VINYL ORTHOFORMATES AND VINYL ACETALS. PART I
Synthesis of divinylorthoxymethyl and diaryloxymethyl halides and carboxylates

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Diaryloxymethyl halides and divinylorthoxymethyl halides, formates, trichloroacetates, monochloroacetates, and acetates have been synthesized from corresponding vinyl orthoformates.

In a previous paper\(^1\) the mechanism of the acid-catalyzed decomposition of compounds \(R^1R^2CXOCH(OR^3)_2\) (I) according to

\[
\begin{align*}
\text{RX} & \quad \text{OR}^3 \\
\text{R}^1\text{C}-\text{O} & \quad \text{C}-\text{OR}^3 \\
\text{R}^2 & \quad \text{H} \\
\text{I} & \quad \text{II} & \quad \text{III}
\end{align*}
\]

\(\text{R}^1\text{C}(-\text{O})\text{R}^2 + \text{X}(-\text{C})\text{OR}^3\) (1)

has been discussed for \(X = \text{OR}\) or \(\text{CN}\). The decomposition proceeds \(\text{via} R^1R^2\text{C(OH)}X\) (IV) as an intermediate and is faster as the intermediate has lower stability.

This decomposition might be used for the synthesis of reactive compounds III with \(X = \text{halogen or acyloxy}\), provided a suitable preparative method for the appropriate derivative (I) be available. For this purpose we decided to investigate reactions of orthoformates, containing one or more vinyl residues* \(\text{HC(OR)}_2\text{OCH}==\text{CH}_2\) (V), with hydrogen chloride, hydrogen bromide, or carboxylic acids.

Dimethoxy-vinlyloxymethane \(\text{HC(OCH}_3)_2\text{OCH}==\text{CH}_2\) could be obtained by exchange of an alkoxy group\(^2\,^3\) from methyl orthoformate and \(\beta\)-chloroethanol,

* Trivinyl orthoformate and triisopropenyl orthoformate have recently also been prepared by H. Stetter and E. Reske, Chem. Ber. 103, 639 (1970).
followed by dehydrohalogenation with tert-butoxide,

\[
\text{HC(OCH}\text{\textsubscript{3}}\text{)}\text{\textsubscript{2}}\text{OCH\textsubscript{2}}\text{CH}\text{\textsubscript{2}}\text{Cl} + \text{t\textcdot BuO}^- \rightarrow \text{HC(OCH}\text{\textsubscript{3}}\text{)}\text{\textsubscript{2}}\text{OCH=CH}\text{2} + \text{tert-BuOH} + \text{Cl}^- \quad (3)
\]

The latter reaction was performed by dropping the β-chloroethyl orthoformate into a suspension of sodium hydride in 1,2-dimethoxyethane to which some tert-butyl alcohol had been added. During the reaction the eliminating agent, tert-butoxide, is continuously re-formed. Dimethyl sulfoxide could be used as the solvent instead of 1,2 dimethoxyethane, but in that case the product was contaminated with an unpleasantly smelling impurity.

The synthesis of the corresponding ethyl compound, HC(O\text{C\textsubscript{2}}\text{H}\text{\textsubscript{5}}\text{)}\text{\textsubscript{2}}\text{OCH\textsubscript{2}}\text{CH}\text{\textsubscript{2}}\text{Cl}, via reaction (2) failed because of serious disproportionation. The compound could be obtained, however, from diethoxymethyl acetate, HC(O\text{C\textsubscript{2}}\text{H}\text{\textsubscript{5}}\text{)}\text{\textsubscript{2}}\text{OCOCH}\text{3}, and β-chloroethanol in the presence of an excess of triethylamine.

The reactivity of vinyl orthoformates (V) towards acids was followed by NMR measurements on reaction mixtures of HC(O\text{C\textsubscript{2}}\text{H}\text{\textsubscript{3}}\text{)}\text{\textsubscript{2}}\text{OCH=CH}\text{2} or HC(O\text{C\textsubscript{2}}\text{H}\text{\textsubscript{5}}\text{)}\text{\textsubscript{2}}\text{OCH=CH}\text{2} and formic acid. They revealed that dialkoxymethyl formates (III, X = \text{HCOO}) and acetaldehyde (II, R\textsuperscript{2} = H) were formed.

