PDF hosted at the Radboud Repository of the Radboud University
Nijmegen

The following full text is a publisher's version.

For additional information about this publication click this link.
http://hdl.handle.net/2066/142319

Please be advised that this information was generated on 2020-01-29 and may be subject to change.
VINYL ORTHOFORMATES AND VINYL ACETALS. PART III
The reactions of alkenyl acetals and 1-alkenyloxyalkyl carboxylates
with carboxylic acids

BY
J. E. W. VAN MELICK, T. POSNO,
J. W. SCHEEREN and R. J. F. NIVARD

(Department of Organic Chemistry, Catholic University, Toernooiveld, Nijmegen,
The Netherlands)

Treatment of alkenyl acetals IV with one equivalent of a carboxylic acid leads
to the formation of either addition products, 1-acyloxyalkyl alkenyl acetals VI,
or substitution products, 1-alkenyloxyalkyl carboxylates V, dependent upon
the structure of the starting compounds. In a similar way the 1-alkenyloxyalkyl
carboxylates V give 1,1'-diacyloxy ethers X, or acylals XI, with a second equivalent
of a carboxylic acid. Synthetic applications and tentative reaction schemes
of these conversions are given.

In previous papers of this series1,2 we described the reactions of alkenyl
orthoesters I with carboxylic acids. It was shown that treatment of
orthoesters with one equivalent of an acid leads to the formation of
substitution products II via an intermediate carboxonium ion ac­
cording to the overall equation [1].

\[
\begin{align*}
R_3-C(OCR_1=CHR_2)_3 + RCOOH & \rightarrow R_3-C(O)CR_1=CHR_2OCOR + R'-CO-CH_2R_2 \\
(\text{I}) & \quad (\text{II})
\end{align*}
\]

Addition of a second equivalent of the acid did not lead to further
substitution. Instead, the 1,1-dialkenyloxyalkyl carboxylates II decomposed2,3. Dependent on the nature of the substituents R, R1, R2, R3,
either alkenyl carboxylates (R3-CO-OCR1=CHR2) and acylals

   (1971).
   (1973).
[R^{1–C(OCOR)}_2CH_2R^2] or alkenyl carboxylates, acid anhydrides \((\text{RCO}_2\text{O})\) and carbonyl compounds \((R^{1–CO–CH_2R^2})\) were obtained as the main products. It was suggested that a trioxenium cation III participates in the fragmentations.

![Chemical structure](III)

In this paper our investigation of the reactions between the corresponding alkenyl acetals IV and carboxylic acids are presented.

The reaction: \(R^2–CH(OCR^{1}=CH_2)_2 + R\text{COOH}\)

The addition of one equivalent of an acid to alkenyl acetals did not lead always to the expected substitution products. In general, these hitherto unknown 1-alkenyloxyalkyl carboxylates \(V^{4}\) were only obtained if \(R^2\) was a phenyl residue (preferably with an electron-donating substituent), or if \(R^1\) contained an electron-withdrawing group (e.g. \(R^1=\text{CH}_2\text{OCH}_3\)). In other cases addition products, 1-acyloxyalkyl alkenyl acetals VI, always accompanied by small amounts of di-addition products VII, arose. Table I summarises the reaction products obtained from several alkenyl acetals. The results can best be explained by the reaction scheme [2].

---

4 Some products \(V\) of the type \(R^{1–CH(OCOR)OCH=CHR^2}\) (\(R^1\) and \(R^2\) are aliphatic) are described by L. W. McTeer in a U.S. Patent, 3,383,374 (1968). See also C.A. 69, 26768 (1968).
Table I

