The following full text is a publisher’s version.

For additional information about this publication click this link.
http://hdl.handle.net/2066/142134

Please be advised that this information was generated on 2019-08-18 and may be subject to change.
CHEMISTRY OF STRAINED POLYCYCLIC COMPOUNDS—VI\textsuperscript{1,2}

STEREOSPECIFIC BASE INDUCED HOMOKETONIZATION OF CUBANE, HOMOCUBANE AND 1,3-BISHOMOCUBANE BRIDGEHEAD ALCOHOLS

A. J. H. KLUNDER and B. ZWANENBURG\textsuperscript{*}

Department of Organic Chemistry, University at Nijmegen, Toernooiveld, Nijmegen, The Netherlands

(Received in the UK 28 October 1972; Accepted for publication 9 January 1973)

Abstract—The base induced homoketonization of bridgehead cubane alcohol\textsuperscript{10}, homocubane alcohol 1 and acetate 7, and 1,3-bishomocubane alcohols 13 and 17 has been studied. Under protic conditions, homocubane alcohol 1 and acetate 7 are converted quantitatively into half cage ketones 3 and 8, respectively, by exclusive cleavage of the C\textsubscript{2}—C\textsubscript{3} bond. Similarly, homoketonization of 1,3-bishomocubane alcohols 13 and 17 leads to half cage ketones 15 and 18, respectively, by exclusive cleavage of the C\textsubscript{2}—C\textsubscript{3} bond. As shown by deuterium labeling experiments homoketonization of 1, 7, 13 and 17 proceeds with high stereospecificity and with retention of configuration (> 96%) at the carbon of substitution. The cubane acetate 10 gave, under similar conditions, complex mixtures of ring-opened products. Under aprotic conditions, base treatment of homocubane alcohol 1 leads to cleavage of the C\textsubscript{2}—C\textsubscript{3} and C\textsubscript{2}—C\textsubscript{5} bond giving the tricyclo[4.2.1.0\texttextsuperscript{2,5}]nonene 21. The mechanism of homoketonization is discussed.

Bridgehead alcohols and amines in highly strained polycyclic systems show a remarkable sensitivity to base as exemplified by the homoketonization reaction of 1-hydroxynortricyclene,\textsuperscript{3} a birdcage bridgehead alcohol,\textsuperscript{4-5} and an anionic fragmentation of a chlorinated birdcage amine.\textsuperscript{6} Apparently, ring strain in these type of compounds has a unique influence on the chemical properties of the hydroxy and amino function.

This paper deals with the homoketonization reaction of some bridgehead cage alcohols which show a diversity in cage strain, viz cubane, homocubane and 1,3-bishomocubane alcohols. The synthesis of these alcohols was described in previous parts of these series.\textsuperscript{2,7}

