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Spectra of water dimer from a new ab initio potential with flexible monomers

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We report the definition and testing of a new ab initio 12-dimensional potential for the water dimer with flexible monomers. Using our recent accurate CCpol-8s rigid water pair potential [W. Cencek, K. Szalewicz, C. Leforestier, R. van Harrevelt, and A. van der Avoird, Phys. Chem. Chem. Phys. 10, 4716 (2008)] as a reference for the undistorted monomers’ geometries, a distortion correction has been added, which was taken from a former flexible-monomer ab initio potential. This correction allows us to retrieve the correct binding energy $D_e = 21.0$ kJ mol$^{-1}$, and leads to an equilibrium geometry in close agreement with the one obtained from benchmark calculations. The kinetic energy operator describing the flexible-monomer water dimer has been expressed in terms of Radau coordinates for each monomer and a recent general cluster polyspherical formulation describing their relative motions. Within this formulation, an adiabatic scheme has been invoked in order to decouple fast (intramolecular) modes and slow (intermolecular) ones. Different levels of approximation were tested, which differ in the way in which the residual potential coupling between the intramolecular modes located on different monomers and the dependence of the monomer rotational constants on the dimer geometry are handled. Accurate calculations of the vibration-rotation-tunneling levels of (H$_2$O)$_2$ and (D$_2$O)$_2$ were performed, which show the best agreement with experiments so far for any water potential. Intramolecular excitations of the two monomers were calculated within two limiting cases, to account for the lack of non-adiabatic coupling between intramolecular modes due to the intermolecular motion. In the first model, the excitation was assumed to stay either on the donor or the acceptor molecule, and to hop between the two moieties upon donor-acceptor interchange. In the second model, the excitation remains on the same molecule whatever is the dimer geometry. Marginal frequency differences, less than 2 cm$^{-1}$, were obtained for all modes, and the resulting infrared shifts are in good agreement with experiments. © 2012 American Institute of Physics.

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I. INTRODUCTION

Water dimer is the subject of continuous theoretical and experimental interest.1 This is due to both the importance of the isolated dimer, in particular for atmospheric science,2,3 and to the fact that the dimer interactions constitute the leading and most important term in a many-body expansion for water clusters and condensed phases of water.1,4 The main source of experimental information on water dimer comes from spectral measurements.5–9 The quest to reproduce the intermolecular rovibrational transitions began in the late 1990s when six-dimensional (6D) “exact” quantum nuclear dynamics calculations became possible.10–13 In such calculations, the monomers are kept rigid, preferably at their average ground rovibrational state configuration.14,15 The vibrations and hindered rotations of the monomers in the water dimer have large amplitudes and are highly anharmonic, so that nearly exact quantum dynamics is absolutely necessary to quantitatively describe the transitions between the intermolecular rovibrational levels. In addition, these levels are split by quantum tunneling and the harmonic approximation fails to give even a qualitatively correct spectrum. The standard methods improving upon this approximation predict transitions with large errors. It has been found in the early calculations that even if fully anharmonic quantum dynamics is used, none of the existing water dimer potentials gives even a qualitatively correct description of the measured spectrum. However, it was possible to fit a potential to spectral data and to model the spectrum well.11 The first ab initio potential to achieve a fairly accurate description of the spectrum was the SAPT-5s potential developed in Refs. 13, 16, and 17. With two recent water pair potentials,18,19 it became possible to reach nearly perfect agreement of the intermolecular mode frequencies and tunneling splittings with the measured high-resolution spectra.

The next challenge in investigations of water dimer was to go beyond the rigid-monomer approximation. This extension allows in particular the computation of the shifts of intra-monomer fundamental infrared (IR) vibrational modes.

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transitions. These shifts are subjects of numerous measurements for hydrogen-bonded dimers and constitute a signature of the presence of hydrogen bonds. Although some of the IR shifts in the water dimer have been measured more than two decades ago, new experimental data are still becoming available and part of the existing data have been reassigned. The agreement between the previously calculated frequency shifts and the experimental data is still not as good as it is for the intermolecular modes. Whereas “exact” 12-dimensional (12D) calculations are still not possible, a [6+6]D adiabatic decoupling method developed by one of the present authors and co-workers is sufficiently accurate for the water dimer. This method was initially used to fit an empirical potential to the spectra. The only flexible-monomer water-dimer potentials available until the early 2000s were empirical ones fitted to bulk water properties with the inter-intramolecular couplings achieved only by the “atom-following” approach in atom-atom potential functional forms, i.e., by allowing the monomers to deform without changing the interatomic parameters. Such an approach gives only a qualitative description of the couplings. The first fully ab initio 12-dimensional water pair potentials were developed in Refs. 25 and 26. The potential in Ref. 25 was based on interaction energies by numerical differentiation. The SAPT-5s potential was obtained from a refitting of the rigid-monomer calculations were very similar to those from rigid-monomer ones, which shows that the effects of monomer flexibility on these transitions are small. The predictions from flexible-monomer calculations were very similar to those from rigid-monomer ones, which shows that the effects of monomer flexibility on these transitions are small.

In the present paper, we combine the currently most accurate rigid-monomer water dimer potential, CCpol-8s, with SAPT-5s/fIR. Such a combination is possible since SAPT-5s/fIR was constructed in such a way that if both monomers are in their average ground vibrational state geometries, SAPT-5s/fIR becomes identical with the SAPT-5s’ rigid-monomer potential. Thus, if one subtracts SAPT-5s’ and adds CCpol-8s, such a potential extends CCpol-8s to 12 dimensions by including monomer deformation effects at the level of SAPT-5s/fIR. The new potential has been used by us to compute the spectra of water dimer. The outline of this paper is as follows. In Sec. II, we briefly recall our previous high accuracy rigid CCpol-8s and flexible-monomer SAPT-5s/fIR potentials which we combine here to define a new flexible-monomer potential. Section III introduces the adiabatic formulation which allows us to handle flexible monomers. Section IV presents the resulting microwave (MW) and far infrared spectra, as well as the infrared shifts, and comparison to experiments. Finally, some conclusions are drawn in Sec. V.

II. A FLEXIBLE CCPOL-8S POTENTIAL

Recently, Cencek et al. proposed an accurate analytical representation of the water pair potential, denoted CCpol-8s. This rigid-monomer potential was obtained from a refitting of previous supermolecular calculations that used second-order perturbation theory with the Moller-Plesset decomposition of the Hamiltonian in combination with the CCSD(T) method. The functional form of this potential is given by

\[ V^{(2)} = \sum_{a \in A} \sum_{b \in B} \mu_{ab}^\text{ind} \cdot F_{ab} + V_2^\text{ind}(A, B), \]  

(1)

where

\[ V_2^\text{ind}(A, B) = -\frac{1}{2} \left\{ \mu_A^\text{ind} \cdot F_A + \mu_B^\text{ind} \cdot F_B \right\} \]  

(2)

represents the induction term with the induced dipole moments \( \mu_A^\text{ind} \) and \( \mu_B^\text{ind} \) and electric fields \( F_A \) and \( F_B \) on the monomers. The fields are produced by only the static (permanent) fractional point charges. In contrast, the induced dipole moments are fully iterated, i.e., are proportional to the electric fields produced by both the static charges and the induced
dipole on the interacting partner. The site-site potentials $u_{ab}$ depend only on the distances $r_{ab}$ and have the following general form:

\begin{equation}
  u_{ab}(r_{ab}) = f_1(\Delta^{(1)}_{ab}, r_{ab}) + \sum_{m=0}^{3} \frac{f_2(\Delta^{(m)}_{ab}, r_{ab})}{r_{ab}} + e^{-\beta_{0} r_{ab}} \sum_{m=0}^{3} \frac{f_3(\Delta^{(m)}_{ab}, r_{ab})}{r_{ab}^{m}},
\end{equation}

with $f_i(\delta, r)$ representing Tang-Toennies damping functions. The improvement in the relative mean-square error of the fit with respect to the ab initio data decreased from 0.089 to 0.010 kcal/mol for the data points with negative interaction energies. This new rigid-monomer potential, based on vibrationally averaged monomer geometries, produced a dimer vibration-rotation-tunneling (VRT) spectrum in substantially better agreement with high-resolution experimental data than the already quite accurate one predicted by CCpol-5s. Although CCpol-8s is close to the limit that might be expected for a rigid monomer potential, the lack of monomer flexibility precludes consideration of the IR frequency shifts of the monomer vibrations which constitute a signature of hydrogen bonding in the water dimer. Also other properties, such as the second virial coefficient, are sensitive to flexibility as it changes the binding energy $D_e$.

Previously, we obtained a flexible-monomer water dimer potential from ab initio calculations performed with SAPT at about a quarter million grid points with non-rigid monomers. A flexible-monomer 12-dimensional potential, denoted SAPT-5s/IR, was fitted to these points by means of a similar site-site form as in Eqs. (1)–(3), but using only five nonequivalent sites on each monomer and a perturbative (non-iterated) induction term

\begin{equation}
  V_2^{\text{ind}}(A, B) = -\frac{1}{2} \left[ \tilde{\alpha}^A \mathcal{F}^A \right]^2 + \tilde{\alpha}^B \mathcal{F}^B \right]^2 f_0(\delta^\text{ref}_q, R),
\end{equation}

with $\tilde{\alpha}^A$ and $\tilde{\alpha}^B$ being the molecular polarizabilities and $R$ the distance between polarization centers. This induction term contains a similar type of damping function as used in the other long range terms, while the CCpol-8s potential has an undamped but fully iterated induction term. To account for the energy variation upon distortion of the monomers, all the parameters $(q_a, q_b, \beta_{ab}, c_{ab}^m, C_{ab}^{(n)})$ entering Eqs. (1), (3) and (4) were made dependent on the distortion $\mathbf{q}^x - q^\text{ref}$ of each monomer. The symbol $\mathbf{q}^x$ stands for the set of three internal coordinates of monomer $X(A, B)$, and $q^\text{ref}$ refers to a reference geometry, for which we take the vibrationally averaged structure. The quality of this flexible-monomer potential surface can be checked, among other things, by comparing the characteristics of the global minimum with the benchmark calculations of Tschumper et al. Table I shows (the dimer interatomic distances and angles are defined in Fig. 1) that intermonomer parameters are reproduced by SAPT-5s/IR with a few-percent errors: the binding energy $D_e$ with 5.2%, the equilibrium distance $O_a - H_b$ with 3.0%, and the angle $\angle HO_aO_f$ with 1.7% error. It was demonstrated by Bukowski et al. that deficiencies of SAPT-5s compared to CCpol-5s result mainly from the smaller basis set used to develop the former potential.