Reaction products from $R^2-\text{CH(OCR}^1\text{)}=\text{CH}_2 + \text{RCOOH}$

<table>
<thead>
<tr>
<th>$R^1$</th>
<th>$R^2$</th>
<th>main product</th>
<th>minor products</th>
</tr>
</thead>
<tbody>
<tr>
<td>H</td>
<td>H</td>
<td>VI</td>
<td>VII</td>
</tr>
<tr>
<td>H</td>
<td>CH$_3$, $n$-C$_3$H$_7$, 2,6-Cl$_2$-C$_6$H$_3$</td>
<td>VI</td>
<td>VII, V</td>
</tr>
<tr>
<td>H</td>
<td>$p$-X-C$_6$H$_4$*</td>
<td>V</td>
<td>-</td>
</tr>
<tr>
<td>CH$_3$</td>
<td>CH$_3$Cl, CH$_3$*</td>
<td>VI</td>
<td>VII, V</td>
</tr>
<tr>
<td>CH$_3$</td>
<td>$p$-X-C$_6$H$_4$*</td>
<td>V</td>
<td>-</td>
</tr>
<tr>
<td>CH$_2$OCH$_3$</td>
<td>CH$_3$Cl</td>
<td>V</td>
<td>VI, VII</td>
</tr>
<tr>
<td>CH$_2$OCH$_3$</td>
<td>$CH_3$, $n$-C$_2$H$_7$, $p$-X-C$_6$H$_4$*</td>
<td>V</td>
<td>-</td>
</tr>
</tbody>
</table>

* $X = H$, Cl, CH$_3$ or OCH$_3$.

** After prolonged reaction time $V$ becomes the main product (see text).

The scheme suggests that the carboxonium ion VIII, formed in an initial protonating step, does not always produce a carboxonium...
ion IX. This ion IX is less stable than the corresponding cation \((H_2C = CR^1-O\longrightarrow CR^2=O-CR^1=CH_2)\), formed in the analogous reactions of alkenyl orthoesters with acids. IX arises only if strongly stabilizing, donating groups \(R^2\) are present, or if electron-withdrawing substituents in \(R^1\) make its precursor VIII less stable.

The addition of a carboxylic acid to VI proceeds at a rate comparable to the formation of VI, because VII was always found as a second product in reactions where VI was the main product. This can be understood because substituent influences on the reactivity of the double bond in IV and VI will be about the same. A similar addition of acid to the substitution product V was generally not observed under the applied reaction conditions, using only one equivalent of the acid (see, however, the following section).

The formation of VI from VIII is probably a reversible process. In one case \([\text{CH}_3-\text{CH(OCCCH}_3 = \text{CH}_2)_2]\) the amount of the initial addition product decreased gradually after prolonged reaction times, and the substitution product (V, \(R^1 = R^2 = \text{CH}_3\)) became the main product. The conversion of VI into V could have occurred directly via a cyclic transition state, but in the presence of acid a pathway via the cations VIII and IX seems more probable.

The reactions between alkenyl acetals and carboxylic acids have been followed by NMR to get an idea of the influence of \(R, R^1\) and \(R^2\) on the rate of product formation. It appeared that the rate decreased in the following sequences:

\[
\begin{align*}
R & : \text{CH}_3\text{Cl} > \text{H} > \text{CH}_3
R^1 : \text{CH}_3 > \text{H} > \text{CH}_2\text{OC}_2\text{H}_5 > \text{CH}_2\text{OCH}_3 > \text{CH}_3\text{Cl}
R^2 : p-\text{CH}_3\text{OC}_6\text{H}_4 > p-\text{CH}_3\text{C}_6\text{H}_4 > \text{C}_6\text{H}_5 > p-\text{ClC}_6\text{H}_4 > n-\text{C}_3\text{H}_7, \text{C}_2\text{H}_5, \\
& \quad \text{CH}_3 > \text{H} > \text{CH}_2\text{Cl}, 2,6-\text{Cl}_2\text{C}_6\text{H}_3
\end{align*}
\]

These data are in accordance with the proposed reaction scheme, in so far as both supposed intermediate cations are better stabilized by stronger electron-donating groups in \(R^1\) or \(R^2\), and protonation, probably involved in the rate-determining step, becomes faster with stronger acids.