The homocubane bridgehead alcohol 1 and its acetate 2 were extremely base labile. Upon treatment with NaOMe in MeOH, alcohol 1 or its acetate 2 reacted almost instantaneously, giving half cage ketone 3 in quantitative yield (Scheme 1). To this ketone which was isomeric with alcohol 1, structure 3 was assigned on basis of spectral evidence. The IR spectrum shows a C==O stretching frequency at 1765 cm\textsuperscript{-1}, typical for a cyclobutane none. The NMR spectrum (C\textsubscript{6}D\textsubscript{6}) displays one half of an AB quartet as a doublet of doubletts centered at 8 1·30 ppm\textsuperscript{*} ascribed to inside proton H\textsubscript{n}. This H\textsubscript{n} signal is split by H\textsubscript{g} (J ~ 13 Hz) and by H\textsubscript{e} (J ~ 2 Hz). Coupling with H\textsubscript{g} is negligible probably because the dihedral angle between H\textsubscript{n} and H\textsubscript{g} is about 90°. The lowfield half of the AB pattern for H\textsubscript{n} appears as a broad multiplet centered at 8 2·1 ppm.\textsuperscript{*} This absorption coincides with that of H\textsubscript{g}. The ethylene ketal protons appear as an unsymmetrical multiplet between 8 3·41 and 4·0 ppm, while the remaining protons are found at 8 2·4—3·0 (H\textsubscript{d,e}) and 8 3·10—3·40 (H\textsubscript{f,g}) as a complex pattern. The isomeric structure 4 which can be envisioned by scission of the central C\textsubscript{2}—C\textsubscript{3} bond, must be rejected on basis of the NMR spectrum, since: (i) a doublet of triplets would be expected for the endo proton H\textsubscript{n} in 4 as the result of coupling with H\textsubscript{t} and the equivalent protons H\textsubscript{h} and H\textsubscript{e}, (ii) in the symmetrical ketone 4 the ethylene ketal protons are expected to appear as a symmetrical AA'BB' absorption,\textsuperscript{†} (iii) the upfield shift for H\textsubscript{g} in 3 as compared with 1 is in agreement with relief of strain around C\textsubscript{g} in 3 (in 4 the congestion around C\textsubscript{g} has hardly changed). Further confirmation of structure 3 was obtained from its behaviour during LAH reduction and from the Grignard reaction with MeMgl. As expected for such half-cage structures,\textsuperscript{4} reaction with these reagents proceeds with a high degree of steric approach control to yield exclusively the oxygen-inside alcohols 5 and 6, respectively. Because of the congestion in the half cage
alcohols 5 and 6, the endo proton Hₙ will experience a strong deshielding effect of the OH-function. The spectra of 5 and 6 indeed show the Hₙ signal at a deshielded position, viz as a half of an AB quartet centered at δ 2.64 ppm for 5 and at δ 3.04 ppm for 6. The outside protons Hₓ appear as complex multiplets at δ 1.85-2.35 and 2.05-2.45 ppm, respectively. Unequivocal assignment of the Hₓ and Hₙ signals was accomplished with deuterium labeling (vide infra).

With the structure of 3 resolved, attention was turned to the stereochemistry of the cleavage of the cyclobutane ring. Hence, acetate 2 was treated with NaOMe in MeOD giving mono deuterated ketone 3a in quantitative yield. The NMR spectrum revealed that deuterium was introduced exclusively (> 96%) at the C₇ endo position because of the absence of the AB pattern at δ 1.30 and a simplified two proton absorption for Hₓ and Hₙ at δ 2.1 ppm (the signals of Hₓ, Hₙ, Hₐ, H₅, H₆ and the ketal protons remained unchanged). Under the same conditions, treatment of ketone 3 with NaOMe in MeOD did not lead to any H/D exchange. Evidently, homoketonization of 1 is a highly stereospecific process and proceeds with retention (> 96%) of configuration at C₇.

From inspection of molecular models, it became apparent that the bulky ketal group could seriously hamper the approach of the proton donating solvent molecule, i.e. methanol, from the exo-side and therefore would promote endo-attack. For this reason the homoketonization of deketalized acetate 7 was investigated. Treatment of 7 with NaOMe in MeOH at room temperature gave quantitatively the half cage ketone 8. The IR spectrum shows the characteristic C=O absorption at 1765 cm⁻¹. Unexpectedly, the NMR spectrum did not display the AB pattern for C₇ methylene protons, but instead a narrow, almost unsplit two proton absorption (a degenerated AB pattern) was observed at δ 1.72 ppm for the protons Hₓ and Hₙ. Apparently, the inside proton Hₙ in half cage ketones 3 and 8 is not significantly shielded by the C=O function. The difference between the Hₙ, Hₓ absorption in 3 and 8 is most likely due to a deshielding of Hₓ in 3 by the anisotropy of the ethylene ketal function and not to
the anticipated shielding effect of the $\text{C} = \text{O}$ group at $\text{C}_4$. The remaining part of the spectrum was entirely consistent with the proposed structure 8, viz the methylene protons at $\text{C}_9$ appear as an AB quartet ($J \sim 12$ Hz) centered at $\delta 2.03$, confirming the asymmetric structure of 8, and the bridgehead protons $\text{H}_2$, $\text{H}_3$, $\text{H}_5$, $\text{H}_6$ absorb at $\delta 3.35$–$3.80$ and $\delta 2.70$–$3.15$ ppm as complex multiplets.