In order to add flexibility to the recent very accurate CCpol-8s potential, we used the following simple prescription to define the pair potential $V^{(2)}_{\text{CCpol-8s}}$:

\begin{equation}
  V^{(2)}_{\text{CCpol-8s}}(\mathbf{d}) \equiv V^{(2)}_{\text{CCpol-8s}}(\mathbf{u}) + V^{(2)}_{\text{SAPT-5s/IR}}(\mathbf{d}) - V^{(2)}_{\text{SAPT-5s}}(\mathbf{u}),
\end{equation}

where $\mathbf{d}$ and $\mathbf{u}$ refer to geometries with distorted and the corresponding undistorted monomers, respectively. The latter implies that both monomers are at their reference geometry $q^\text{ref}$. Merging the two potentials as indicated above is actually possible because the $q^\text{ref}$ geometry defined in SAPT-5s/IR corresponds to the same vibrationally averaged H$_2$O geometry as used for the rigid monomers in CCpol-8s. The definition given above relies on the assumption that the intrinsic errors of SAPT-5s/IR, due to the lower accuracy level of the method, are similar for distorted and undistorted geometries. The difference $V_{\text{SAPT-5s/IR}}(\mathbf{d}) - V_{\text{SAPT-5s}}(\mathbf{u})$ thus represents the flexibility correction to be applied to the $V_{\text{CCpol-8s}}(\mathbf{u})$ potential obtained for rigid monomers.

In the bound-state calculations, the 12D geometry of the system is described by two sets of coordinates: (i) the intermolecular coordinates, denoted by $Q = (\mathbf{R}, \Omega^A, \Omega^B)$, which determine the center-of-mass separation $\mathbf{R}$ and the relative orientations of the two monomer-fixed frames, and (ii) the intramolecular ones $(\mathbf{q}^A, \mathbf{q}^B)$ that specify the geometry of water.
each monomer. These two sets of coordinates are not independent as the Euler angles \( \Omega \) that determine the orientation of monomer \( X \) with respect to a dimer frame depend on the embedding of the monomer frame in the (vibrationally distorted) monomer. The SAPT-5s/IR potential depends only on the Cartesian coordinates of all atoms which uniquely define the 12D geometry of the system. This 12D geometry can then be represented by a set of inter- and intra-monomer coordinates based on arbitrary embedding. Thus, we first associate with any 12D distorted geometry (\( \mathbf{d} \)), specified by Cartesian coordinates of all atoms, the set \((q^A, q^B, Q)\) based on some assumed embedding. The undistorted geometry is then \((\mathbf{u}) \equiv (\mathbf{q}_{\text{ref}}, \mathbf{q}_{\text{ref}}, \mathbf{Q})\), where \( \mathbf{q}_{\text{ref}} \) corresponds to the vibrationally averaged geometry mentioned before. Equation (5) can thus be rewritten more explicitly as

\[
V_{\text{CCpol-8s}/fIR}^{(2)}(q^A, q^B, Q) = V_{\text{CCpol-8s}}^{(2)}(Q) + V_{\text{SAPT-5s}/IR}^{(2)}(q^A, q^B, Q) - V_{\text{SAPT-5s}}^{(2)}(Q) .
\]

Note that \( V_{\text{SAPT-5s}}^{(2)}(Q) = V_{\text{SAPT-5s}/IR}(\mathbf{q}_{\text{ref}}, \mathbf{q}_{\text{ref}}, Q) \). As the above definition refers to the two-body term, it must be supplemented by the one-body terms \( V^{(1)}(q^A) \) to obtain the full potential energy of the water dimer

\[
V_{\text{CCpol-8s}/fIR}^{(2)}(q^A, q^B, Q) = V^{(1)}(q^A) + V^{(1)}(q^B) + V_{\text{CCpol-8s}/fIR}^{(2)}(q^A, q^B, Q) .
\]

For the one-body terms, we used the spectroscopically accurate water PJT2 potential of Polansky et al.\(^{30}\) It should be noted that these one-body terms were already included in the fitting of the flexible-monomer SAPT-5s/IR potential to the set of \( ab \text{initio} \) points.

Table I shows that this new flexible-monomer potential gives significantly better characteristics of the global minimum than SAPT-5s/IR. In particular, the 5.2% error in the binding energy \( D_e \) is reduced to zero and the errors of inter-monomer coordinates are all below 1%.

### III. VRT LEVELS FROM A FLEXIBLE-MONOMER POTENTIAL

In this section, we first define the coordinates used to describe the fully flexible water dimer and the resulting Hamilton operator. We then explain the adiabatic decoupling between the fast (intramolecular) and slow (intermolecular) modes, and the different approximations which can be invoked depending on the level of accuracy needed. In particular, we will consider how the time-consuming calculation of the adiabatic potential can be drastically accelerated by neglecting a part of the potential coupling between the intramolecular modes located on different monomers. We then discuss two limiting cases for the definition of intramolecular excitations of the two monomers to account for the breakdown of the adiabatic separation between the intramolecular modes on different monomers.

Finally, it is shown that this adiabatic scheme allows us to recast the calculation into a pseudo-rigid monomer one using an adiabatic potential energy surface, except for the rotational constants which depend on the intermolecular geometry.

#### A. Full flexible-monomer Hamiltonian

As formulated in a previous paper,\(^{22}\) the exact Hamiltonian for the fully flexible water dimer is

\[
\hat{H} = -\frac{\hbar^2}{2\mu_{AB} R^2} \frac{\partial^2}{\partial R^2} R + \hat{T}^{A}_{VR} + \hat{T}^{B}_{VR} + V(q^A, q^B, Q) + \frac{1}{2\mu_{AB} R^2} (\hat{J}^2 + 2\hat{J}_{AB}^2 - 2J_{AB} \cdot \hat{J})
\]

where \( J_{AB} = J_A + J_B \) with \( J_A \) and \( J_B \) the internal rotation angular momenta of the monomers, \( J = J_{AB} + \hat{L} \) is the total angular momentum with \( \hat{L} \) being the end-over-end angular momentum of the monomers’ centers of mass and \( \mu_{AB} \) is the reduced mass of the dimer. The operator \( \hat{T}^{X}_{VR} \) represents the vibration-rotation kinetic energy operator (KEO) of monomer \( X \), to be specified in Sec. III C. In order to proceed with the derivations, we re-express this KEO in matrix form\(^{37}\) (suppressing the \( X \) label)

\[
\hat{T}_{VR} = \frac{1}{2} [\hat{p}^j] \begin{pmatrix} \Sigma & \sigma \end{pmatrix}[\sigma^t \Gamma][\hat{p}^j],
\]

with \( \hat{p} = (\hat{p}_1, \hat{p}_2, \hat{p}_3) \) being the momenta conjugate to the internal coordinates \((q_1, q_2, q_3)\) and \( \hat{J} = (\hat{J}_x, \hat{J}_y, \hat{J}_z) \) the rotational angular momentum components. The matrices \( \Sigma, \Gamma, \) and \( \sigma \) are defined in Sec. III C.

An exact quantum calculation explicitly treating these 12 internal degrees of freedom would probably be possible nowadays, due to the huge computing power of massively parallel machines. However, one can resort to an adiabatic separation between the “fast” intramolecular coordinates \((q^A, q^B)\) and the “slow” intermolecular coordinates \( Q \), as applied before by Klopper, Quack, and Suhm\(^{38,39}\) in a \([4 + 2]D\) treatment of the HF dimer, and more recently by one of us and co-workers for the water dimer.\(^{22}\)

#### B. Adiabatic approximation

To define here such a \([6+6]D\) formulation, i.e., to decouple the intramolecular coordinates from the intermolecular ones, we make two basic approximations:

(i) we set the intramolecular vibration-rotation coupling matrix \( \sigma \) to zero, so that

\[
\hat{T}_{VR} \simeq \hat{T}^{(ad)}_{VR} = \hat{T}^{A}_{VR} + \hat{T}^{B}_{VR},
\]

which will be justified later on. This allows a partitioning of the total Hamiltonian according to

\[
\hat{H}^{(ad)} = \hat{H}_{\text{inter}} + \hat{H}_{\text{intra}}(Q)
\]

with

\[
\hat{H}_{\text{inter}} = \hat{T}_R + \hat{T}^{A}_{R} + \hat{T}^{B}_{R} + \hat{T}_{CC},
\]

\[
\hat{H}_{\text{intra}} = \hat{T}^{(ad)}_{VR} + \hat{T}_{CC},
\]

\[
\hat{T}_{R} = \frac{1}{2} \frac{\partial^2}{\partial q^2} + \frac{1}{2} \frac{\partial^2}{\partial q^2} + \frac{1}{2} \frac{\partial^2}{\partial q^2}.
\]
where $\hat{T}_R$ stands for the first term of Eq. (8) and $\hat{T}_{CC}$ (centrifugal and Coriolis coupling) for the last term in this equation.

(ii) the $\mathbf{Q}$-dependence of the intramolecular modes is treated in the “clamped $\mathbf{Q}$” approximation, nonadiabatic coupling terms are neglected. The calculations of the intramolecular modes proceed as presented below.

C. 6D intramolecular calculations

The intramolecular (adiabatic) eigenfunctions $\Phi_N(q^A, q^B; Q)$ at fixed $Q$ are obtained from the eigenvalue problem

$$\{\hat{T}_{V}^{(ad)} + \hat{T}_{V}(\mathbf{q}^A, \mathbf{q}^B; \mathbf{Q})\} \Phi_N(q^A, q^B; Q) = V_N^{(ad)}(Q) \Phi_N(q^A, q^B; Q),$$

with the intramolecular Hamiltonian of Eq. (13). The subscript $N$ is a composite index representing the six vibrational quantum numbers associated with the two monomers, as will be discussed in Sec. III D.