The reaction: \(R^2-\text{CH(OCR}^1=\text{CH}_2)\text{OCOR} + \text{RCOOH}\)

The 1-alkenyloxalkyl carboxylates V behave similarly towards carboxylic acids as the alkenyl acetals, although their reactivity is lower.
Dependent upon the nature of the substituents (R, R¹, R²) and also upon temperature, either addition products, viz. 1,1'-diacyloxy-ethers X, or substitution products, viz. acylals XI, were obtained. These results suggest that the reactions proceed according to scheme [3], in outline similar to the scheme [2] given above. Slight modifications will be explained in the subsequent discussion.

At temperatures not exceeding 50°, addition, leading to compounds X, appeared to be the usual route in this reaction. Acylals XI were only formed in substantial amounts under these conditions if R² was an aryl group, strongly stabilizing a carboxonium ion such as XIII, and at the same time R¹ was an electron-withdrawing group (CH₂OCH₃, CH₂Cl). However, in every case the amount of XI increased at higher temperatures (80-100°) at the cost of previously formed X.

It was shown by means of NMR studies that pure 1,1'-diacyloxy ethers X decompose at higher temperatures into the acid (RCOOH) and the parent compound V, especially when R¹ ≠ H. Subsequently acylals XI and ketones XIV appear in the reaction mixtures. On prolonged
heating the acylals themselves gradually decompose\(^5\), giving rise to aldehydes XV and acid anhydrides XVI.

The influence of substituents R, R\(^1\) and R\(^2\) on the rate of thermal decomposition of X differs from that on its formation-rate. In both cases R\(^2\) has no significant influence, and both reactions slow down in the series \( R = \text{CH}_2\text{Cl} > \text{H} > \text{CH}_3 \). For the influence of R\(^1\), however, we found

\[
\begin{align*}
\text{CH}_3 & > \text{CH}_2\text{OCH}_3, \text{CH}_2\text{OC}_2\text{H}_5, \text{CH}_2\text{Cl} \geq \text{H} \text{ in the decomposition of X, but } \\
\text{CH}_3 & > \text{H} > \text{CH}_2\text{OCH}_3 \geq \text{CH}_2\text{OCH}_3 \geq \text{CH}_2\text{Cl} \text{ in the formation of X from V.}
\end{align*}
\]

A possible explanation of these data is that, at higher temperatures and at the low acid concentrations in the initial stage of the reaction, the decomposition of X to V proceeds \textit{via} a pathway comparable with that in ester pyrolyses\(^6\) rather than \textit{via} a reversal of its synthetic route. The rate of this way of elimination will decrease as the olefins formed become less stable (R\(^1\) = CH\(_3\)>CH\(_2\)OCH\(_3\)>H). Similar eliminations in compounds VI (Scheme 2) have not been observed, but cannot be excluded under appropriate conditions.

The relatively slow decomposition of 1,1'-diacyloxy ethers X, with R\(^1\) = H, explains why the elimination of acid in the example Xa represented in scheme [4] is restricted to that side of the molecule bearing the CH\(_2\)OCH\(_3\) group (the scheme represents only overall reactions). Two products XVII and XVIII appeared to be formed, the latter in \textit{cis} and \textit{trans} configuration. The presumably less stable product XVII

\[
\begin{align*}
\text{CH}_3 & = \text{CH}-\text{O}-\text{C}-\text{CH}_3 \\
\text{OCOR} & \quad \text{(Xa)} \\
\text{CH}_3 & \quad \text{H} \quad \text{CH}_2\text{OCH}_3 \
\end{align*}
\]

\[
\begin{align*}
\text{H} & \quad \text{CH}_2\text{OCH}_3 \\
\text{CH}_3 & \quad \text{C}-\text{O}-\text{C}-\text{CH}_3 \\ 
\text{OCOR} & \quad \text{(XVII)} \\
\text{H} & \quad \text{CHOCH}_3 \\
\text{CH}_3 & \text{C}-\text{O}-\text{C}-\text{CH}_3 \\
\text{OCOR} & \quad \text{(XVIII; cis and trans)}
\end{align*}
\]


arose faster, but was also converted more rapidly into an acylal or back into Xa. On careful decomposition more XVIII than XVII was eventually obtained, due to the "reversibility" of the reaction. However, XVIII also undergoes substitution by the acid present, giving mainly the ketone, \( CH_3-CO-CH_2OCH_3 \), and an acylal, \( CH_3-CH(OCOR)_2 \).