It is not yet clear whether the products arise via an intermediate addition product I, R\textsuperscript{1} = \text{CH}\text{3}, R\textsuperscript{2} = \text{H} (reaction 4), which decomposes in an acid medium (reaction 1), or by simple acid-catalysed substitution.

\[
\text{HC(OR)}\text{\textsubscript{2}}\text{OCH=CH}\text{2} + \text{HX} \rightarrow \text{HC(OR)}\text{\textsubscript{2}}\text{OCHXCH}\text{3} \quad (4)
\]

The fact that even a weak electrophile like bromine converts HC(O\text{C\textsubscript{2}}\text{H}\text{\textsubscript{3}}\text{)}\text{\textsubscript{2}}\text{OCH=CH}\text{2} nearly quantitatively into bromoacetaldehyde and the decomposition products of HC(O\text{C\textsubscript{2}}\text{H}\text{\textsubscript{3}}\text{)}\text{\textsubscript{2}}\text{Br}, methyl formate, and methyl bromide, in this case seems to be more consistent with the former possibility. The exact mechanism is under further investigation, together with that of analogous reactions with vinyl acetals.

The preparation of dialkoxymethyl carboxylates or halides by the reactions mentioned is complicated by disproportionation\textsuperscript{4} of the vinyl orthoformate (V) in the presence of acids according to:

\[
\text{HC(OR)}\text{\textsubscript{2}}\text{OCH=CH}\text{2} \xrightleftharpoons{\text{H}^+} \text{HC(OR)}\text{\textsubscript{3}} + \text{HCOR(OCH=CH}\text{2})\text{\textsubscript{2}} \quad (5)
\]

and by a reaction between the products:

$$\text{HC(OR)₂X + CH₃CHO} \rightarrow \text{CH₃CH(OR)X + HCOOR} \quad (6)$$

Compounds $\text{HC(OR)₂X}$ (R = alkyl, X = acetate or formate), therefore, can be synthesized more effectively by a procedure previously described\(^5\).

Because aryl orthoformates disproportionate much more slowly than alkyl orthoformates, and diaryloxymethyl halides are more stable than their alkoxy analogues\(^6\), several diaryloxymethyl halides could be obtained from the corresponding diaryloxy-vinylloxymethanes, $\text{HC(OAr)₂OCH=CH₂}$ (VI), and hydrogen chloride or hydrogen bromide.

The method appeared to be especially useful for the synthesis of compounds such as $\text{HC(OCH=CH₂)₂X}$ (VII) and $\text{HC(OCC₂H₅)₂X}$, (VIII), which have not been described before. Because reaction (6) is very slow with acetone instead of acetaldehyde\(^7\), the yields of the latter compounds (VIII) are generally higher (Table III).

Divinylloxymethyl halides are interesting compounds for the synthesis of tetravinylxythenes. Diaryloxymethyl halides give high, dialkoxy-methyl halides low yields of “ethenes” on treatment with a strong base. The reactivity of divinylloxymethyl halides towards bases lies between those of the corresponding aryloxy and alkoxy compounds. These results will be discussed more fully in a subsequent paper, together with the possible occurrence of carbenes as intermediates.

**Experimental** (with the technical assistance of Mrs. F. Janssen).

1.1 *β*-chloroethyl orthoformates and 2-chloro-1-methylethyl orthoformates (Table I).

Ethyl or methyl orthoformate was mixed with a proper alcohol\(^*\) in the ratio required for the exchange of one, two, or three alkoxy groups. Some drops of trifluoro-acetic acid were added, and ethanol or methanol was removed at 15 mm or 5 cm pressure, depending on the boiling points of the components used. After enough alcohol had been collected (4–8 h), the *β*-chloroethyl orthoformates were distilled at reduced pressure after neutralization of the mixture with sodium methoxide.