For efficient convergence, we express the $\hat{T}_{V}^{(ad)}$ operators in terms of Radau’s coordinates $(r_1, r_2, \beta)$ (Ref. 40)

$$\hat{T}_{V}^{(ad)} = \frac{1}{2} \hat{p}^\dagger \Sigma \hat{p},$$

$$= \frac{1}{2} [\hat{p}_r_1, \hat{p}_r_2, \hat{p}_\beta] \begin{pmatrix} \frac{1}{m_H} & 0 & 0 \\ 0 & \frac{1}{m_H} & 0 \\ 0 & 0 & \frac{1}{m_H} \end{pmatrix} \begin{pmatrix} \hat{p}_r_1 \\ \hat{p}_r_2 \\ \hat{p}_\beta \end{pmatrix},$$

(15)

where $m_H$ is the mass of hydrogen atom which are ideally suited to the H$_2$O molecule as shown by Bačić et al.41 We use the $f$-embedding formulation of Wei and Carrington,42,43 with the value of the $f$ parameter equal to 1, i.e., choosing the $z$ axis of each monomer frame as the vector that bisects the angle between the two Radau vectors. This results in a $\Gamma$

$$\hat{H}_{\text{int}}(\mathbf{Q}) = \hat{T}_{V}^{(ad)} + \hat{T}_{V}(\mathbf{q}^A, \mathbf{q}^B; Q),$$

(13)

where $c = \cos \beta$. The resulting Coriolis matrix $\sigma$ entering Eq. (9) has only zero elements except for

$$\sigma_{\beta \xi} = \sigma_{\xi \beta} = \frac{1}{2m_H} \sqrt{1 - c^2} \left[ \frac{1}{r_1^2} - \frac{1}{r_2^2} \right],$$

(17)

which vanish at equilibrium ($r_1 = r_2$), and stay very small except for highly distorted monomer geometries. In the work presented here, this term was neglected, which allowed for the decoupling of the intermolecular calculations from the intramolecular ones. Actually, this intramolecular Coriolis coupling gives rise to the terms

$$\hat{p}_\beta^\dagger \sigma_{\beta \xi} \hat{j}_\xi + \hat{j}_\xi^\dagger \sigma_{\xi \beta} \hat{p}_\beta, $$

which would lead to a correction that vanishes in first-order perturbation theory. This is easily shown by expressing $\hat{j}_\xi$ in terms of ladder operators $\hat{j}_\xi = \frac{1}{2}(\hat{j}_+ + \hat{j}_-)$ that have vanishing diagonal matrix elements in the asymmetric rotor eigenfunction basis of the H$_2$O monomers (see Sec. III E).

The adiabatic intermolecular potential $V_N^{(ad)}(Q)$ results from a six-dimensional calculation of Eq. (14), to be performed for every six-dimensional intermolecular geometry ($\approx 10^6$ $Q$-points). In order to make these calculations efficient, we used the following two-step procedure at each geometry $Q$:

(i) We first optimize the intramolecular geometry $(q^A_{\text{opt}}, q^B_{\text{opt}})$ for a given geometry $Q$, so that

$$\frac{\partial V(q^A_{\text{opt}}, q^B_{\text{opt}}; Q)}{\partial q^X} \bigg|_{q^X = q^X_{\text{opt}}} = 0 \quad (X = A, B).$$

(18)

We use the Powell method44 which does not require analytic energy derivatives. The associated potential energy is $V_{\text{opt}}(Q) = V(q^A_{\text{opt}}, q^B_{\text{opt}}; Q)$.

(ii) We then expand the intramolecular potential around the minimum

$$V(q^A, q^B; Q) = V(q^A_{\text{opt}}, q^B_{\text{opt}}; Q)$$

$$+ V(q^A_{\text{opt}}, q^B_{\text{opt}}; Q) - V_{\text{opt}}(Q)$$

$$+ V(q^A_{\text{opt}}, q^B_{\text{opt}}; Q) - V_{\text{opt}}(Q)$$

$$+ V(q^A, q^B; Q) - V^A(q^A_{\text{opt}}, q^B; Q) - V^B(q^A; Q) - V_{\text{opt}}(Q)$$

$$\equiv V^{AB}(q^A, q^B; Q).$$

(19)
which allows one to define single-monomer terms $V^X(q^X; Q)$ at fixed optimized geometry of the other monomer and a residual coupling term $V^{AB}(q^A, q^B; Q)$ between the two monomers.

(iii) The vibrational states of monomer $X$ are exactly computed in the field of the other monomer frozen at its optimized geometry, by means of a sequential truncation-reduction scheme\textsuperscript{25,46}

$$\{ \dot{\psi}^X_{n_3}(q^X; Q) \} = \frac{1}{\sqrt{\rho X(q^X; Q)}} \left( \frac{\partial}{\partial q^X} - V^X(q^X; Q) \right) \psi_{n_3}(q^X; Q) = E^X_{n_3}(Q) \psi_{n_3}(q^X; Q) .$$

The eigenvalue $E^X_{n_3}(Q)$ represents the energy of mode $n_3$ with respect to the instantaneous energy minimum. The six-dimensional zero-order intramolecular states $\Phi^{(0)}_A(q^A, q^B; Q)$ are then defined as

$$\Phi^{(0)}_A(q^A, q^B; Q) = \psi_{n_3}(q^A; Q) \psi_{n_3}(q^B; Q) ,$$

with associated energies

$$V^{[3+3]D}_N(Q) = V_{opt}(Q) + E^A_{n_3}(Q) + E^B_{n_3}(Q) - 2E_0^\infty ,$$

where $E_0^\infty$ is the zero-point energy of an isolated monomer. This approach corresponds to the $[3+3]D$ adiabatic formulation in which one ignores the potential coupling between intramolecular modes. The ground state energy $V^{[3+3]D}_0(Q)$ reflects the change in the zero-point energies of both monomers when the intermolecular geometry $Q$ varies. With this definition, the dissociation limit $V^{[3+3]D}_0(R \to \infty)$ equals zero provided that $\lim_{R \to \infty} V_{opt}(Q) = 0$.

(iv) The remaining term $V^{AB}(q^A, q^B; Q)$ can then be taken into account either by perturbation theory

$$V^{6Dp}_N(Q) = V_{opt}(Q) + E^A_{n_3}(Q) + E^B_{n_3}(Q) - 2E_0^\infty + \langle \psi_{n_3}(Q) | \psi_{n_3}(Q) \rangle V^{AB} \langle \psi_{n_3}(Q) | \psi_{n_3}(Q) \rangle ,$$

which we call the $6Dp$ adiabatic formulation, or within a full variational treatment in which the potentials $V^{6Dp}_N(Q)$ are defined as the eigenvalues of the Hamiltonian matrix

$$\hat{H}_{n_3n_3n_3} = \left[ V_{opt}(Q) + E^A_{n_3}(Q) + E^B_{n_3}(Q) - 2E_0^\infty \right] \delta_{n_3n_3} \delta_{n_3n_3} + \langle \psi_{n_3}(Q) | \psi_{n_3}(Q) \rangle V^{AB} \langle \psi_{n_3}(Q) | \psi_{n_3}(Q) \rangle ,$$

which is named the $6D$ adiabatic formulation.

These three different definitions of the adiabatic potential will be compared further on.

D. Excited intramolecular states

In our $[6+6]D$ adiabatic model for the vibrational ground state of the monomers, the $6D$ adiabatic intermolecular potential has eight equivalent minima. Each of these minima corresponds to a hydrogen-bonded structure in which one of the monomers is the donor in the hydrogen bond and the other monomer is the acceptor. In four of the minima, monomer $A$ is the donor and $B$ the acceptor, in the other four minima the monomers are interchanged, so that $B$ is the donor and $A$ the acceptor. The intermolecular VRT states calculated on the ground state potential also correspond to hydrogen-bonded structures; they are adapted to the irreducible representations (irreps) of the full permutation-inversion group $G_{16}$. The states that carry the one-dimensional irreps of $G_{16}$ are either symmetric or antisymmetric with respect to the monomer interchange operator $P_{AB}$, which leads to a small ($\approx 1 \text{ cm}^{-1}$) interchange tunneling splitting of the intermolecular rovibrational levels. The two components of the states that carry the two-dimensional $G_{16}$ irreps are interchanged by the action of $P_{AB}$ and the corresponding energy levels are not split by interchange tunneling.

If one of the monomers is vibrationally excited, the situation is more complicated. If the excitation is localized on one of the monomers, say $A$, the corresponding adiabatic intermolecular potential is no longer symmetric under donor-acceptor interchange $P_{AB}$. An equivalent adiabatic state exists with monomer $B$ excited, which is not symmetric under $P_{AB}$ either, but which may be obtained from the adiabatic state in which $A$ is excited by acting on it with $P_{AB}$. The interchange tunneling splitting in the monomer excited state can be calculated from a two-state model in which the two equivalent adiabatic potentials and the non-adiabatic coupling are included, as it was done, for example, by Fraser.\textsuperscript{47} The excited state interchange tunneling splitting is smaller by an order of magnitude than the corresponding ground state splitting\textsuperscript{5} because the monomer excitation is localized either on the donor or on the acceptor, and donor-acceptor interchange also involves hopping of the excitation between the monomers $A$ and $B$. So, it is clear that our adiabatic $[6+6]D$ model would have to be extended to compute reliable values of the excited state interchange tunneling splitting. This is not a goal of the present paper, however. Here, we perform monomer excited state calculations to obtain infrared frequency shifts of the monomer vibrations in the dimer. In a previous paper,\textsuperscript{34} two of us discussed and applied two approximate methods to do this, while maintaining the $[6+6]D$ adiabatic model.

(i) One option is to identify the donor (D) and the acceptor (A) monomers at each intermolecular geometry $Q$, and to define donor- or acceptor-excited adiabatic potentials. The acceptor can, for example, be identified as the monomer displaying the shortest $d_{O-H}$ distance between its oxygen atom and a hydrogen atom located on the other monomer. Such a definition retains the full $G_{16}$ symmetry of the adiabatic intermolecular potentials, and we will therefore call it the $G_{16}$ model. It implies that upon donor-acceptor interchange (which corresponds to bringing the complex from one minimum to an equivalent one) the excitation energy simultaneously migrates from one monomer to the other. The interchange tunneling splitting that would result from this model for the excited state would be roughly of the same size as the ground state interchange tunneling splitting, in contrast with high-resolution spectra\textsuperscript{5} which show that the excited state tunneling splitting is reduced by one
order of magnitude with respect to the ground state splitting.