As to the intermediate carboxonium ions in scheme [3], it might be expected that XII and XIII will be less stable than the corresponding cations VIII and IX in scheme [2], respectively. The lower reactivity of the 1-alkenyloxyalkyl carboxylates V in comparison with alkenyl acetals IV seems in accordance with this expectation. The ion XII might be stabilized by isomerisation, giving a trioxenium ion XIIa, comparable to that (III) suggested in reactions of alkenyl orthoesters with two equivalents of an acid\(^2\).

However, the ion XIIa might also equilibrate with the carboxonium ion XIIb. This should lead to the same addition product X as obtained via XII, but by substitution other products, \( R^1-C(OCOR)_2CH_3 \) and \( R^2CHO \), should arise. These have never been found in our reaction mixtures. Hence, the occurrence of a trioxenium ion as an intermediate in this reaction remains uncertain.

A trioxenium ion\(^7\) such as XIIa can surely be expected as an intermediate in the reaction of methylenetrioxanes with acids. We found that treatment of 2,4-dimethyl-6-methylene-trioxane (XIX) with acetic acid at room temperature gives quantitatively di(l-acetoxy-ethyl) ether (XX) in a strongly exothermic reaction. NMR studies using 20% solutions in CDCl\(_3\) showed that at low temperature (below \(-60^\circ\)) the addition product XXI, which is only stable below \(-40^\circ\) (Scheme 5), was formed quantitatively.

Finally, a brief remark may be made concerning the incidental occurrence of other side-products in reactions of 1-alkenylxoyalkyl carboxylates with acids. If $R^1 = H$ and $R^2 = p-X-C_6H_4$ ($X = H, CH_3, OCH_3$) in the parent compound, a cinnamic aldehyde XXIII also appeared in the reaction mixture. The amount of this side-product was higher when a stronger acid was used ($R = CH_2Cl$) and when $X$ was more electron donating ($CH_3O > CH_3 > H$). A possible explanation is given in scheme [6]. It was shown that condensation of an aliphatic and aromatic aldehyde, which might also explain formation of XXIII did not occur under the reaction conditions used.
Experimental

In general, isolation and purification of reaction products was performed by distillation with 70 x 1.2, 20 x 1.0 or 20 x 0.8 cm. Vigreux columns or with a spinning band column (Normag) of 25 plates. The use of other techniques is explicitly mentioned where appropriate.

Products were identified by NMR (Varian HA-100 or Varian T-60) from 10% solutions in carbon tetrachloride with TMS as internal standard. A mass spectrum (Varian-Mat-SM1-B) was sometimes also recorded.

In view of the large number of new compounds prepared, only a small number of each type are listed in Table II and III. Boiling points, refractive indices, NMR-data, and yields of preparations for many other examples from the various classes are available.

