Alkali Metal Picrate Complexes of "Crowned"-Phthalocyanines. Solid State Structures and Electrical Properties

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Molecules of copper phthalocyanine substituted with four 18-crown-6 moieties, were aggregated with the help of K+, Rb+, and Cs+ picrates. Possible structures were proposed based on computer analysis of X-ray powder diffraction patterns: for the crowned phthalocyanine an orthorhombic cell with: a = 50.77, b = 25, c = 8.7 Å. For its metal picrate complexes, tetragonal structures were found: a = b = 22.76, c = 12.99 Å (K+-complex); a = b = 29.44, c = 10.63 Å (Cs+-complex); a = b = 29.44, c = 10.40 Å (Rb+-complex). A.c. impedance spectroscopy showed an increase in the electronic conductivity of 2 to 3 powers of ten, if crowned phthalocyanine was aggregated by means of a metal picrate salt. In the K+ complex only electronic conductivity could be detected. The Rb+ and Cs+ complexes exhibited cation conductivity as well.

Introduction

Recently much effort has been directed to the construction of low dimensional organic conductors comprising phthalocyanines (Pcs). Well known are single crystals of partially oxidized (metallo-) Pcs [1], Pc's as discotic liquid crystals [2], and metallo-Pc's polymerized through suitable bidentate ligands [3]. All of these systems consist of columns of face to face stacked macrocycles with small inter-planar distances. Charge transport of electrons or holes occurs via π—π overlap of the macrocyclic ligands, via a one dimensional chain of metal centres, or via a metal-ligand-metal chain. We achieved an alternative way of stacking by adding cations to Pc's to which four crown ether rings are attached [4].

In solution "crowned"-phthalocyanine 1 tends to form aggregates. Aggregation is induced by alkali metal ions which coordinate to the crown ether rings. This aggregation is particularly notable when the diameter of the ion exceeds the inner diameter of the crown ether rings. In these aggregates the cations are sandwiched between two crown ether rings. Also, by using cations with different diameters, we are able to influence the inter-planar distance and staggering angles between the Pc ligands. In an earlier communication [5], we reported on the electrical conductivity of 1 in its solid state. This conductivity increases when 1 is complexed with alkali metal picrates. Picrate salts were used to reduce anion conduction by steric hindrance.

In the present paper we propose a structure of 1 and of its K+, Rb+, and Cs+ picrate complexes on the basis of X-ray powder diffraction patterns. Moreover, the electrical properties of the solid state compounds were studied by a.c. impedance spectroscopy. CuPc was used as a reference compound.

Experimental

Preparation of alkali metal picrate complexes of (4,5,4',5',4'',5'',4''',5'''-tetakis(1,4,7,10,13,16-hexaoxa-hexadecamethylene)phthalocyaninato)copper(II) (1)

Compound 1 was synthesized as described previously [4a]. Its K+, Rb+ and Cs+ picrate complexes were prepared by mixing 1 (0.07 mol) with the appropriate picrate salt in a 1:4 ratio in 20 ml dry chloroform-methanol (1:1 v/v) and stirring for two days at 40°C under an atmosphere of dry nitrogen. A precipitate was formed which was isolated by filtration and washed with hot chloroform, and subsequently with hot methanol until a colourless filtrate was obtained. The product was dried in vacuum over P2O5 for two days at 150°C yielding a green, insoluble, micro-crystalline powder. Yield: 80%.

K+-picrate complex of 1. T.G.A. 306°C (decomp.); Anal. Calcd. for CuK6C6H42N4S2O32 (K+-picrate): K+, 6.10; C, 44.66; H, 3.75; O2.50/0; N, 9.96. Yield: 80%.
N, 10.85; O, 32.22. Found: K+, 6.3; C, 44.54; H, 3.82; N, 10.68; O, 32.21.

Rb+-picrate complex of 1. T.G.A. 306°C (decomp.); Anal. Calcd. for CuxRbyCpxHpyNqOq (1. (Rb+-picrate)z): C, 41.67; H, 3.50; N, 10.13; O, 30.05. Found: C, 41.49; H, 3.70; N, 10.08; O, 30.03.

X-Ray Diffraction

X-ray powder diffraction patterns of all polycrystalline compounds studied were recorded on a Guinier Johansson FR 552 powder diffractometer (ENRAF-NONIUS-CuKα), and on a Kossiakoff powder diffractometer using Ni-filtered CuKα radiation (100 and 400 nm sample-film distances were used). Unit cells of the studied compounds were generated using CPK-models, and a local computer program based on the trial and error method of Henry, Upson, and Wooster [5].