(ii) A second possibility is to stipulate that, whatever is the
intermolecular geometry $Q$, the excitation energy stays
on the same monomer, say A. In that case, the excitation
does not migrate upon donor-acceptor interchange.
The symmetry of the adiabatic potentials is broken be-
cause the minima at geometries $Q_{eq}$ differ slightly in en-
ergy from those at $P_{AB}Q_{eq}$. As the interchange $P_{AB}$ is no
longer a symmetry operation, the $G_{16}$ symmetry of the
intermolecular VRT states is lost and one must use the
subgroup $G_8$ instead. If this scheme, which we call the $G_8$
model, were used to compute excited state interchange
tunneling splittings, the resulting splittings would be too
small.

In Sec. IV, we will present results corresponding to these
two extreme possibilities, and show that they produce nearly
identical infrared frequency shifts of the monomer vibrations
in the water dimer.

E. 6D intermolecular calculations

The full dimer wave function $\psi_N(q^A, q^B, Q)$ is a solution
of the equation

$$\left(\hat{H}_{\text{inter}} + \hat{H}_{\text{intr}(Q)}\right)\psi_N = \mathcal{E}^N \psi_N,$$

with the operators $\hat{H}_{\text{inter}}$ and $\hat{H}_{\text{intr}(Q)}$ defined in
Eqs. (12) and (13). Within the adiabatic approximation,
the full dimer wave function is written as

$$\psi_N(q^A, q^B, Q) = \phi_N(q^A, q^B; Q) \psi^N(Q).$$

Neglecting all non-adiabatic coupling terms, the Schrödinger
equation to be solved in the coordinates $(q^A, q^B, Q)$ thus
reads as

$$\left[\hat{T}_R + \hat{T}^{(\text{ad})}_R + \hat{T}^{(\text{ad})}_B + \hat{T}_{CC} + V^{(\text{ad})}_N(Q)\right] \phi_N \psi^N

= \mathcal{E}^N \phi_N \psi^N.$$

The operators that differentiate with respect to the intermolecu-
lar coordinates $Q$ act only on $\psi^N$ and not on the function
$\phi_N$, as implied by the adiabatic approximation. The adiabatic
potential $V^{(\text{ad})}_N(Q)$ in this equation was obtained by solving
the eigenvalue problem, Eq. (14), in the intramolecular coordi-
nates $q^A, q^B$ with clamped $Q$. Projecting Eq. (27) onto
the adiabatic eigenfunction $\phi_N(q^A, q^B; Q)$, one obtains the inter-
atomic equation

$$\left[\hat{T}_R + \hat{T}^{(\text{ad})}_R + \hat{T}^{(\text{ad})}_B + \hat{T}_{CC} + V^{(\text{ad})}_N(Q)\right] \psi^N(Q) = \mathcal{E}^N \psi^N(Q).$$

(28)

The notation $\hat{T}^{(\text{ad})}_R$, with the operators $\hat{T}_R^{(\text{ad})}$ defined in
Eqs. (9) and (10), implies averaging over the adiabatic monomer
states $\phi_N(q^A, q^B; Q)$, i.e., integration over the
monomer coordinates $q \equiv (q^A, q^B)$

$$\hat{T}^{(\text{ad})}_R = \frac{1}{2} \sum_{j} \left| j^X \right| \left\langle \phi_N | \Gamma^X \phi_N \right\rangle q j^X \equiv \frac{1}{2} \sum_{j} \left| j^X \right| \Gamma^X j^X.$$ (29)

In Sec. IV, we will also present calculations in which this term
was obtained without averaging, but instead evaluated for the
optimized intramolecular coordinates $q_{\text{opt}}$, see Eq. (18),

$$\hat{T}^{(\text{ad})}_R = \frac{1}{2} \sum_{j} \left| j^X \right| \Gamma^X j^X,$$

(30)

with $\Gamma^X(q_{\text{opt}}) = \Gamma^X(q_{\text{opt}})$. Such term is denoted as “opti-
mized” in contrast to the “averaged” version of Eq. (29). In
both cases, the operator $\hat{T}_R$ contains off-diagonal terms, cf.
Eqs. (9) and (16)

$$\Gamma_{\gamma} \{ j^A_\gamma, j^B_\gamma \}.$$ (31)

These terms will be neglected as $\Gamma_{\gamma}$ is much smaller than the
diagonal elements in Eq. (16) when the monomer geometries
are not too strongly distorted and $r_1 \approx r_2$. Moreover, their
contribution vanishes in first-order perturbation theory due to the
fact that $\bar{j}_\gamma$ only contains step-up and step-down operators.

Equation (28) formally corresponds to the formulation of
Brocks et al.48 for rigid monomers, but it uses an adiabatic po-
tential which reflects the variation of the monomer zero-point
energies as a function of the intermolecular geometry $Q$. A
second, important difference stems from the $Q$-dependence of
the rotational constant matrices $\Gamma^X(N)$ of the monomers.
The coordinates $(r_1, r_2, \cos \beta)$ entering Eq. (16) correspond to ei-
ther the optimized values $\phi_{\text{opt}}(Q)$ given by Eq. (18) or they
are averaged over $Q$-dependent adiabatic monomer eigen-
functions as shown in Eq. (29). Explicitly handling these $Q$
dependent terms makes the energy levels calculation about
one order of magnitude more expensive than in the rigid case.

Therefore, we have investigated the approximation of making the $\Gamma^X(N)$ (or $\Gamma^X(q_{\text{opt}})$) matrices independent of the Euler angles
$\Omega^A$ and $\Omega^B$ through the averaging procedure

$$\tilde{\Gamma}^X(N)(R) = \int d\Omega^A \int d\Omega^B \Gamma^X(N)(R, \Omega^A, \Omega^B)$$

$$\times \exp \left\{ -\alpha \left[ V^{(\text{ad})}_N(R, \Omega^A, \Omega^B) - V_{\text{min}} \right] \right\}.$$ (32)

while explicitly retaining the $R$-dependence. In Eq. (32), $V_{\text{min}}$
is the minimum value of the potential and $\alpha$ is some constant
($\approx 10^5$ Hartree$^{-1}$). This approximation allows us to recast
the fully flexible formulation into a rigid one, except for the rota-
tional matrices $\tilde{\Gamma}^X(N)$ which only depend on the separation $R$.

Within this averaging approximation, which yields

$$\tilde{T}^X(N)_R = \frac{1}{2} \sum_{j} \left| j^X \right| \tilde{\Gamma}^X(N)(R) j^X,$$ (33)

we will consider the two possibilities of either ignoring the
angular dependence of the $\Gamma^X(N)$ matrices

$$\left[ \hat{T}_R + \tilde{T}^{(\text{ad})}_R(N) + \tilde{T}^{(\text{ad})}_B + \hat{T}_{CC} + V^{(\text{ad})}_N(Q) \right] \psi^N(Q) = \mathcal{E}^N \psi^N(Q)$$

(34)

or subsequently retrieving this dependence with

$$\Delta \Gamma^X(N)(Q) = \Gamma^X(N)(Q) - \tilde{\Gamma}^X(N)(R)$$

(35)

by first order perturbation theory

$$\Delta \varepsilon^N \equiv \frac{1}{2} \langle \psi^N | \Delta \Gamma^{(A)} \left| j^A \right\rangle \Delta \Gamma^{(B)} \left| j^B \right\rangle | \psi^N \rangle.$$ (36)

We will show in Sec. IV that this latter correction essentially
gives the exact transition energies obtained from Eq. (28).
A detailed description of the implementation of the calculation of the intermolecular VRT states has been given previously. For the sake of clarity, we briefly recall it. The intermolecular KEO of Eq. (28) leads to simple matrix elements in the overall basis set

\[ \mathcal{B} = \{|n\rangle \otimes \{|j_A, k_A, m_A\rangle \otimes \{|j_B, k_B, m_B\rangle \otimes \{|J, K, M\rangle \rangle \rangle = (m_A + m_B = K), \] (37)

where \(|n\rangle \) is an appropriate basis in the interfractional distance coordinate \( R \), the Wigner functions \( \{ \Omega |j, k, m\rangle \) = \( \sqrt{2J + 1}/8\pi^2 D_{m}(\Omega) \) form a basis for the rotation of each monomer with the Euler angles \( \Omega = \{ \phi, \theta, \chi \} \), and \(|\{J, K, M\rangle \) is the Wigner basis associated with the overall rotation of the complex. The basis \( \mathcal{B} \) can be projected onto the different irreps \( \Gamma \) of the molecular symmetry group \( G_{16} \) (or \( G_{8} \)) used to label the energy levels \( \{ \nu \rangle \)

\[ \mathcal{B} = \bigoplus_\Gamma \{|n\rangle \otimes \{|\nu; \Gamma \rangle \}, \]

where \( \nu \) stands for the combination of all angular quantum numbers. The most compact representation for the adiabatic potential energy \( V_{\text{ad}}(R, \Omega^A, \Omega^B) \) is the six-dimensional direct-product grid \( \{ \phi \times \chi^A \times \chi^B \times \theta^A \times \theta^B \times R \} \), where the symbols \( \phi \), etc., denote the grids in the respective coordinates and the coordinate \( \phi = \phi^A - \phi^B \). This grid is restricted to points where the potential energy is lower than some threshold value \( V_{\text{max}} \), and only non-symmetry-equivalent points are computed, which typically corresponds to a dimension of about \( 10^6 \). As the six-dimensional intramolecular eigenvalue problem has to be solved for each point of this grid, this is an expensive step of the whole calculation. It is straightforwardly parallelized by means of the Open MP protocol. More specifically, the inner do-loop running over the \( \phi, \chi^A \), and \( \chi^B \) grid points was declared to run in parallel. On a 12-core processor, a speed up factor of 5.9 was obtained, close to the best performance expected due to the limited size of the common cache memory. Energy levels are computed from an iterative Lanczos procedure, parallelized on the outer do-loop running over the radial basis set index. A similar speed-up factor of 6.8 was obtained for this step. When needed, intermolecular eigenvectors \( |\psi^N_{\alpha}\rangle \) were computed afterwards from the previously stored Lanczos vectors \( |u_n\rangle \)

\[ \begin{align*}
|\psi^N_{\alpha}\rangle &= \sum_n T_{\alpha n}|u_n\rangle ,
\end{align*} \] (38)

where \( T \) corresponds to the eigenvector matrix resulting from the diagonalization of the tridiagonal Hamiltonian matrix built by the Lanczos scheme. Transformations between the spectral and grid representations, as required to act with the potential energy operator on the successive Lanczos vectors, are performed by means of a six-dimensional pseudospectral scheme involving a three-dimensional fast Fourier transform. Depending on the method, \([3+3]D, 6Dp, \) or \( 6D \), used to compute the adiabatic potential, the whole calculation will be termed \([6+[3+3]D], [6D+6Dp] \) or \([6+6]D \), respectively, as the intermolecular calculation is effectively six-dimensional.