1.2 *β*-chloroethyl diethyl orthoformate (Table I).

20 g (0.25 mole) of *β*-chloroethanol dissolved in 75 ml of pentane was added to a mixture of 40.5 g (0.25 mole) of diethoxymethyl acetate\(^5\) and 34 g (0.35 mole) of triethylamine. After standing for 30 minutes, the solution was washed with water four times, dried on sodium sulfate, and distilled.

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\(^*\) Pure 1-chloro-2-propanol was obtained from a sample containing 25% of 2-chloro-1-propanol (Fluka AG., Chemische Fabrik, Buchs SG) by distillation with a Nester Faust spinning band column.


Table I

\(\beta\)-chlo-roethyl orthoformates, HC(OR)\(_2\)OCHR'CH\(_2\)Cl

<table>
<thead>
<tr>
<th>R</th>
<th>R'</th>
<th>Bp</th>
<th>(n_D^{20})</th>
<th>Yield</th>
<th>Method</th>
</tr>
</thead>
<tbody>
<tr>
<td>methyl</td>
<td>H</td>
<td>69°/14 mm</td>
<td>1.4230</td>
<td>45%</td>
<td>1.1</td>
</tr>
<tr>
<td>ethyl</td>
<td>H</td>
<td>85°/13 mm</td>
<td>1.4204</td>
<td>80%</td>
<td>1.2</td>
</tr>
<tr>
<td>phenyl</td>
<td>H</td>
<td>185°/1.3 mm</td>
<td>1.5525</td>
<td>60%</td>
<td>1.3</td>
</tr>
<tr>
<td>p-chlorophenyl</td>
<td>H</td>
<td>not isolated</td>
<td>1.3</td>
<td>1.3</td>
<td></td>
</tr>
<tr>
<td>p-tolyl</td>
<td>H</td>
<td>190°/1.3 mm</td>
<td>1.5425</td>
<td>50%</td>
<td>1.3</td>
</tr>
<tr>
<td>(\beta)-chlo-roethyl*</td>
<td>H</td>
<td>125°-127°/0.9 mm</td>
<td>1.4685</td>
<td>80%</td>
<td>1.1</td>
</tr>
<tr>
<td>2-chloro-1-methylethyl**</td>
<td>CH(_3)</td>
<td>107°/0.6 mm</td>
<td>1.4622</td>
<td>86%</td>
<td>1.1</td>
</tr>
</tbody>
</table>

* Lit. Bp 163°-165°/11 mm.⁹
** Lit. Bp 105°-107°/0.5 mm.⁹

1.3 Aryl \(\beta\)-chlo-roethyl orthoformates (Table I).

Tri-\(\beta\)-chlo-roethyl orthoformate was mixed with 2 equivalents of a phenol. The mixture was boiled at reduced pressure (15 mm) and \(\beta\)-chlo-roethanol was distilled off (Vigreux column 100 × 15 mm) until about \(\frac{1}{2}\) equivalent had been removed (24-48 h). From the relative residues, diphenyl and di-p-tolyl \(\beta\)-chlo-roethyl orthoformate were isolated by distillation. Because of the high boiling point of the corresponding p-chlorophenyl compound the p-chlorophenyl orthoformates HC(OCH\(_2\)CH\(_2\)Cl)\(_2\)OC\(_6\)H\(_4\)Cl and HC(OCH\(_2\)-CH\(_2\)Cl)(OC\(_6\)H\(_4\)Cl)\(_2\) were not separated and used as a mixture in the reaction with NaH. The resultant vinyl orthoformates boil at a much lower temperature and can easily be separated.