A.c. Impedance Spectroscopy

To prepare samples the materials were powdered in a mortar and pressed at a load of 3 tons in disc-shaped compacts with a diameter of 6 mm, and a thickness of 0.6 to 1.2 mm. Sputtered platinum electrodes were applied using an Edwards Sputter Coater S 150B. The samples were mounted in a conventional conductivity interface. Reliable conductance values could be measured from Arrhenius plots of the temperature dependence of the conductivity, \( \sigma = \sigma_0 \exp\left(-E_a/RT\right) \). Dielectric constants \( \varepsilon \) were calculated from the high-frequency capacitance \( C_m = \varepsilon_0 A/d \) where \( \varepsilon_0 = 8.85 \times 10^{-12} \text{ F/m} \), while \( d \) and \( A \) are the thickness, and electrode surface area of the samples.

Seebeck Measurements

The type of charge carriers in the samples was determined between 150 and 170°C using the hot point probe technique. Seebeck voltages were measured with a Keithly 614 electrometer.

Results and Discussion

Description of the Structure

The sharp X-ray powder diffraction patterns of 1 and of its metal picrate complexes indicate the crystalline nature of the samples. Using a computer program, CPK models, and literature data on Pc's, we tried to derive the structure of each compound from its diffraction data. Since powder patterns do not provide a unique structure, we were only able to propose structures that are not in contradiction with the X-ray patterns.

For 1 and its copper free analogue we observed lattice spacings similar to the crystalline phase observed by Simon et al. [9]. We did not observe a mesophase. However, the crystalline structure did not explain the presence of an additional spacing at 50.77 Å in our diffraction patterns. This 50.77 Å spacing indicates a larger unit cell. An orthorhombic unit cell with parameters \( a = b = 20.8 \), \( c = 8.2 \) Å \( (\alpha = \beta = \gamma = 90°) \). They also observed a mesophase comprising staggered Pc's with unit cell parameters \( a = b = 20.8 \), \( c = 8.2 \) Å \( (\alpha = \beta = \gamma = 90°) \). For both structures an interplanar Pc distance of 3.4 Å and a 4.2 Å distance for the crown ethers was proposed.

To propose structures that are not in contradiction with the X-ray patterns.
picrate anions. The diffraction patterns of the rubidium, and cesium picrate complexes of 1 were very similar, but different from their potassium counterpart. The powder diffraction patterns of both the rubidium and cesium complex indicated that ordering of the Pc stacking was very high, higher than the Pc stack of the potassium complex of 1. In UV/Vis experiments we observed aggregation upon adding rubidium and cesium salts to a solution of 1 [4a]. In these aggregates the cations are sandwiched between the crowns, and the Pc macrocycles have an eclipsed conformation. Rubidium and cesium ions have diameters that exceed the inner diameter of a planar 18-crown-6 ring. We assumed a similar structure to be possible in the solid state. Because of the resemblance of the powder diffraction patterns and the UV/Vis experiments, we assume the same stacking for both the rubidium and cesium complex of 1. Remarkable in the diffraction pattern of the cesium complex were four sharp reflections in the region between 3.8 and 3.4 Å. In the diffraction pattern of the rubidium complex similar reflections were found, all of them at 0.1 Å shorter distances, i.e. between 3.7 and 3.3 Å. This indicates an interplanar distance that is 0.1 Å smaller in the rubidium as compared to the cesium complex. The data of the cesium picrate complex could be fitted to a tetragonal unit cell \( a = b = 29.44, c = 10.63 \) Å, and those of the rubidium picrate complex to a tetragonal unit cell \( a = b = 29.44, c = 10.40 \) Å. A probable molecular arrangement that fits these unit cells is schematically given in Fig. 3.

Although the X-ray measurements gave sharp powder diffraction patterns, it proved to be impossible to determine exact positions of the molecules in the proposed cells.

**Electrical Measurements**

The electrical properties of phthalocyanine 1, and of its K\(^+\), Rb\(^+\), and Cs\(^+\) picrate complexes were studied in the solid state by a.c. impedance spectroscopy. CuPc was used as a reference compound. Fig. 4 shows typical impedance spectra for the rubidium picrate complex of 1, \( 1 \cdot \text{Rb}^+ \), at 155°C. Similar spectra were obtained for 1 and for its other metal picrate complexes. All spectra suggest that electrode polarisation phenomena are absent and that electronic conductivity predominates.