By performing the homoketonization of acetate 7 with NaOMe in MeOD monodeuterated ketone $\text{8a}$ was obtained in quantitative yield. The NMR spectrum exhibits a one proton absorption at $\delta 1.55$–$1.85$ ppm for either the $\text{exo}$ proton $\text{H}_x$ or the $\text{endo}$ proton $\text{H}_n$ (the other signals remained unchanged). To differentiate between $\text{H}_x$ and $\text{H}_n$, ketone 8 and its monodeuterated form $\text{8a}$ were reduced with LAH. In both cases exclusive formation of oxygen inside alcohols viz 9 and 9a was observed (Cf. formation of alcohol 5). The NMR spectrum of 9 exhibits one half of an AB quartet ($J \sim 12$ Hz) as doublets ($J \sim 1.5$ Hz) at $\delta 2.78$ and 2.60 for proton $\text{H}_n$, and a doublet of complex multiplets at $\delta 1.2$–$1.6$ ppm for proton $\text{H}_x$, whereas in the spectrum of $\text{9a}$ the AB pattern at $\delta 2.78$ and 2.60 was absent and the $\text{H}_n$ proton ($\equiv \text{H}_x$) appeared as a simplified multiplet at $\delta 1.38$–$1.58$ ppm.

Comparison of these data indicates that the $\text{H}_n$ proton in $\text{8a}$ was replaced by D. Hence, the introduction of the ethylene ketal group at $\text{C}_9$ in the homocubane system does not alter the stereochemical course of the homoketonization. The acetate 2 as well as 7 gives exclusively cleavage of the $\text{C}_4$–$\text{C}_7$ bond (or the equivalent $\text{C}_3$–$\text{C}_4$ bond) with a subsequent protonation which is highly stereospecific (> 96% retention of configuration at $\text{C}_7$).

The cubane alcohol 10 and its acetate 11 appeared to be even more sensitive towards base than the homocubane bridgehead alcohols and acetates (Scheme 2). On mild treatment with NaOMe or KOH in MeOH, alcohol 10 and acetate 11 reacted instantly to a complex mixture of cage degradation products. The IR and NMR spectra showed several carbonyl and olefinic absorptions. Attempts to separate these products (GLC and TLC) were unsuccessful. Presumably, the half cage ketone 12 has been formed, but because of the large constraint in the half cage system, subsequent fragmentation will take place readily, leading to mixtures of less strained compounds. Thus far, we were unable to isolate or identify any of these compounds.

The far less strained 1,3-bishomocubane bridgehead alcohol 13 and formate 14 did not show homoketonization upon treatment with NaOMe in MeOH at room—or even at reflux temperature. However, another increase in reaction temperature was sufficient to bring about the homoketonization reaction. Alcohol 13 or formate 14 gave, when treated with NaOMe in refluxing EtOH, a sole ketone to which structure 15 was assigned (Scheme 3). The IR spectrum shows a $\text{C} = \text{O}$ absorption at $1730$ cm$^{-1}$ indicative of a cyclopentanone. The NMR spectrum (CDCl$_3$) displays one half of an AB pattern as a doublet of doublets ($J \sim 13$ Hz, $J_2 \sim 1$ Hz) for the inside proton $\text{H}_n$ and $\text{H}_x$ centered at $\delta 2.32$, singlets for the Me groups at $\delta 1.11$ and 0.94, a complex multiplet for $\text{H}_n$ at $\delta 2.0$, a multiplet for the remaining bridgehead protons at $\delta 2.58$–$3.60$, and a symmetrical AA'BB' multiplet for the ethylene ketal function at $\delta 3.78$–$4.43$ ppm. Although the IR and NMR spectra are consistent with the proposed structure 15, they can not differentiate between the ketone 15 and its isomer 16, arising from $\text{C}_4$–$\text{C}_5$ bond cleavage. Unambiguous evidence for structure 15 was obtained from the NMR spectrum of deketalized ketone 18, which was produced in quantitative yield from alcohol 17. The outside proton $\text{H}_n$ and inside proton $\text{H}_x$ in 18 appear as a broad multiplet at $\delta 1.77$. The large down field shift ($\Delta \delta \sim 0.6$ ppm) observed for $\text{H}_x$ in 15 as compared with $\text{H}_x$ in 18 is largely due to the magnetic anisotropy of the ethylene ketal function (vide supra), and can only be reconciled with structures 15 and 18, since in half cage ketone 16 no such deshielding effect of the ethylene ketal group on $\text{H}_x$ is expected. Evidently, homoketonization of 1,3-bishomocubane alcohols proceeds exclusively by cleavage of the $\text{C}_2$–$\text{C}_3$ bond in the direction of the least strained ketone.