2.1 Vinyl orthoformates (Table II).

To a suspension of NaH in 1,2-dimethoxyethane (100 ml per mole of NaH) 20 g of tert-buty]l alcohol per 100 ml of solvent was added. The \(\beta\)-chlo-roethyl orthoformate was added dropwise and the reaction system was then refluxed for about 2 h until the evolution of hydrogen had stopped. In all the experiments NaH was used in about 10% excess. After reaction, the 1,2-dimethoxy-ethane was removed at 15 mm pressure; then water and ether were added carefully, and the ether solution was washed several times with water, dried, and distilled.

Di-p-chlorophenyl vinyl orthoformate was prepared from the mixture of p-chlorophenyl orthoformates described under 1.3. This procedure can also be used with the other aryl \(\beta\)-chlo-roethyl orthoformates. In these cases also arylvinyl orthoformates were isolated.

\[
\begin{align*}
\text{HC(OC}_6\text{H}_4)(\text{OCH}==\text{CH}_2)\_2 & \quad \text{Bp 112°/13 mm, } n_D^{20} 1.5043 \\
\text{HC(OC}_6\text{H}_4-\text{CH}_3-\text{p})(\text{OCH}==\text{CH}_2)\_2 & \quad \text{Bp 82°/0.7 mm, } n_D^{20} 1.5028 \\
\text{HC(OC}_6\text{H}_4-\text{Cl-}\text{p})(\text{OCH}==\text{CH}_2)\_2 & \quad \text{Bp 94°/0.7 mm, } n_D^{20} 1.5165
\end{align*}
\]

The NMR spectra in CCl\(_4\)δ[HC(OR)\(_3\)] showed 6.00, 5.90, and 6.35 ppm, respectively.

3.1 Divinyl(oxy)- and diaryloxy(methyl) halides.

The chlorides were prepared by adding a 4 N HCl solution in ether (in 10% excess) to the proper vinyl orthoformates at room temperature. After standing for 15 minutes, the ether was removed at reduced pressure and the compounds were distilled.
### Table II

Vinyl orthoformates, $\text{HC(OR)}_2\text{OCR'}=\text{CH}_2$

<table>
<thead>
<tr>
<th>$\text{R}$</th>
<th>$\text{R'}$</th>
<th>$\delta \text{D}^{20}$</th>
<th>$\text{Bp}$</th>
<th>$\delta \text{HC(OR)}_2$</th>
<th>Yield</th>
</tr>
</thead>
<tbody>
<tr>
<td>methyl</td>
<td>H</td>
<td>1.4011</td>
<td>119°</td>
<td>5.55</td>
<td>65%</td>
</tr>
<tr>
<td>ethyl</td>
<td>H</td>
<td>1.4027</td>
<td>42°/14 mm</td>
<td>5.50</td>
<td>65%</td>
</tr>
<tr>
<td>phenyl</td>
<td>H</td>
<td>1.5430</td>
<td>135°/1.4 mm</td>
<td>6.45</td>
<td>75%</td>
</tr>
<tr>
<td>$p$-chlorophenyl</td>
<td>H</td>
<td>1.5642</td>
<td>162°/0.7 mm</td>
<td>6.20</td>
<td>45%</td>
</tr>
<tr>
<td>$p$-tolyl</td>
<td>H</td>
<td>1.5429</td>
<td>136°/0.2 mm</td>
<td>5.85</td>
<td>70%</td>
</tr>
<tr>
<td>vinyl*</td>
<td>H</td>
<td>1.4313</td>
<td>38°/13 mm</td>
<td>70%</td>
<td></td>
</tr>
<tr>
<td>1-methylvinyl**</td>
<td>CH$_3$</td>
<td>1.4381</td>
<td>72°/14 mm</td>
<td>6.20</td>
<td>85%</td>
</tr>
</tbody>
</table>