1) Synthesis of β-chloroalkyl acetals (Table II)

a. β-Chloroalkyl acetals of acetaldehyde, propionaldehyde and butyraldehyde were synthesized from the appropriate aldehydes and a β-chloroalcohol in the presence of granulated calcium chloride.

b. β-Chloroalkyl acetals of benzaldehydes and chloroacetaldehyde were obtained via an acid catalyzed alcohol exchange from appropriate ethyl acetals. The products were isolated after neutralization of the reaction mixture with sodium methoxide.

c. The β-chloroethyl acetal of formaldehyde was obtained by refluxing a mixture of equimolar amounts of paraformaldehyde, β-chloroethanol and β-chloroethyl orthoformate in the presence of a catalytic amount of p-toluenesulfonic acid for 8 hours.

d. 6-Bromomethyl-2,4-dimethyltrioxane was prepared by slow addition of 0.6 moles of bromine to 1 mole of paraldehyde, in the presence of enough potassium carbonate to neutralize the hydrogen bromide formed, and at a temperature below –10°. The mixture was left until it was discoloured and the pH had increased to about 5. It was then filtered, and the filtrate distilled at reduced pressure, b.p. 78°/14 mm, yield 75%, nD 1.4626.

2) Synthesis of alkenyl acetals IV (Table II)

a. Elimination of hydrogen chloride from β-chloroalkyl acetals was performed as described previously for β-chloroalkyl orthoesters. With acetals of formaldehyde and acetaldehyde sodium hydride was used and the alkenyl acetals were directly distilled from the reaction mixtures. With those of propionaldehyde, butyraldehyde, benzaldehydes and chloroacetaldehyde sodium hydride in 1,2-dimethoxyethane, to which some tert-butyl alcohol had been added, was used.

---

9 Ref. 8, p. 250.
10 Ref. 8, p. 222.
11 A. Stepanow, N. Preobraschensky and M. Schtschukina, Ber. 59, 2533 (1926).
### Table II

<table>
<thead>
<tr>
<th>R²</th>
<th>R¹</th>
<th>B.p. °C/mm</th>
<th>n_D</th>
<th>Yield</th>
<th>δ Hₘ</th>
</tr>
</thead>
<tbody>
<tr>
<td>β-Chloroalkyl acetals, R²—CH₄(OCHR¹=CH₂Cl)₂</td>
<td></td>
<td></td>
<td></td>
<td></td>
<td></td>
</tr>
<tr>
<td>H</td>
<td>H</td>
<td>73°/1.7</td>
<td>1.4550</td>
<td>65 %</td>
<td>4.71</td>
</tr>
<tr>
<td>CH₃</td>
<td>CH₃</td>
<td>76°/0.8</td>
<td>1.4492</td>
<td>45 %</td>
<td>4.85</td>
</tr>
<tr>
<td>n-C₅H₇</td>
<td>CH₃Cl</td>
<td>130°/0.11</td>
<td>1.4799</td>
<td>35 %</td>
<td>4.74</td>
</tr>
<tr>
<td>CH₂Cl</td>
<td>CH₃</td>
<td>92°/0.5</td>
<td>1.4706</td>
<td>70 %</td>
<td>4.77</td>
</tr>
<tr>
<td>C₆H₅</td>
<td>CH₂Cl</td>
<td>158°/0.2</td>
<td>1.5322</td>
<td>80 %</td>
<td>5.76</td>
</tr>
<tr>
<td>p-Cl—C₆H₄</td>
<td>CH₃</td>
<td>162°/0.8</td>
<td>1.5188</td>
<td>80 %</td>
<td>5.75*</td>
</tr>
<tr>
<td>2,6-Cl₂—C₆H₃</td>
<td>H</td>
<td>151°/0.1</td>
<td>1.5450</td>
<td>80 %</td>
<td>6.05</td>
</tr>
</tbody>
</table>

| Alkenyl acetals, R²—CH₄(OCHR¹=CH₂)₂ |
| CH₃| CH₃| 48°/0.26 | 1.4261 | 70 % | 5.46 |
| CH₃| CH₃OCH₃| 75°/0.6 | 1.4476 | 90 % | 5.56 |
| CH₃| CH₂Cl| 76°/0.2 | 1.4801 | 40 % | 5.50 |
| CH₂Cl| CH₃ | 70°/1.8 | 1.4480 | 70 % | 5.40 |
| C₆H₅| CH₃OCH₃| 115°/0.3 | 1.5069 | 55 % | 6.25 |
| p-Cl—C₆H₄| CH₃ | 95°/0.9 | 1.5175 | 75 % | 6.12 |
| 2,6-Cl₂—C₆H₃| H  | 93°/0.1 | 1.5463 | 80 % | 6.62 |

* At both sides of the main peak a small peak from diastereoisomers is present.