The stereochemistry of the cage opening of alco-

![Scheme 2](image-url)
holts 13 and 17 was studied by performing the homoketonization with NaOMe in refluxing EtOD (Scheme 3). The position of deuterium in the so-obtained ketones 15a and 18a was established to be exclusively endo in both cases. The NMR spectrum of 15a was lacking the Hn signal, while the Hx absorption had a simplified pattern as compared with 15. Similarly as described for 9a, the spectrum of the inside alcohol 19a provides a means to show that deuterium had entered exclusively the endo position in 18a. Thus, the presence of the ethylene ketal function at C10 does not alter the stereochemistry of the homoketonization of the 1,3-bishomocubane alcohols.

These homoketonization experiments with the homocubane and 1,3-bishomocubane bridgehead alcohols clearly demonstrate the occurrence of a directional-selective bond cleavage and a strict stereochemical control in the subsequent proton uptake. In Scheme 4 a mechanism for this base induced process is proposed (shown for 20). The high selectivity in the direction of the bond cleavage is most likely governed by relief in cage strain. Inspection of molecular models indeed confirms that cleavage of the C4—C7 (or C3—C4) bond in the homocubane skeleton provides more relief of strain than rupture of the central C4—C5 bond. Similarly, breaking of the C5—C3 bond is most favourable with regard to relief of cage strain in the bis-homocubane system. Two other types of cage compounds, the only ones mentioned in literature so far, viz the birdcage alcohol of Howe and Wistein, and the cage amine of Stedman et al., also show a base induced C—C bond cleavage directed to the least strained half cage compound.

According to Cram’s terminology these cage opening reactions represent examples of an electrophilic substitution at saturated carbon with carbon as leaving group. The stereochemistry of such reactions, particularly the SEi type, has extensively been studied for open systems and monocyclic alcohols (inclusive cyclopropanols). It was found that the change in configuration at the carbon of substitution depends on the nature of the solvent (dissociating power and dielectric constant). However, the base-catalyzed ketonization of the strained 1-hydroxytricyclicene to bornan-2-one takes place with inversion of configuration independent of solvent. Most likely, the views developed for simple systems cannot adequately account for the stereochemical results in their strained counterparts. The stereospecific cleavage of the cyclobutane ring in homocubane and 1,3-bishomocubane takes place with retention of configuration at C7 and C2, respectively. In both
cases proton uptake occurs from the endo-side of the molecule, even in the case of bishomocubane where the endo approach of the proton donating EtOH molecule might be hindered by the gem dimethyl group at C6. This stereospecificity requires either that the cleavage reaction produces a non-inverting carbanion or that the proton donor is incorporated in the transition state of the bond cleavage (concerted process). The very high degree of stereospecificity observed in these cage cleavage reactions argues against the intermediacy of free carbanions. Carbanion mechanisms involving preferential endo-protonation because the departing carbonyl group gives the intermediate carbanion a polar and non-polar face or generation of an energetically unfavourable ion pair during exo-protonation are in our opinion too subtle to explain the exclusive retention of configuration. Cyclobutane ring opening in other cage molecules during homoketonization reactions was also found to occur very predominantly with retention, viz for a bicycage amine and for 7-phenyltricyclo[3.2.0.02,6]heptan-7-ol. Opening of a cyclopentane ring in 3,7-dimethyl tricyclo[3.3.0.03,7]octan-1-ol showed 98% net retention of configuration during the base-induced fragmentation process independent of solvent. We are inclined to believe that the stereochemistry of 90% inversion for the homoketonization of Nickon's nortricyclanol, opposite to the cases described above, must be attributed to the special nature of the cyclopropane ring present in this system.