Seebeck measurements revealed compound 1 and CuPc to be p-type electronic conductors, as has earlier been observed for CuPc\(^{11}\), whereas the K\(^+\), Rb\(^+\), and Cs\(^+\) picrate complexes of 1 exhibit n-type electronic conductivity. The rubidium picrate complex of the metal free derivative of 1 showed n-type conductivity as well. For copper phthalocy-
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parallel constant phase element (CPE) \( Q_p \), with an impedance \( Z = (\varepsilon_0 \sigma)^{-1} \) has values for \( \sigma \) of 0.5 to 0.7. This element is related to ionic diffusion. Usually, ionic conduction should be represented with an \( R_C \) branch, \( R_C^{-1} \) being the ionic conductance and \( C \) the double layer capacitance representing the ion blocking effect of the Pt electrode. Fits with the \( R_C \) branch included, led to insignificant values for \( R \), and values for \( C \) comparable to \( \varepsilon_0 \). This may indicate that the mechanism of ionic conduction is dominated by diffusion. Fig. 6 shows the temperature dependence of \( Q_p \), plotted as \( \ln \left( k \cdot \frac{d}{A} \cdot T \right) \) versus \( 1/T \). This figure points to a thermally activated transport of \( \text{Rb}^+ \) and \( \text{Cs}^+ \). The slopes of the curves correspond to a conduction activation enthalpy, \( \Delta H(Q_p) \), of 0.64 eV for the \( \text{Rb}^+ \), and 0.59 eV for the \( \text{Cs}^+ \) complex of 1. This may be related with binding of \( \text{Rb}^+ \) and \( \text{Cs}^+ \) to 18-crown-6 [4a]. The larger \( \text{Cs}^+ \)-ion has a lower binding energy with the 18-crown-6 rings of 1 than the smaller \( \text{Rb}^+ \)-ion. It is, therefore, easier to transport \( \text{Cs}^+ \) through the crystal lattice than \( \text{Rb}^+ \). As can be seen in Fig. 6, the \( \text{K}^+ \)-ion exhibits a different behaviour. The slope of this curve does not refer to thermally activated ion conduction. In fact, CPE \( Q_p \) in Fig. 5, is not required for fitting the impedance spectra of \( \cdot \text{K}^+ \). Adequate fit results are obtained also with the equivalent circuit \( C_{m} R_s p Q_2 s \), Po-
potassium has a much larger binding energy with 18-crown-6 than rubidium and cesium. As can be seen in Fig. 2, the potassium ions fit in the crown-ether rings and are closely packed with picrate anions in between. Hence, the K⁺ ions are practically immobile. Potassium conduction has been demonstrated to occur through an 18-crown-6 channel [14]. For further examination of a possible contribution of K⁺ ions to the conductivity, the open-circuit voltage (OCV) of the concentration cell K-amalgam(I) |1 • K⁺| Hg(II) was measured during discharge of this cell. The OCV did not decrease in a subsequent period of discharge of 40 h. From this it is concluded that potassium ions are not involved in d.c. transport in 1 • K⁺.

The temperature dependence of the electrical conductivity σ of the prepared compounds, plotted as ln(σ(T)) versus 1/T is shown in Fig. 7. The isothermal electrical conductivity at 180°C and the conductivity activation energy, Eₐ, are presented in Table 1. As can be seen in Table 1 and Fig. 7 compound 1 shows a higher electrical conductivity and a lower activation energy than CuPc. A substantial increase in electrical conductivity occurs when 1 is complexed with an alkali metal ion. This increase is larger for Rb⁺ and Cs⁺, than for K⁺. This can be explained by the fact that complexation of an alkali metal ion results in aggregates with eclipsed phthalocyanine units. Rb⁺ and Cs⁺ ions slightly decrease the interplanar spacing and change the angle of staggering between the macrocycles as is concluded from the X-ray powder diffraction patterns, resulting in an increased π–π overlap [15]. This overlap could be higher for Rb⁺ than for Cs⁺ because the latter has a larger diameter. In the K⁺ complex of 1, picrate anions are intercalated between eclipsed macrocycles, but the ordering of the Pc stacking is less than for the rubidium and cesium complexes. The overall effect is a higher conductivity activation energy. From the presented data it can be concluded that the electrical properties are affected by a variation in the angle of staggering between the Pc macrocycles.