1,3-Dichloroisopropyl acetals were converted into the corresponding 3-alkoxyisopropenyl acetals with sodium alkoxide in alcohol. With potassium hydroxide in water/1,2-dimethoxyethane (1/1) the same starting compounds gave 3-chloroisopropenyl acetals. In the reaction between 1,3-dichloroisopropyl acetals of a benzaldehyde and sodium methoxide in methanol the dimethyl acetal of the benzaldehyde used was a substantial side-product.

b. 2,4-Dimethyl-6-methylenetrioxane (XIX) was obtained by the addition of 0.1 mole of 6-bromomethyl-2,4-dimethyltrioxane to a suspension of 0.13 moles of sodium hydride in 35 ml hexamethylphosphonamide. The mixture was heated at 80° for four hours and then distilled at reduced pressure. The fraction collected from 45–55° (40 mm) was redistilled; b.p. 52° (48 mm), yield 40 %, n_D 1.4265. The cis-trans mixture was not separated.

### 3) Reactions of alkenyl acetals with carboxylic acid

Syntheses of 1-alkenyoxyalkyl carboxylates V (Table III) and 1-acyloxyalkyl alkenyl acetals VI.
Vinyl orthoformates and vinyl acetals. Part III

Anhydrous formic or acetic acid (0.055 moles), or a solution of 0.055 moles of chloroacetic acid in 10 ml of dry ether, was added dropwise to 0.05 moles of an alkenyl acetal over half an hour. Volatile compounds (ether, acetaldehyde), if present, were evaporated at reduced pressure. In most cases the reaction mixture was then heated at about 50° until the starting compound had for the most part (>90%) disappeared. The reaction times varied from 1 to 36 hours; slow reactions (R = CH₃, R¹ = CH₂OCH₃ or CH₂Cl, Scheme 2) could be catalyzed by addition of a small amount of trifluoroacetic acid. Reactions of compounds with R¹ = CH₃ were performed at 20°. In cases where 1-acyloxyalkyl alkenyl acetals VI were isolated as the main product, small amounts of 1-acyloxyalkyl acetals (VII) were obtained as side products. Physical constants of some isolated mono-addition products VI are given below.

\[
\begin{align*}
\text{CH₂C-CH(OCH=CH₂)OCH(OOCCH₃)CH₃} & \text{ b.p. 81°/14 mm; } n_0° = 1.4376; \text{ yield 50%;} \\
n-C₃H₇-C[H(OCH=CH₂)-OCH(OOC)CH₃] & \text{ b.p. 78°/14 mm; } n_0° = 1.4228; \text{ yield 50%}.
\end{align*}
\]

4) Synthesis of 1-acyloxyethyl acetals of aliphatic aldehydes (VII)

1-Acyloxyethyl acetals can be isolated in good yields from reaction mixtures similar to those used in the foregoing procedure, but containing a vinyl acetal of an aliphatic aldehyde and 2.5 instead of 1.1 equivalent of an acid, e.g.

\[
\begin{align*}
\text{H₂-C[O-CH(OOC)CH₃]₂} & \text{ b.p. 107°/16 mm; } n_0° = 1.4161; \text{ yield 80%;} \\
\text{CH₂C-CH[O-CH(OOC)CH₃]₂} & \text{ b.p. 131°/16 mm; } n_0° = 1.4386; \text{ yield 90%}.
\end{align*}
\]