More information about the fate of the presumed carbanion was obtained when the reaction with base was conducted under aprotic conditions. Accordingly, homocubane alcohol 1 was treated with LiN(iPr)2 in THF (Scheme 5). A sole product was obtained to which structure 21 was assigned. The IR spectrum shows a C=O absorption at 1785 cm⁻¹, typical for a cyclobutanone, and a weak olefinic C—H band at 3080 cm⁻¹. The NMR spectrum displays a narrow multiplet at δ 6.17 ppm for the olefinic protons H2 and H3, an unsymmetrical multiplet at δ 3.6-4.35 ppm for the ethylene ketal group which coincides with one of the bridgehead protons, and a complex multiplet between δ 2.4 and δ 3.3 ppm for the remaining protons. Similar cage fragmentation products were observed in the homocubane methylcyanylide and a 4-homocubane methyl sulfone. A mechanistic pathway for the base induced cage fragmentation reaction is outlined in Scheme 5. Evidently, under the applied aprotic conditions the homoketonization does not stop at the stage of the half cage system, but another C—C bond is broken producing a less congested carbanion. Bishomocubane alcohol 13 on treatment with LiN(iPr)2 in THF or dioxane did not lead to anionic fragmentation neither at room temperature nor in refluxing solvent. Apparently under aprotic conditions the reaction conditions need to be more drastic to achieve cage fragmentation than with a protic solvent.

This extended fragmentation to 21 provides a good means to gather some additional evidence for the concerted mechanism proposed for the homoketonization under protic conditions. When a free carbanion is a true intermediate one would expect at least some further fragmentation of the homocubane alcohol to a structure like 21 when such fragmentation is placed in a fair competition with protonation. However, by performing the homoketonization of 1 with NaOMe as base in dioxane or acetonitril as solvents, containing a trace of MeOH, a quantitative yield of the half-cage ketone 3 was isolated. This result substantially supports a synchronous mechanism for the homoketonization reaction.

EXPERIMENTAL

IR spectra were taken on a Perkin Elmer 125 or 257 grating spectrometer. NMR spectra were recorded on a Varian A60 or T60 spectrometer, using TMS as internal standard. All m.ps are uncorrected and determined on a Kofler hot stage. Elemental analyses were carried out in duplicate (their average values are reported), in the micro analytical department of the University at Groningen under supervision of Mr. W. M. Hazenberg.