**IV. RESULTS**

In this section, we present the results obtained from the new potential for the MW and far IR spectra, and the frequency shifts of the monomer vibrations observed in IR spectra of the \((H_2O)_2\) dimer, and compare with available experimental data. We also present results for the \((D_2O)_2\) dimer, for which accurate VRT spectra are available as well. First, we describe the basis set specifications used in the calculations.

**A. Specification of basis sets and convergence tests**

**1. Intramolecular basis set**

The intramolecular calculations used a local biharmonic basis set \( \{ H_m(r_1; \beta_1) \otimes H_n(r_2; \beta_2) \} \) on each monomer which was contracted at each sampled value of the grid for the Radau angle \( \beta \). The three-dimensional resulting eigenvalues and eigenvectors are then obtained by means of a sequential truncation-reduction scheme. Two different monomer intramolecular bases were considered in the calculations: the first basis \( (I) \) consisted of 7 harmonic functions along each radius \( r \) and 19 sampled angles \( \beta \). For the second basis \( (II) \), these numbers were increased to 9 and 25, respectively. Furthermore, for convergence checking only, we also performed test calculations with an enlarged basis \( (III) \) corresponding to 11 harmonic functions and 31 \( \beta \)-angle sampling points. The completeness of these bases is assessed in the upper part of Table II by comparing the calculated fundamental frequencies of the water molecule with experimental values.

The lower part of Table II reports the zero-point energy and fundamental intramolecular frequencies at the water dimer equilibrium geometry. The zero-point energies of both monomers, which actually define the ground state adiabatic potential, are practically the same in bases I and II. Basis set I has thus been used for the MW and far IR calculations.

<table>
<thead>
<tr>
<th>H2O</th>
<th>Basis I</th>
<th>Basis II</th>
<th>Basis III</th>
<th>Exp.</th>
</tr>
</thead>
<tbody>
<tr>
<td>b</td>
<td>1594.68</td>
<td>1594.67</td>
<td>1594.67</td>
<td>1594.7</td>
</tr>
<tr>
<td>ss</td>
<td>3658.29</td>
<td>3657.23</td>
<td>3657.19</td>
<td>3657.1</td>
</tr>
<tr>
<td>as</td>
<td>3755.99</td>
<td>3755.87</td>
<td>3755.86</td>
<td>3756.0</td>
</tr>
</tbody>
</table>

**TABLE II. IR excitation frequencies (in cm⁻¹), calculated for the isolated monomer and the water dimer equilibrium geometry, as a function of the intramolecular basis set selected:** b denotes the bend mode, ss the symmetric stretch, and aa the antisymmetric stretch. The upper part of the table refers to the isolated monomer, the lower one to the dimer equilibrium geometry: O–H₂[D] and O–H₂[D] correspond to the two donor OH stretches (see Fig. 1), zpe[X] to the zero-point energy of monomer X.


2. Intermolecular basis set

Convergence with respect to the intermolecular basis set, as in Eq. (37), was assessed by systematically increasing its size. For each intermolecular vibration considered in this work, convergence was tested on both the subfork positions and the associated splittings, as depicted in Fig. 2.

For the (H₂O)₂ dimer, we used a Wigner basis up to maximum values of $j_A$, $j_B = 11$ on the monomers, and a primitive radial basis set of 20 sine functions spanning the range [4.2 ≤ $R$ ≤ 10]Å, contracted to 9 functions by means of the Harris, Engerholm, and Gwinn procedure. A further increase in the Wigner basis, up to $j_A = j_B = 12$, leads to changes less than 0.08 (0.025) cm⁻¹ for the subfork positions (splittings). For the ground state results, these values are reduced to 0.06 (0.002) cm⁻¹.

Concerning the (D₂O)₂ dimer, a larger intermolecular basis (with maximum $j_A$, $j_B = 13$ values) had to be used in order to converge the tunneling splittings, which in (D₂O)₂ are smaller than in (H₂O)₂ by an order of magnitude. Increasing the basis to $j = 14$ changes the excited subfork positions (splittings) by less than 0.14 (0.01) cm⁻¹. These values are reduced to 0.06 (0.001) cm⁻¹ for the ground state.

The intramolecular calculations were systematically conducted in the 6D full variational approach, which takes into account the vibrational coupling between monomers and leads to the global [6+6]D formulation. We will, however, mention how the [6+[3+3]]D and [6D+6Dp] approximations compare with it.

2. Test of some approximations for flexible monomers

We have shown in Sec. III B that, within the adiabatic approximation, one can recast the 12-dimensional calculation for flexible monomers into a series of calculations of the intramolecular vibrations on a grid of intermolecular coordinates $Q$, which yields a set of adiabatic potentials, and a pseudo-rigid monomer calculation of the intermolecular VRT states on one of these adiabatic potentials. However, the resulting $Q$-dependence of the monomer rotational constant matrices $\Gamma$ renders the Lanczos iterative diagonalization scheme about one order of magnitude more costly. The reason is that one has to switch from the spectral to the grid representation many times in order to evaluate the effect of the monomer
rotation operators $\hat{R}^X_R$. The perturbative approach defined by Eqs. (32)–(36) allows one to retrieve the exact transition energies and splittings to within 0.03%. If we ignore the final perturbation step in Eq. (36), the errors increase to about 1%.

The second series of tests concerns the calculation of the adiabatic potential $V^{(ad)}_V(Q)$ as detailed in Sec. III C. We compared the results for the MW and far IR transitions obtained, respectively, in the [6+[3+3]]D approximation, i.e., completely neglecting the six-dimensional $V^{AR}(q^4, q^R; Q)$ correction term of Eq. (19), and in the full variational treatment [6+6]D of this term by means of Eq. (24).

We found that, for excited states, the transition energies $o_i(K)$ are lower in the [6+6]D treatment than in the [6+[3+3]]D approximation, but the decrease does not exceed 1 cm$^{-1}$. The associated interchange splittings $i_i(K)$ differ by 0.5 cm$^{-1}$ at most. For the ground state, these values are 0.25 cm$^{-1}$ and 0.036 cm$^{-1}$, respectively. The calculations presented in the following were systematically conducted in the full variational [6+6]D treatment. The results corresponding to the [6D + 6Dp] perturbative approach, not shown here, are essentially indistinguishable from those obtained within the [6+6]D variational one.

Finally, we compare the results obtained from either $q$-averaging the rotational constant matrices $\Gamma^X(0)$ as given by Eq. (29), or using their optimized values $\Gamma^X(\text{opt})$ from Eq. (30). It should be kept in mind that the adiabatic potentials are identical in both cases, only the rotational constant matrices $\Gamma^X$ differ. For the excited states, the differences in the $o_1(K)$ levels are marginal, between 0 and 0.6 cm$^{-1}$ ($\leq 0.5\%$), smaller than the deviation from experimental results when available. The corresponding splittings $i_i(K)$ show changes of about 4%, with the exception of the very small acceptor wag $i_2(0)$ value. For the ground state, the changes are about 1% for the energy levels $o_1(K)$. The associated splittings $i_i(K)$ are systematically decreased by 0.02 cm$^{-1}$ when moving from optimized to averaged rotational constants.

In Table III, we show the comparison between optimized $q_{\text{opt}}$ and vibrationally averaged $q_{\text{avg}}$ internal monomer coordinates at the water dimer equilibrium geometry $Q_{\text{eq}}$. Also given are the values $q_{\text{ref}}$ for the monomer reference geometry obtained by averaging over the isolated monomer ground state vibrational wave functions and used in the rigid CCpol-8s potential. Not surprisingly, the optimized values of $r_1$ and $r_2$ are systematically smaller than all the values obtained by vibrational averaging. Among the latter, we observe that the $r_1$ and $r_2$ values are larger when averaged over the adiabatic monomer wave functions in the dimer than when averaged over the isolated monomer wave functions. The lower part of the table lists the associated monomer rotational constants $\Gamma^X_{\alpha\alpha}(Q_{\text{eq}})$. With the exception of $\Gamma_{zz}$, the adiabatically averaged constants are closer to the reference values.

### 3. Flexible-monomer results

In Fig. 3, we report the energy levels, i.e., the origins $o_1(K)$ and $o_2(K)$ for $K=0$ and $K=1$, and the interchange tunneling splittings $i_1(K)$ and $i_2(K)$ calculated with the full flexible potential in the [6+6]D adiabatic approximation. Also the available experimental data are included in this figure.

