.

2. $AH_3 + P \rightarrow AP$,

3. $AH_3 + S \rightarrow AH_3S$,

which gives expression (2):

$$\frac{1}{[A_0]} \frac{d[S]}{dt} = \frac{k_2[H_2]^{3/2}[S]}{1 + K_1[H_2]^{3/2} + K_2K_3[H_2]^{3/2}[P] + K_4[P]}.$$  \( \text{(2)} \)

Equation (2) is identical in form to the experimental expression (1). Comparison of Eqs. (1) and (2) gives the result that $K_2$ and $K_3$ are equal. This means that the addition of free PPh₃ to the complex is independent of the presence of hydride ligands.

In view of the fact that all three hydride ligands are involved in the mechanism, the nature of $A$ has to be either Co(PPh₃)₂(I), or Co(PPh₃)₃(II), or the solvated species of I or II. Co(PPh₃)₂ is less favorable, because this would imply that cyclohexene is activated on a coordinatively saturated compound. So we conclude that $A$ has to be Co(PPh₃)₂, or, more likely, its solvated species. Although the experimental Eq. (1) gives a fair description of the rate of hydrogenation as a function of the added PPh₃ concentration, the phosphine contribution from the dissociated complex has been neglected. With small additions of PPh₃ this effect will be significant. Allowance for the dissociated PPh₃ can be made as follows:

A balance over the PPh₃ concentrations gives:

$$[P] = [P_a] + [A_0] \frac{1 + K_1[H_2]^{3/2}}{1 + K_1[H_2]^{3/2} + (K_1K_2[H_2]^{3/2} + K_4)[P]}.$$  \( [P_a] = \text{added } \text{PPh}_3 \).  \( \text{(3)} \)

If we assume that also in this case $K_2$ and $K_4$ are not very different and therefore suppose them to be equal, Eq. (3) simplifies to:

$$[P] = \frac{1}{2}([P_a] - 1/K_4) + \frac{1}{2}([P_a] - 1/K_4)^2 + 4/K_4([P_a] + [A_0])^{1/2}.$$  \( \text{(4)} \)

Substitution of Eq. (4) into Eq. (2) gives a very good approximation of the rate of hydrogenation on the added amount of PPh₃. From the slope and intercept in Fig. 5 the numerical value for $K_4$ has been calculated. The required solubilities of hydrogen in benzene were taken from the data reported by Cook, Hanson and Alder (8). The value for $K_4$ has been calculated from Eqs. (9) and (4) with the aid of an iterative computer program. The plot of the rate of hydrogenation against the total triphenylphosphine concentration is given in Fig. 7.
The resulting values for $K_1$, $K_2$, $K_3$ and $k_r$ are shown in Table 1.

Although Eq. (2) describes our kinetic results fairly well, the proposed mechanism is not fully complete, because it does not include a catalytic cycle. A better mechanism is:

\[
\begin{align*}
\text{TABLE 1} \\
\begin{array}{|c|c|c|}
\hline
\text{Constant} & \text{Unit} & \text{Numerical value} \\
\hline
K_1 & \text{liters/mole} & 140 \\
K_2 & \text{liters/mole} & 2260 \\
K_3 & \text{liters/mole} & 2260 \\
k_r & \text{liters/mole sec} & 0.031 \\
\hline
\end{array}
\end{align*}
\]

If $K_4 \gg K_4[H_2]^{1/2}$, Eq. (5) simplifies to (2), which is the rate expression of the framed part of the mechanism.

Computer calculations show that this simplification is not fully allowed, because fits of $- [A_0]/(d[S]/dt)$ vs $(1/P_{H_2})^n$ give the best linear plot for $n = 1.4$. However, our experimental data to date are not sufficient for solving this problem in detail. In a later paper we shall give a more detailed mechanism, based on more kinetic data. Also the mechanism of insertion of the activated cyclohexene in the Co–H bond is not clear, because our experiments do not give information about it.

The ir spectra confirm the results obtained from the kinetic measurements. When adding free PBu₃ to a solution of CoH₃(PPh₃)₃, a shift of both bands due to Co–H stretching modes was seen. This shift
cess of $\text{PBu}_3$ with respect to $\text{CoH}_3(\text{PPh}_3)_3$. Then the band positions were 1920 and 1720 cm$^{-1}$, which are the locations of the Co–H stretching modes in $\text{CoH}_3(\text{PBu}_3)_3$ (9). A similar exchange of the phosphine ligands has been published by Yamamoto et al. (10) for the closely related complex $\text{CoHN}_2(\text{PPh}_3)_3$. The observed exchange supports the assumption that there is a dynamic equilibrium between the phosphine ligands and the free phosphine in the solution. On the other hand, the $\text{H}_2$–$\text{D}_2$ exchange was not inhibited by a tenfold excess of free phosphine. So the conclusion is that free phosphine inhibits the addition of cyclohexene to the catalyst, but not the activation of hydrogen to hydride. This then confirms the assumption that cyclohexene is activated on a phosphine site, which is probably temporarily occupied by a benzene molecule, while the sites occupied by hydride ligands cannot be occupied by cyclohexene or phosphine.

In the proposed mechanism the essential step is the dissociation of a complex molecule into the active form of the catalyst and one $\text{PPh}_3$ ligand. This phenomenon has also been observed with other homogeneous hydrogenation catalysts [e.g., $\text{RhCl}(\text{PPh}_3)_2$ and $\text{IrH}($CO$)(\text{PPh}_3)_2$]. A discrepancy is observed with the hydrogen pressure dependence of the rate of hydrogenation for $\text{CoH}_3(\text{PPh}_3)_3$ and $\text{RhCl}(\text{PPh}_3)_2$ and $\text{IrH}($CO$)(\text{PPh}_3)_2$. The participation of three hydride ligands, following from our kinetic data, has not been reported before.

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