### Table 1

<table>
<thead>
<tr>
<th>Compound</th>
<th>σ_{180°C} (S • m⁻¹)</th>
<th>Eₐ (eV)</th>
<th>Type of conductivity</th>
</tr>
</thead>
<tbody>
<tr>
<td>CuPc</td>
<td>6.10⁻⁴</td>
<td>1.39</td>
<td>p</td>
</tr>
<tr>
<td>1</td>
<td>5.10⁻⁷</td>
<td>1.18</td>
<td>p</td>
</tr>
<tr>
<td>1 • K⁺</td>
<td>1.10⁻⁵</td>
<td>1.33</td>
<td>n</td>
</tr>
<tr>
<td>1 • Rb⁺</td>
<td>3.10⁻⁴</td>
<td>1.02</td>
<td>n</td>
</tr>
<tr>
<td>1 • Cs⁺</td>
<td>7.10⁻⁴</td>
<td>1.09</td>
<td>n</td>
</tr>
</tbody>
</table>

For all samples the CPE Q₂, with an impedance Z = (k₀/iω)⁻¹, has a value for a between 0.1 and 0.4, and has, therefore, the character of a resistance. We relate the presence of this CPE to current inhomogeneities, which occur due to barriers from cross-linking of the metal picrate aggregates of 1 as visualised in Fig. 8. Charge transport requires hopping between phthalocyanine stacks. The activation energies Eₐ(Q₂) for this hopping process are given in Table 2. These values are similar to the activation energy, found by Belarbi et al. [16], required for intermolecular electron hopping between Pc₂Lu subunits. The conduction activation energies, Eₐ, related to charge transport within

![Graph 1](image1.png)

**Fig. 7**

Temperature dependence of the electrical conductivity (σ), plotted as ln(σ) versus 1/T for 1, and of its metal picrate complexes

![Graph 2](image2.png)

**Fig. 8**

Cross-linking of phthalocyanine stacks
The metal picrate complexes of 1 are intrinsic semiconductors with a larger mobility of the electrons than that of the holes.

The n-type K⁺, Rb⁺, and Cs⁺ picrate complexes of 1 show Meyer-Neldel behaviour, i.e. a linear relationship between the isothermal conductivity plotted as log(σ(T)) and Eᵣ [17]. However, the slope of the log(σ(180°C)) versus Eᵣ curve is smaller than the expected value of 0.434/eV/kT, indicating an additional temperature independent term preceding the term exp(−Eᵣ/kT) as has been found for semiconducting oxidized cholesterol, retinal, RNA and DNA [18]. A possible explanation for this behaviour is tunneling of thermally excited electrons through intermolecular barriers, according to the Many, Harnik, and Gerlich model. This tunneling leads to an activation energy dependent mobility [19].

The dielectric constant (ε), obtained from the high-frequency capacitance Cₒₒ of polycrystalline 1, its K⁺, Rb⁺, and Cs⁺ complexes, and CuPc are listed in Table 3. These values are higher than earlier published data on phthalocyanines [20]. The εₒₒ values are, however, strongly dependent on the morphology, preparation, and compaction methods of the samples. In this respect, the present samples differ from the samples in Ref. [20].

<table>
<thead>
<tr>
<th>Compound</th>
<th>Eᵣ(Qₑ) (eV)</th>
</tr>
</thead>
<tbody>
<tr>
<td>1 - K⁺</td>
<td>0.55</td>
</tr>
<tr>
<td>1 - Rb⁺</td>
<td>0.45</td>
</tr>
<tr>
<td>1 - Cs⁺</td>
<td>0.43</td>
</tr>
</tbody>
</table>

The metal picrate complexes of 1, its K⁺, Rb⁺, and Cs⁺ picrate complexes, and CuPc are listed in Table 3. These values are higher than earlier published data on phthalocyanines [20]. The εₒₒ values are, however, strongly dependent on the morphology, preparation, and compaction methods of the samples. In this respect, the present samples differ from the samples in Ref. [20].

<table>
<thead>
<tr>
<th>Compound</th>
<th>ε</th>
</tr>
</thead>
<tbody>
<tr>
<td>1</td>
<td>82.4</td>
</tr>
<tr>
<td>1 - K⁺</td>
<td>74.0</td>
</tr>
<tr>
<td>1 - Rb⁺</td>
<td>70.9</td>
</tr>
<tr>
<td>1 - Cs⁺</td>
<td>62.2</td>
</tr>
<tr>
<td>CuPc</td>
<td>67.0</td>
</tr>
</tbody>
</table>

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References