5) Synthesis of 1,1'-diacyloxy ethers¹² (X) (Table III)

a. Reaction mixtures similar to those used in the reactions between alkenyl acetals and acids were employed, using an alkenyloxalkyl carboxylate instead of an alkenyl acetal. Generally, best results were obtained at low reaction temperatures (0° for R¹ = CH₃, 50° for R¹ = H, room temperature in other cases). Reaction times needed for at least 90% conversion varied considerably; 2-36 hours with mono-chloroacetic acid, 2-72 hours with formic acid, 6 hours – two weeks with acetic acid. In general the fastest conversions were found for R¹ = CH₃ and the slowest for R¹ = CH₂OCH₃ or CH₂Cl. Higher temperatures or catalysis with trifluoroacetic acid have to be avoided, however, in order to prevent formation of acylals. Because the products can decompose during distillation at too high temperature, isolation and purification were performed at low pressure and in some cases by molecular distillation.

b. 1-Acyloxyethyl acetals (VII, R¹ = H) can be converted into 1,1'-diacyloxy ethers (X) by heating them with 0.5 equivalent of the acid corresponding with the acyloxy residue at 100–120° for 2–6 hours.

6) Synthesis of acylals of benzaldehydes (XI)

1.1 Equivalents of a carboxylic acid were added to an 1-alkenyloxibenzyl carboxylate (V, R² = p-XC₆H₄, R¹ ≠ H) or 2.5 equivalents to an alkenyl acetal of benzaldehyde

¹² For a simple synthesis of symmetrical 1,1'-diacyloxy ethers of the type R¹–CH(OCOR)OCH(OCOR)R¹ (R¹ is aliphatic) see ref. 4.
Table III

<table>
<thead>
<tr>
<th>R²</th>
<th>R¹</th>
<th>R</th>
<th>B.p. °C/mm</th>
<th>nD²⁰</th>
<th>yield</th>
<th>δ H₂</th>
</tr>
</thead>
<tbody>
<tr>
<td>CH₃</td>
<td>H</td>
<td>CH₂Cl</td>
<td>43°/0.7</td>
<td>1.4415</td>
<td>10%</td>
<td>6.18</td>
</tr>
<tr>
<td>CH₃</td>
<td>CH₂OCH₃</td>
<td>H</td>
<td>53°/1.0</td>
<td>1.4273</td>
<td>80%</td>
<td>6.40</td>
</tr>
<tr>
<td>CH₃</td>
<td>CH₂Cl</td>
<td>CH₂Cl</td>
<td>74°/0.2</td>
<td>1.4682</td>
<td>70%</td>
<td>6.38</td>
</tr>
<tr>
<td>CH₃Cl</td>
<td></td>
<td>CH₂Cl</td>
<td>91°/3.5</td>
<td>1.4533</td>
<td>70%</td>
<td>6.42</td>
</tr>
<tr>
<td>C₆H₅</td>
<td>H</td>
<td>H</td>
<td>76°/1.0</td>
<td>1.5091</td>
<td>80%</td>
<td>6.90</td>
</tr>
<tr>
<td>C₆H₅</td>
<td>H</td>
<td>CH₂Cl</td>
<td>113°/0.4</td>
<td>1.5154</td>
<td>75%</td>
<td>7.04</td>
</tr>
<tr>
<td>C₆H₅</td>
<td>CH₃</td>
<td>CH₃</td>
<td>102°/0.3</td>
<td>1.5010</td>
<td>80%</td>
<td>7.07</td>
</tr>
<tr>
<td>C₆H₅</td>
<td>CH₂OCH₃</td>
<td>CH₃</td>
<td>H</td>
<td>101°/0.2</td>
<td>1.5183</td>
<td>85%</td>
</tr>
<tr>
<td>2,6-Cl₂-C₆H₅</td>
<td>H</td>
<td>CH₃</td>
<td>93°/0.1</td>
<td>1.5392</td>
<td>5%</td>
<td>7.52</td>
</tr>
</tbody>
</table>