1-Bromotetracyclo[4.3.0.02,6]nonan-9-one (3). NaOMe (0.05 g, 0.92 mmole) was added to a stirred soln of acetate 2 (0.1 g, 0.3 mmole) in MeOH (5 ml). After stirring at room temp for 1 hr, MeOH was removed in vacuo, the residue diluted with water and
either extracted. The ether layer was dried (MgSO₄) and the solvent evaporated to give ketone 3 (0.015 g, 0.4 mmole) in ether (5 ml). After stirring overnight, the mixture was diluted with water and ether extracted. The ether layer was dried (MgSO₄) and the solvent evaporated affording half cage ketone 8 (0.015 g, 0.4 mmole). Recrystallization from hexane gave a pure sample, m.p. 75-77°C; IR νmax 1765 (C=O) cm⁻¹; NMR (CDCl₃) δ 3.8-3.4 (m, 3H, CH₃), 2.9 (s, 3H, CH₃), 2.6-2.1 (m, 4H, CH₂), 2.0-1.5 (m, 4H, CH₂), 1.5-1.0 (m, 4H, CH₂), 1.0-0.5 (s, 3H, CH₃), 0.8-0.4 (t, J ~ 12 Hz, 1H, proton H₄). m/e 131 (M⁺, 1 Br). (Found: C, 48.73; H, 4.05; Br, 29.30%; Calc. for C₃H₃BrO: C, 48.37; H, 4.05; Br, 29.30%). The same result was obtained when alcohol 1 was used as starting material; NMR (CDCl₃) δ 4.12 (t, J ~ 6 Hz, 1H, proton HO, 1H), 3.85-3.45 (m, 4H, CH₂), 2.3-3.0 (m, 5H), 1.8-1.5 (m, 2H, methylene protons at C10), 1.1-0.9 (m, 3H, CH₃), 0.8-0.4 (t, J ~ 12 Hz, 1H, proton H₄). m/e 131 (M⁺, 1 Br). 6,6-Dimethyl-5-bromotetrahydro-2-endo-deuterio-tetracyclo[5.3.0.0²,8]dodecan-5-one (5a) was prepared as described above using monodeuterated ketone 8a as starting material; NMR (CDCl₃) δ 3.8-3.4 (m, 3H, CH₃), 2.6-2.1 (m, 4H, CH₂), 2.0-1.5 (m, 4H, CH₂), 1.5-1.0 (m, 4H, CH₂), 1.0-0.5 (s, 3H, CH₃), 0.8-0.4 (t, J ~ 12 Hz, 1H, proton H₄). m/e 131 (M⁺, 1 Br). (Found: C, 53.76; H, 5.66; Br, 25.77; Calc. for C₅H₇BrO₂: C, 53.69; H, 5.47; Br, 25.19%). 6,6-Dimethyl-5-bromo-2-endo-deutero-tetracyclo[5.3.0.0²,8]dodecan-5-one (5a) was prepared as described above using 6,6-dimethyl-5-bromotetrahydro-2-endo-deuterio-tetracyclo[5.3.0.0²,8]dodecan-5-one (5a) as starting material; NMR (CDCl₃) δ 4.12-4.00 (m, 4H, CH₂), 3.85-3.45 (m, 4H, CH₂), 2.3-3.0 (m, 5H), 1.8-1.5 (m, 2H, methylene protons at C10), 1.1-0.9 (m, 3H, CH₃), 0.8-0.4 (t, J ~ 12 Hz, 1H, proton H₄). m/e 131 (M⁺, 1 Br). (Found: C, 56.48; H, 5.93; Br, 31.32%). 6,6-Dimethyl-5-endo-deuterio-tetracyclo[5.3.0.0²,8]dodecan-5-one (1b) was prepared as described above using 5,7-dimethyl-5-endo-deuterio-tetracyclo[5.3.0.0²,8]dodecan-5-one (1b) as starting material; NMR (CDCl₃) δ 4.12-4.00 (m, 4H, CH₂), 3.85-3.45 (m, 4H, CH₂), 2.3-3.0 (m, 5H), 1.8-1.5 (m, 2H, methylene protons at C10), 1.1-0.9 (m, 3H, CH₃), 0.8-0.4 (t, J ~ 12 Hz, 1H, proton H₄). m/e 131 (M⁺, 1 Br). (Found: C, 56.48; H, 5.93; Br, 31.32%; Calc. for C₅H₇BrO₂: C, 56.48; H, 5.93; Br, 31.32%). 6,6-Dimethyl-5-bromo-2-endo-deuterio-tetracyclo[5.3.0.0²,8]dodecan-5-one (1b) was prepared as described above using 6,6-dimethyl-5-bromotetrahydro-2-endo-deuterio-tetracyclo[5.3.0.0²,8]dodecan-5-one (5a) as starting material; NMR (CDCl₃) δ 4.12-4.00 (m, 4H, CH₂), 3.85-3.45 (m, 4H, CH₂), 2.3-3.0 (m, 5H), 1.8-1.5 (m, 2H, methylene protons at C10), 1.1-0.9 (m, 3H, CH₂), 0.8-0.4 (t, J ~ 12 Hz, 1H, proton H₄). m/e 131 (M⁺, 1 Br). (Found: C, 56.48; H, 5.93; Br, 31.32%). 6,6-Dimethyl-5-bromo-2-endo-deuterio-tetracyclo[5.3.0.0²,8]dodecan-5-one (1b) was prepared as described above using 5,7-dimethyl-5-endo-deuterio-tetracyclo[5.3.0.0²,8]dodecan-5-one (1b) as starting material; NMR (CDCl₃) δ 4.12-4.00 (m, 4H, CH₂), 3.85-3.45 (m, 4H, CH₂), 2.3-3.0 (m, 5H), 1.8-1.5 (m, 2H, methylene protons at C10), 1.1-0.9 (m, 3H, CH₂), 0.8-0.4 (t, J ~ 12 Hz, 1H, proton H₄). m/e 131 (M⁺, 1 Br). (Found: C, 56.48; H, 5.93; Br, 31.32%; Calc. for C₅H₇BrO₂: C, 56.48; H, 5.93; Br, 31.32%).
5-one (18a) was prepared as described for 18, using EtOD instead of EtOH; NMR (CDCl₃) δ 2.25-3.40 (m, 5 H), 1.95 (AB quartet, J₁ ~ 11 Hz, J₂ ~ 1.5 Hz, 2H, methylene protons at C₁₁₀), 1.58-1.83 (m, 1H, proton H₂), 1.09 (s, 3H, CH₃), 0.90 (s, 3H, CH₃).