* Lit. Bp 141.5–143°
** Lit. Bp 70–73°/10 mm

### Table III

Dialkoxy methyl compounds, $\text{HC(OR)}_2X$

<table>
<thead>
<tr>
<th>$\text{R}$</th>
<th>$\text{X}$</th>
<th>$\text{B.p.}$</th>
<th>$\delta \text{HC(OR)}_2X$</th>
<th>Yield</th>
</tr>
</thead>
<tbody>
<tr>
<td>vinyl</td>
<td>CH$_3$COO</td>
<td>84°/48 mm</td>
<td>6.80</td>
<td>40%</td>
</tr>
<tr>
<td>vinyl</td>
<td>CH$_3$CICOO</td>
<td>98–100°/15 mm</td>
<td>6.90</td>
<td>55%</td>
</tr>
<tr>
<td>vinyl</td>
<td>CCl$_3$COO</td>
<td>66°/2.5 mm</td>
<td>6.95</td>
<td>60%</td>
</tr>
<tr>
<td>vinyl</td>
<td>HCOO</td>
<td>50–51°/44 mm</td>
<td>6.90</td>
<td>55%</td>
</tr>
<tr>
<td>vinyl</td>
<td>Cl</td>
<td>37–38°/17 mm</td>
<td>6.90</td>
<td>50%</td>
</tr>
<tr>
<td>phenyl</td>
<td>Cl</td>
<td>134°/0.8 mm*</td>
<td>7.25</td>
<td>60%</td>
</tr>
<tr>
<td>phenyl</td>
<td>Br</td>
<td>150–155°/1 mm</td>
<td>7.85</td>
<td>50%</td>
</tr>
<tr>
<td>$p$-chlorophenyl</td>
<td>Cl</td>
<td>152°/0.4 mm</td>
<td>**</td>
<td>50%</td>
</tr>
<tr>
<td>$p$-tolyl</td>
<td>Cl</td>
<td>155°/1 mm</td>
<td>7.40</td>
<td>65%</td>
</tr>
<tr>
<td>1-methylvinyl</td>
<td>Cl</td>
<td>59°/15 mm</td>
<td>7.15</td>
<td>90%</td>
</tr>
<tr>
<td>1-methylvinyl</td>
<td>HCOO</td>
<td>72°/15 mm</td>
<td>7.20</td>
<td>85%</td>
</tr>
<tr>
<td>1-methylvinyl</td>
<td>CH$_3$COO</td>
<td>78°/15 mm</td>
<td>7.10</td>
<td>70%</td>
</tr>
<tr>
<td>1-methylvinyl</td>
<td>ClCH$_3$COO</td>
<td>82°/3.3 mm</td>
<td>7.20</td>
<td>80%</td>
</tr>
<tr>
<td>1-methylvinyl</td>
<td>Cl$_3$CCOO</td>
<td>85°/1.7 mm</td>
<td>7.10</td>
<td>85%</td>
</tr>
</tbody>
</table>

* Bp lit.$^8$ 128°/0.15 mm.
** Absorption is hidden under the aromatic proton absorptions. The structure of the compound appeared from its mass spectrum; peaks at $m/e = 302$, $\text{HC(OC}_6\text{H}_4\text{Cl)}_2\text{Cl}^+$; $m/e = 267$ ($\text{HC(OC}_6\text{H}_4\text{Cl)}_2\text{H}^+$; $m/e = 175$ ($\text{HC(OC}_6\text{H}_4\text{Cl)}\text{Cl}^+$).

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The monochloro- and trichloro-acetates were prepared in a similar way. These reaction mixtures were kept in a water bath at about 50°C for half an hour, and worked up as the halides.

Formic acid and acetic acid were added pure (in 10% excess), and the reaction mixtures were kept at 50° for one hour at a pressure of 10 cm and then distilled at reduced pressure. In those cases where acetaldehyde was formed, it had to be removed as quickly as possible at reduced pressure.

Purification and analyses

Analytical samples of all the compounds were obtained by distillation with Vigreux columns of 50 cm x 12 mm. They were identified by NMR spectroscopy (Varian HA 100) in CCl₄ (±10% solutions), using hexa-methyl-disiloxane as external standard (see the table). In all the cases measured the differences with TMS as internal standard were 20 Hz to lower field.

Acknowledgement

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