1,1'-diacyloxy ethers, R²—CH₄(OCR¹)=CH₂OOCR

| H | H | H | 80°/15 | 1.4142 | 75%   | 5.40 |
| CH₃ | H |  CH₂Cl | 106°/0.6 | 1.4540 | 75%   | 6.13 |
| CH₃ | CH₃ |  H | 56°/1.5 | 1.4190 | 65%   | 6.41 |
| CH₃ | CH₂OCH₃ |  CH₃ | 81°/1.3 | 1.4246 | 80%   | 6.28 |
| n-C₆H₇ | CH₂OCH₃ |  CH₂Cl | 81°/1.3 | 1.4590 | 50%   | 6.31 |
| CH₃Cl | H |  CH₂Cl | 136°/1.3 | 1.4698 | 70%   | 6.13 |
| CH₃Cl | CH₃ |  CH₃ | 69°/0.1 | 1.4365 | 80%   | 6.28 |
| C₆H₅ | H |  CH₃ | 117°/1.1 | 1.4823 | 60%   | 6.83* |
| C₆H₅ | CH₃ |  CH₃ | 108°/0.1 | 1.4842 | 65%   | 7.09 |
| 2,6-Cl₂-C₆H₅ |  CH₃ |  H | ** | 1.5155 | 50%   | 7.23 |
| 1,1'-diacyloxy ethers (X) |

| * The main peak is accompanied by a small peak due to diastereoisomers. |
| ** Isolated by molecular distillation at ~40°/<10⁻³ mm. |

(IV, R² = p-XC₆H₄, R¹ ≠ H), and the mixture was heated to 50° for about 2 hours for R¹ = CH₃, or about 12 hours for R¹ = CH₂OCH₃; in this way we isolated, e.g.
C₆H₅—CH(OCOH)₂ b.p. 80°/0.1 mm; nD²⁰ 1.5065; yield 70% ;
p-ClC₆H₄—CH(OCOH)₂ b.p. 121°/1.5 mm; nD²⁰ 1.5261; yield 60% .

7) Decomposition of 1,1'-diacyloxy ethers (X)

Pure 1,1'-diacyloxy ethers were heated at such a pressure that the acid eliminated evaporated from the reaction mixture at a suitable rate. Temperatures were 100–120° for
R\(^1\) = CH\(_3\) and about 200° for R\(^1\) = H. Elimination products R\(^2\)CH(OCR\(^1\) = CH\(_2\))OCOR could be isolated by distillation and redistillation, e.g.

- CH\(_2\)Cl—CH(OCCCH\(_3\) = CH\(_2\))OCOOH b.p. 68°/15 mm; \(n^2_0\) 1.4416; yield 80%;
- CH\(_3\)—CH(OCCCH\(_3\) = CH\(_2\))OCOCH\(_3\) b.p. 96°/18 mm; \(n^2_0\) 1.4149; yield 70%;
- H—CH(OCH=CH\(_2\))OCOCH\(_3\) b.p. 75°/100 mm; \(n^2_0\) a.4105; yield 60%.

To obtain elimination product of general structure R\(^2\)—CH(OCCCH\(_3\) = CHOCH\(_3\))OCOR (XVIII) appropriate 1,1'-diacyloxy ethers were heated at slightly lower temperatures (90° for R = H or CH\(_2\)Cl, 100° for R = CH\(_3\)) until the starting compound had disappeared (3–9 hours). By distillation of the residue acylals XI, as well as the alkenyloxy-alkyl carboxylates R\(^2\)CH(OCCCH\(_3\) = CHOCH\(_3\))OCOR, could be isolated. The latter compounds could not be separated completely from traces of acylals XI and 1-alkenyloxy-alkyl carboxylates V, even with a spinning band column, but their purity was at least 90%. Boiling points were about the same as those of their isomers XVII.

**Acknowledgements**

We wish to thank Drs. F. Gerhartl and Mr. H. Mous for recording the mass spectra, Mrs. L. van Herpen for recording NMR spectra and Mr. H. Couvreur for his contribution to a part of this work.

(Received April 5th, 1973)