6,6-Dimethyl-9-bromo-5-endo-hydroxytetracyclo[5.3.0.0³,9.0]decane (19). The same procedure as for the preparation of alcohol 5 was used. Ketone 18 gave crude alcohol 19 in 90% yield. Recrystallization from hexane and subsequent sublimation (80°/12 mm) afforded an analytically pure sample: m.p. 84-85°; IR νmax 3400 (O—H) cm⁻¹; NMR (CDCl₃) δ 3.4 (d, J ~ 6 Hz, 1H, proton H₇), 2.2-3.2 (m, 3H), 2.55 and 2.78 (d, one half of an AB quartet, J ~ 12 Hz, 1H, proton H₈), 1.8 (m, 4H, cage protons and methylene protons at C₁₀), 1.3 (m, 1H, proton H₉). (Found: C, 56.71; H, 6.88; Br, 30.70; Calc. for C₁₃H₁₅BrO: C, 56.04; H, 6.67; Br, 31.07%).

6,6-Dimethyl-9-bromo-2-tendo-deuterio-5-tendo-hydroxytetracyclo[5.3.0.0³,9.0]decane (19a) was prepared in the same way, utilizing monodeuterated ketone (18a) as starting material; NMR (CDCl₃) δ 3.4 (d, J ~ 6 Hz, 1H, proton H₉), 2.2-3.2 (m, 3H), 1.8 (m, 4H, cage protons and methylene protons at C₁₁₀), 1.3-1.5 (m, 1H, proton H₈).

1-Bromotricyclo[4.2.1.0⁴,⁸]non-7-ene-4,9-dione 9-ethylene ketal (21). To a stirred ice-cooled soln of diisopropylamine (0.4 g, 4 mmole) in THF (5 ml) was added (N₂ atmosphere) 1.5 ml of 2N n-BuLi in hexane. After 15 min a soln of alcohol 1 (0.4 g, 1.5 mmole) in THF (5 ml) was added gradually. After stirring at room temp for 1 hr, water was added and the mixture ether extracted. The extracts were washed with diluted HCl aq and dried (MgSO₄). Solvent was removed yielding crude 21 (0.23 g, 58%) as an oil. Crystallization from hexane and subsequent sublimation (100°/12 mm) afforded a pure sample, m.p. 98-102°; IR νmax 3080 (C=CH), 1785 (C=O) cm⁻¹; NMR (CDCl₃) δ 6.17 (d, degenerated AB quartet, 2H, protons H₁, H₉), 3.6-4.35 (m, 5H, ethylene ketal protons and one of the ring protons), 2.4-3.3 (m, 4H). m/e 271 (M⁺, 1 Br), 149 (M—Br—CH₃=C=O). (Found: C, 48.78; H, 4.24; Br, 28.87; Calc. for C₁₃H₁₅BrO₂: C, 48.73; H, 4.09; Br, 29.48%).

REFERENCES

5. T. Fukunaga, Ibid. 87, 916 (1965).