When comparing this figure with Fig. 2, we can observe the effects of using flexible monomers, rather than rigid ones. The agreement with the experimental data, which was already very good for the rigid CCpol-8s potential, remains about equally good for the flexible-monomer potential CCpol-8sf. Only for the O–O stretch mode, the result slightly deteriorates by including monomer flexibility: the frequency obtained from CCpol-8sf is lower than the experimental value by about 4 cm$^{-1}$, while the frequency obtained from CCpol-8s agreed to 0.3 cm$^{-1}$. This might be related to the fact that, as shown in Table I, the SAPT-5s/IR hydrogen bond length ($O_\alpha$–$H_\alpha$) is too large by about 0.06 Å as compared to the benchmark calculations of Tschumper et al. Although the $O_\alpha$–$H_\alpha$ hydrogen bond length from the CCpol-8sf potential agrees much better with the corresponding result of Tschumper et al.,

<table>
<thead>
<tr>
<th>$K=0$</th>
<th>$K=1$</th>
</tr>
</thead>
<tbody>
<tr>
<td>Exp.</td>
<td>CCpol-8sf(avg)</td>
</tr>
<tr>
<td>OO</td>
<td>0.1536(1.88) / 149.63(1.23)</td>
</tr>
<tr>
<td></td>
<td>143.20(2.37)</td>
</tr>
<tr>
<td>DT$^2$</td>
<td>0.184.5(18.3)</td>
</tr>
<tr>
<td>AT</td>
<td>0.121.5(14.8)</td>
</tr>
<tr>
<td></td>
<td>0.120.19(9.39)</td>
</tr>
<tr>
<td>AW</td>
<td>0.108.8(0.02) / 0.107.82(1.0)</td>
</tr>
<tr>
<td></td>
<td>0.109.3(9.25) / 0.109.23(3.29)</td>
</tr>
<tr>
<td>DT</td>
<td>0.113.35(5.91)</td>
</tr>
<tr>
<td></td>
<td>0.64.52(2.54) / 0.63.32(2.48)</td>
</tr>
</tbody>
</table>

FIG. 3. Comparison of experimental VRT levels and tunneling splittings with results calculated with the CCpol-8sf potential using averaged values of the $\Gamma^X$ rotational matrices. For definitions, see Fig. 2.
It was found that the energies which was taken from SAPT-5s jA rigid-monomer CCpol-8s potential. These calculations were acceptor wag mode.

13, intermolecular basis in the [6 ≤ C]. Intermolecular VRT levels for (D2O)2 = 1108.2 cm−1 obtained from the CCpol-8s potential.

FIG. 4. Comparison of (D2O)2 experimental VRT levels and tunneling splittings with values obtained from the rigid CCpol-8s potential. For definitions, see Figs. 2 and 4. The symbol TW stands for a combination of the donor torsion and acceptor wag mode.

The inclusion of the correlation VAB(qA, qB; Q) between the intramolecular vibrations, Eq. (19), by perturbation theory, cf. Eq. (23), essentially left the ground state characteristics unchanged, while the energies of the excited states changed at most by a few times 0.1 cm−1 and the splittings by a few times 0.01 cm−1. Again, this effect can be understood from the increased localization of the intramolecular wave functions of D2O, relative to H2O.

Figure 5 presents a comparison of the VRT levels and splittings calculated on the flexible-monomer CCpol-8s potential with the experimental data.8,35,34,56–58 No significant differences are observed with respect to the rigid CCpol-8s results in Fig. 4, the good agreement with the available experimental data remains.
D. Frequency shifts of the monomer modes in (H₂O)₂

In Sec. III D we explained that, in contrast to the ground state, the adiabatic [6+6]D approximation is not completely valid if one of the monomers is vibrationally excited. We proposed two approximate adiabatic models, called \( G_{16} \) and \( G_8 \), that should give a good approximation to the more rigorous “two-state” non-adiabatic model that is appropriate for the vibrationally excited states. It may be expected that a comparison of the results of the \( G_{16} \) and \( G_8 \) models for the vibrational frequency shifts will give us a good indication of the accuracy of the shifts calculated by either model.

The frequency shift of each monomer vibration was obtained by subtracting the energy of the dimer ground state level computed on the lowest adiabatic 6D potential (with both monomers in their ground state) from the dimer ground state level computed on the excited adiabatic potential obtained for the corresponding monomer vibration. In the \( G_8 \) model, one can choose which monomer is excited, A or B; the results are the same. Since a given monomer excited state gives rise to two dimer excitations in the \( G_8 \) model, one on the donor and one on the acceptor, we had to consider the lowest two eigenvalues from each calculation. The acceptor can be identified \textit{a posteriori} as the monomer that has its O atom in the shortest hydrogen bond \( O \cdots H \). A comparison of the results from the \( G_{16} \) and \( G_8 \) models is given in Table IV. Furthermore, we show in this table what is the effect of simplifying the full adiabatic [6+6]D approach to the computationally cheaper [6+[3+3]]D approximation, as well as the effect of taking monomer rotational constant matrices \( \Gamma^{\text{RIK}} \) calculated at the optimized monomer geometries, instead of using the values averaged over the adiabatic monomer wave functions.

Before discussing the results, we must also explain the symmetry of the intramolecular vibrations and the appropriate selection rules. The water dimer in its equilibrium geometry has a mirror plane, point group \( C_\infty \). The donor monomer (D) lies in the plane of reflection and its symmetric and asymmetric OH stretch modes become localized. One of them is the O–H stretch mode of the H atom involved in the hydrogen bond, O–H₂, the other one is the O–H stretch mode of the free H atom, O–H₂. Both modes have \( A' \) symmetry in \( C_\infty \). The two H atoms of the acceptor monomer (A) stick out of the symmetry plane; they are interchanged by the reflection symmetry operator. The acceptor symmetric O–H stretch (ss) mode has \( A' \) and the asymmetric O–H stretch (as) mode has \( A'' \) symmetry. The HOH bend modes of both monomers have \( A' \) symmetry. Since the water dimer tunnels between eight equivalent equilibrium geometries of this type, we must use the molecular symmetry group \( G_{16} \), instead of the point group \( C_\infty \). The intramolecular vibrations are all of \( A' \) symmetry, except for the as[A] mode which is of \( A_2^+ \) symmetry. Actually, if the as[A] mode were adapted to the full \( G_{16} \) symmetry, one would obtain \( A_2^+ \) and \( B_2^+ \) symmetry components, but we already explained above that the adiabatic separation of the monomer vibrational coordinates and the intermolecular coordinates does not fully apply to the excited intramolecular vibrations. Hence, it is reasonable to apply the selection rules of the subgroup \( G_8 \), as it was done in the interpretation of the high-resolution spectra in Ref. 5. In \( G_8 \) symmetry, the \( G_{16} \) irreps \( A_1^+ \) and \( B_2^+ \) become equivalent. If one (or more) of the intermolecular (tunneling or vibrational) modes is excited, their symmetries should be combined with those of the intramolecular modes, as well as with those of the overall rotation functions of the dimer. The transition dipole moment operator has \( A_1^- \) symmetry, which tells us that transitions are allowed between irreps of the same type, except for the \( \pm \) labels which must be reversed. This is a strict selection rule. In addition, there are approximate selection rules based on the separability between the intra- and intermolecular vibrations. Furthermore, there may be approximate selection rules if the vibrations have small amplitudes and can be well separated from the overall rotations. This does not apply to the intermolecular vibrations, as they are in the free monomers. For all modes except as[A], we report in Table IV the shifts associated with the \( A_1^- \left[ J' = 1, K'' = 0 \right] \leftrightarrow A_2^+ \left[ J' = 0, K'' = 0 \right] \) transitions belonging to the lower acceptor tunneling components. These transitions may be considered as pure intramolecular vibrational fundamentals, accompanied only by a change \( \Delta J = 1 \) in total angular momentum that does not alter the internal motions in the dimer. For the as[A] mode, such transitions are forbidden; the excitation of this mode is only allowed in combination with a change in the projection quantum number \( K \). As one can see in Figs. 2 and 3, the VRT levels for \( K = 0 \) are quite different from the levels for \( K = 1 \), so it is clear that the change of \( K \) affects the internal motions of the monomers in the dimer. In Table IV, we report the allowed \( A_1^- \left[ J' = 1, K'' = 1 \right] \leftrightarrow A_2^+ \left[ J' = 0, K'' = 0 \right] \) transition for the as[A] mode, which was also measured. 5

Within the [6+[3+3]]D approximation (rightmost columns in the table), we first assessed the changes in the frequency shifts upon using averaged or optimized \( \Gamma \) matrices in the \( G_{16} \) model. It can be seen that the shifts are considerably modified, in some cases by as much as 5 cm⁻¹. The explanation of these relatively large differences is that these shifts result from transitions to excited adiabatic potentials which correspond to more delocalized intramolecular wave functions than the ground state. Therefore, the averaging can substantially modify the \( \Gamma \) matrices and we used it in all.

<table>
<thead>
<tr>
<th></th>
<th></th>
<th></th>
<th></th>
</tr>
</thead>
<tbody>
<tr>
<td>( b[D] )</td>
<td>+13.4</td>
<td>+13.7</td>
<td>+14.8</td>
</tr>
<tr>
<td>( O-H_2[D] )</td>
<td>−61.6</td>
<td>−57.9</td>
<td>−59.0</td>
</tr>
<tr>
<td>( O-H_2[A] )</td>
<td>−13.5</td>
<td>−17.6</td>
<td>−18.9</td>
</tr>
<tr>
<td>( \pm 7.3 )</td>
<td>+9.6</td>
<td>+7.5</td>
<td>+4.4</td>
</tr>
<tr>
<td>( ss[A] )</td>
<td>−7.1</td>
<td>−8.6</td>
<td>−8.7</td>
</tr>
<tr>
<td>( as[A] )</td>
<td>−6.6</td>
<td>−10.4</td>
<td>−9.0</td>
</tr>
</tbody>
</table>
TABLE V. Calculated frequency shifts of the monomer vibrations in (H2O)2 for J = 0 → 1 and J = 1 → 0 dimer transitions. The symmetries indicated refer to the intramolecular VRT levels that produce allowed transitions in combination with the intramolecular mode. For the K = 1 → 0 transitions in the last column, the symmetries of the initial and final states are reversed.

<table>
<thead>
<tr>
<th>Intramolecular mode</th>
<th>Transition</th>
<th>Frequency shift (cm⁻¹)</th>
</tr>
</thead>
<tbody>
<tr>
<td></td>
<td>Symmetry</td>
<td>K = 0 → 0</td>
</tr>
<tr>
<td>b[D]</td>
<td>A1      → A1</td>
<td>13.39</td>
</tr>
<tr>
<td></td>
<td>B1      → B1</td>
<td>11.84</td>
</tr>
<tr>
<td></td>
<td>A3      → A2</td>
<td>14.45</td>
</tr>
<tr>
<td></td>
<td>B3      → B2</td>
<td>13.13</td>
</tr>
<tr>
<td>O–H[0]D</td>
<td>A1      → A1</td>
<td>-6.63</td>
</tr>
<tr>
<td></td>
<td>B1      → B1</td>
<td>-62.83</td>
</tr>
<tr>
<td></td>
<td>A2      → A2</td>
<td>-63.15</td>
</tr>
<tr>
<td></td>
<td>B2      → B2</td>
<td>-64.17</td>
</tr>
<tr>
<td>b[A]</td>
<td>A1      → A1</td>
<td>7.26</td>
</tr>
<tr>
<td></td>
<td>B1      → B1</td>
<td>5.68</td>
</tr>
<tr>
<td></td>
<td>A2      → A2</td>
<td>7.55</td>
</tr>
<tr>
<td></td>
<td>B2      → B2</td>
<td>6.26</td>
</tr>
<tr>
<td>s[A]</td>
<td>A1      → A1</td>
<td>-7.06</td>
</tr>
<tr>
<td></td>
<td>B1      → B1</td>
<td>-8.59</td>
</tr>
<tr>
<td></td>
<td>A2      → A2</td>
<td>-7.49</td>
</tr>
<tr>
<td>a[A]</td>
<td>A1      → A2</td>
<td>3.48</td>
</tr>
<tr>
<td></td>
<td>B1      → B1</td>
<td>3.22</td>
</tr>
<tr>
<td></td>
<td>A2      → A2</td>
<td>-6.65</td>
</tr>
<tr>
<td></td>
<td>B2      → B1</td>
<td>-6.65</td>
</tr>
</tbody>
</table>

subsequent calculations. We then considered the two limiting models, G16 and G8. The results from these two models are quite similar, the largest difference being about 2 cm⁻¹. The smallest difference one might expect is about half of the size of the ground state interchange tunneling splitting 1 raised, cf. Fig. 3, since this splitting is given by the G16 calculations but not by the G8 model.

The first column in Table IV provides the calculated shifts obtained within the [6+6]D full variational approach, using averaged Γ matrices. Comparison with the second column allowed us to estimate the importance of the residual coupling term VAB(q₃₋₀,q₃₋₂;Q) between the two monomers, cf. Eq. (19). This coupling is particularly important for the donor O–H stretches and the acceptor asymmetric stretch modes, resulting in changes as large as 4 cm⁻¹. In all further calculations, we used the full [6+6]D adiabatic approach and the G16 model.

The frequency shifts of the allowed J = 0 → 1 and J = 1 → 0 transitions are listed in Table V for all six fundamental intramolecular modes. It is assumed that only different intermolecular tunneling components of the ground and excited states take part in these transitions, and that the intramolecular vibrations are not excited. It should be kept in mind that for (H₂O)₂ the A₁, B₁, A₂, and B₂ levels have nuclear spin statistical weights 1, 0, 3, and 6, respectively. In addition, there are transitions between the levels of E₂⁻ symmetry with weight 3 which are not shown; their shifts are intermedia-
TABLE VI. Measured frequency shifts of the monomer vibrations (in cm\(^{-1}\)) in (H\(_2\)O\(_2\)) and the corresponding calculated values from different potentials, see text. The calculations with the CCpol-8sf potential used the adiabatic G16 full [6+6]D model with averaged \(\Gamma\). The values given are averages of values from Table V, as specified in the last column. For definitions, see Table II.

<table>
<thead>
<tr>
<th>mode</th>
<th>Expt.</th>
<th>Calculated</th>
<th></th>
<th></th>
<th></th>
<th></th>
<th></th>
</tr>
</thead>
<tbody>
<tr>
<td></td>
<td></td>
<td>CCpol-8sf</td>
<td>SAPT-5s/JR(^a)</td>
<td>HBB(^b)</td>
<td>HBB(^b)</td>
<td>Specification(^d)</td>
<td></td>
</tr>
<tr>
<td>(b[D])</td>
<td>+20(^f)</td>
<td>+13.2</td>
<td>+5</td>
<td>+9.6</td>
<td>+16.3</td>
<td>(K = 0 \rightarrow 0)</td>
<td></td>
</tr>
<tr>
<td>O–H(_2)[D]</td>
<td>−56,(^i)−59.0(^g)</td>
<td>−62.9</td>
<td>−52</td>
<td>−42.5</td>
<td>−46.3</td>
<td>(K = 0 \rightarrow 0)</td>
<td></td>
</tr>
<tr>
<td>O–H(_2)[D]</td>
<td>−21,(^i)−25.3 and −26.3(^g)</td>
<td>−14.6</td>
<td>−15</td>
<td>−2.1</td>
<td>−23.5</td>
<td>(K = 0 \rightarrow 0)</td>
<td></td>
</tr>
<tr>
<td>(b[A])</td>
<td>+6(^g)</td>
<td>+6.7</td>
<td>+4</td>
<td>+2.8</td>
<td>+6.9</td>
<td>(K = 0 \rightarrow 0)</td>
<td></td>
</tr>
<tr>
<td>(s[A])</td>
<td>−2.2(^g) (−2 to +3(^i))</td>
<td>−8.0</td>
<td>−7</td>
<td>−0.4</td>
<td>−3.0</td>
<td>(K = 0 \rightarrow 0)</td>
<td></td>
</tr>
<tr>
<td>(a[A])</td>
<td>−17.4(^g)</td>
<td>−21.5</td>
<td>−5.5</td>
<td>−20.8</td>
<td>(K = 1 \rightarrow 0)</td>
<td></td>
<td></td>
</tr>
<tr>
<td></td>
<td>−3.4(^g)−3.2(^i)</td>
<td>−6.6</td>
<td>−12</td>
<td>+9.8</td>
<td>−5.9</td>
<td>(K = 0 \rightarrow 1)</td>
<td></td>
</tr>
<tr>
<td></td>
<td>+4.4(^g)</td>
<td>+3.4</td>
<td>+18.5</td>
<td>+3.1</td>
<td>(K = 0 \rightarrow 1)</td>
<td></td>
<td></td>
</tr>
</tbody>
</table>

\(^a\) Calculated in Ref. 25.
\(^b\) Calculated in Ref. 34.
\(^c\) Calculated in Ref. 64.
\(^d\) The \(K = 0 \rightarrow 0\) transition frequencies differ only slightly for different irreps, see Table V; we averaged over all symmetry components.
\(^e\) From cavity ringdown spectra (Ref. 9).
\(^f\) From size-selected cluster molecular beam spectra (Refs. 61 and 62).
\(^g\) From helium droplet spectra (Ref. 9).
\(^h\) Extrapolated value from solid noble gas matrix spectra (Refs. 21 and 63).
\(^i\) From molecular beam high-resolution spectra (Refs. 20 and 5).

\(1A_2^-/B_2^+ \rightarrow A_1^+/B_1^+\), \(2A_2^-/B_2^+ \rightarrow A_1^+/B_1^+\), \(3A_2^-/B_2^+ \rightarrow A_2^+/B_2^+\).

probably due to the \(A_2^-\), \(B_2^+\) and \(A_1^+, B_1^+\) tunneling components of the \(K = 0 \rightarrow 1\) transition.

The frequency of the acceptor symmetric stretch mode was determined in matrix spectra.\(^{21,63}\) Extrapolation of the spectra from different matrices yields a frequency between 3655 and 3660 cm\(^{-1}\), i.e., a shift between −2 and 3 cm\(^{-1}\). In the helium nanodroplet spectra,\(^9\) this band was observed at 3654.4 cm\(^{-1}\), which amounts to a shift of −2.2 cm\(^{-1}\).

The donor bound and free OH stretch modes were observed in lower resolution by Huisken et al.\(^{61,62}\) A narrow peak at 3601 cm\(^{-1}\), shifted by −56 cm\(^{-1}\) with respect to the monomer symmetric stretch mode, was assigned to the bound OH stretch. Apart from the reassignment, this agrees well with the (incomplete) high-resolution data for this band in Ref. 5. Absorption intensity at 3735 cm\(^{-1}\), obtained by subtraction of two spectra, was ascribed to the donor free OH stretch; this amounts to a shift of −21 cm\(^{-1}\) with respect to the monomer asymmetric stretch mode. The He droplet spectra of the water dimer\(^9\) give a shift of −59.0 cm\(^{-1}\) for the donor bound OH stretch mode and two, not completely separated peaks with shifts of −25.3 and −26.3 cm\(^{-1}\) for the donor free OH stretch.

Some additional remarks can be made regarding the He nanodroplet spectra,\(^9\) which show bands for all four OH stretch modes. Only for the acceptor asymmetric stretch mode a large splitting (7.8 cm\(^{-1}\)) was measured, while for the other modes only a small splitting (1.0 cm\(^{-1}\) for the donor free OH stretch) or no splitting was observed. A similar splitting of the acceptor asymmetric stretch band was measured and explained by interchange tunneling\(^58\) for (D\(_2\)O\(_2\)). It should be kept in mind that the acceptor asymmetric stretch mode is the only one for which \(K = 0 \rightarrow 0\) transitions are forbidden. As explained above, we believe that the peaks in the He droplet spectra assigned to the acceptor asymmetric stretch mode belong to a \(K = 0 \rightarrow 1\) band; our calculations predict an acceptor tunneling splitting of the correct size for this band. As also explained above, we predict much smaller tunneling splittings for \(K = 0 \rightarrow 0\) transitions. This implies that the absorption intensities of all bands in the He droplet spectra, except for the acceptor asymmetric stretch band, are probably dominated by \(K = 0 \rightarrow 0\) transitions.

The donor and acceptor bend modes were observed by Paul et al. in cavity ringdown spectra.\(^60\) A peak at 1600.6 cm\(^{-1}\), i.e., a blueshift of 6 cm\(^{-1}\), was assigned to the acceptor bend, probably involving a parallel \(\Delta K = 0\) transition. From three peaks at 1613.8, 1614.7, and 1628.6 cm\(^{-1}\) in the spectrum,\(^60\) it was estimated that the donor bend frequency is 1615–1620 cm\(^{-1}\), a blueshift of more than 20 cm\(^{-1}\) with respect to the monomer bend. According to Ref. 60, these bands are dominated by perpendicular \(\Delta K = \pm 1\) transitions. Our shifts calculated for \(K = 0 \rightarrow 0\) transitions in the donor bend mode are about +13 cm\(^{-1}\), for the \(K = 0 \rightarrow 1\) and \(K = 1 \rightarrow 0\) transitions they vary from −4 to +30 cm\(^{-1}\). It is not clear which components dominate the intensities of these bands, but we included the value calculated for the \(K = 0 \rightarrow 0\) transition in Table VI.

For all transitions, except for the acceptor asymmetric stretch mode, we included in Table VI the averages of the calculated transition frequencies listed in Table V for the \(K = 0 \rightarrow 0\) transitions. The results computed with the HBB and HBB2 potentials in Refs. 34 and 64 were treated similarly. Our reason for doing this is that the transition frequencies of the different symmetry components are quite similar for \(\Delta K = 0\) transitions, so we expect that in the (lower resolution) spectra the absorption bands are dominated by such transitions. Table VI shows that the IR frequency shifts calculated in the [6+6]D model with the CCpol-8sf potential agree quite well with the available experimental data, better than
TABLE VII. Frequency shifts of the intramolecular modes in \( \text{H}_2\text{O}_2 \) calculated in the harmonic approximation (in \( \text{cm}^{-1} \)). For experimental values, see Table VI.

<table>
<thead>
<tr>
<th>Mode</th>
<th>CCpol-8s/</th>
<th>SAPT-5s/IR(^a)</th>
<th>Tschumper et al.(^{30})</th>
</tr>
</thead>
<tbody>
<tr>
<td>( b[D] )</td>
<td>+18.8</td>
<td>+17.5</td>
<td>+26.7</td>
</tr>
<tr>
<td>( \text{O–H}_1[D] )</td>
<td>−48.3</td>
<td>−43.5</td>
<td>−82.7</td>
</tr>
<tr>
<td>( \text{O–H}_2[D] )</td>
<td>−26.8</td>
<td>−24.1</td>
<td>−29.0</td>
</tr>
<tr>
<td>( b[A] )</td>
<td>+4.2</td>
<td>+3.1</td>
<td>+1.8</td>
</tr>
<tr>
<td>( s[A] )</td>
<td>−3.5</td>
<td>−6.1</td>
<td>−5.7</td>
</tr>
<tr>
<td>( a[A] )</td>
<td>−9.3</td>
<td>−10.0</td>
<td>−9.7</td>
</tr>
</tbody>
</table>

\(^{30}\)The reason why these values differ from those in Ref. 25 is explained in the text.

\(^a\)Since the harmonic approximation cannot distinguish between tunneling levels, its prediction for the \( a[A] \) shift is usually compared with the experimental value of about 11 \( \text{cm}^{-1} \) from Table V in Ref. 5.

The results obtained with the SAPT-5s/IR potential\(^{25}\) and the HBB potential,\(^{34}\) although the improvement over the SAPT-5s/IR results is small except for the \( b[D] \) mode. This holds also for the acceptor asymmetric stretch mode, where experimental data are available for the individual symmetry components, both for \( K = 0 \rightarrow 1 \) and \( K = 1 \rightarrow 0 \) transitions. The agreement with experiment achieved by the HBB2 potential is generally comparable to that of CCpol-8s/IR. The former shifts are somewhat better for the donor bend, free O–H stretch, and acceptor symmetric stretch modes, but the latter potential provides a much better agreement with experiment for the most significantly shifted donor bond O–H stretch mode.

As discussed earlier, the harmonic approximation works relatively well for the IR frequency shifts of the intramolecular modes, so we also computed harmonic frequencies for both the CCpol-8s/IR and SAPT-5s/IR potentials. The force constants are the second derivatives of the potential with respect to the same intra- and intermolecular coordinates as used in the calculations of the VRT levels, computed at the dimer equilibrium geometry. The frequencies calculated with the SAPT-5s/IR potential differ slightly from those in Ref. 25, since we discovered that the original SAPT-5s/IR potential had an error in one of the expansions in monomer coordinates that violated the permutational symmetry. We restored the symmetry in the potential, but since the optimizations of the potential parameters were performed with the error, the parameters were not completely optimal. Very recently, we repeated the optimization with the correct symmetry\(^{65}\) and the resulting potential is very close to the original one. The effect of this symmetrization on the spectra is very small: most of the IR shifts change by less than 1 \( \text{cm}^{-1} \), the largest change is 3 \( \text{cm}^{-1} \) for the \( \text{O–H}_2[D] \) mode. The results for the intramolecular modes, listed in Table VII, can be compared with the best harmonic frequencies in the literature computed directly at CCSD(T) level in an augmented basis set of between double- and triple-zeta quality by Tschumper et al.\(^{30}\) Note that in Ref. 30 the energies of the characteristic points are highly converged benchmark values, but the harmonic frequencies can probably be appreciably improved by using larger basis sets. Table VII shows that the shifts from the two potentials are similar and that the agreement with the results in Ref. 30 is fairly good, except for the large frequency shift of the \( \text{O–H}_2[D] \) mode. Also the agreement of the results in Ref. 30 with the measured frequency shifts is worst for this mode. For all other modes, the harmonic shifts, both from Ref. 30 and computed by us, agree quite well with the experimental values. For the intermolecular modes and tunneling splittings the harmonic model is completely inappropriate.

V. DISCUSSION

CCpol-8s/IR, the new flexible-monomer \( \textit{ab initio} \) water dimer potential presented here, constitutes a definite improvement over the previous SAPT-5s/IR potential.\(^{25}\) In particular, it has the depth at the minimum \( D_0 = 21.0 \text{kJ mol}^{-1} \) identical to the number of digits given with an accurate benchmark value. As the one-body \( \mathcal{V}^{(1)} \) term is taken as the spectroscopically accurate PFT2 water potential of Polansky and co-workers,\(^{36}\) it enforces the correct asymptotic monomer properties.

Calculation of the spectroscopic properties of a non-rigid water dimer constitutes a formidable task, as it formally corresponds to a 12-dimensional problem. To describe the vibrating water monomers, we started from the general cluster formulation of Gatti and Iung\(^{37}\) which expresses the KEO in a rigorous, compact form and offers the user a choice of internal coordinates. Due to the light-heavy-light nature of the \( \text{H}_2\text{O} \) monomers, Radau vectors constitute an optimal choice and lead to a compact expression of their KEO.

The large frequency separation between inter- and intramolecular modes allowed us to resort to an adiabatic decoupling, and to recast the flexible-monomer calculations of the dimer VRT states into pseudo-rigid ones performed on adiabatic potentials. These potentials are obtained by solving the six-dimensional intramolecular subsystem eigenvalue problem at fixed intermolecular geometries, and assigning the resulting energies of the vibrational manifold. The pseudo-spectral method used to perform pseudo-rigid dimer VRT calculations renders this adiabatic implementation particularly efficient, as the potential is expressed on the six-dimensional intermolecular grid.

Calculations relying on the exact solution of the intramolecular subsystem problem, termed \( [6+6][D], \) required a large amount of central processing unit time (up to 40 days on a 12-core computer): in fact, one has to solve a six-dimensional vibrational problem at each point of a six-dimensional grid. Therefore, we also investigated the \( [6+[3+3]]D \) approximation which solves the vibrational manifold of each monomer with the other one fixed at its local equilibrium geometry. Even for the accurate CCpol-8s/IR potential used here, the errors resulting from such an approximation represent only a fraction of the deviation from the experimental data. They are, however, more pronounced when considering the frequency shifts of the monomer vibrations observed in IR spectra. This might have been expected since the potential coupling between excited monomers is larger than for ground state monomers.

The intermolecular VRT levels of both \( \text{H}_2\text{O}_2 \) and \( \text{D}_2\text{O}_2 \) calculated with the rigid-monomer CCpol-8s potential agree extremely well with the measured high-resolution spectra.\(^{6,18,53}\) The present calculations show that the flexible-monomer correction which we introduced in the CCpol-8s/IR
potential did not further improve the agreement with experiment for the intramolecular VRT levels; it remains about equally good. Essentially the same conclusion was reached in calculations with the flexible-monomer HBB potential of Bowman and co-workers, when the VRT levels were computed from this potential in [6+6]D calculations with flexible monomers, some of the levels agree slightly better with the experimental data than in 6D calculations with rigid monomers, for other levels the agreement became slightly worse. Only the ground state acceptor tunneling splitting was systematically improved for the HBB potential when the monomers were kept flexible. The present results computed with the CCpol-8s potential may also be compared with similar results obtained recently with the newer HBB2 version of the HBB potential. The flexible-monomer HBB2 potential produces slightly better agreement with experiment for the acceptor tunneling splitting in the ground state of the dimer, but for the excited states and for the ground state interchange tunneling splittings the CCpol-8s results are slightly superior. Disregarding these subtle differences, it may be stated that the potentials discussed above produced VRT levels that agree very well with the MW and far IR data. Earlier calculations of the VRT levels of (H2O)2 and (D2O)2 with the various available water potentials have shown, however, that this is certainly not the case for all water potentials. In fact, in addition to the SAPT, CCpol, and HBB potentials, only the empirical VRT(ASP) potentials fitted to dimer spectra in same league in terms of agreement with experiment. It may be reiterated that the comparison of accurately calculated VRT levels of the water dimer with high-resolution experimental data constitutes a very critical test of the quality of the intermolecular potential, a test that is failed by most of the available water potentials.

Table VI shows that the frequency shifts of the monomer vibrations obtained with the CCpol-8s potential constitute a definite improvement when compared to the results in Ref. computed with the HBB potential. They are also slightly better than the shifts obtained from the SAPT-5s/IR potential. The frequency shifts obtained in Ref. with the HBB2 potential agree equally well with experiment as the shifts obtained here with the CCpol-8s potential. The shifts from the HBB2 potential are somewhat better for the donor bend, free O–H stretch, and acceptor symmetric stretch modes, the CCpol-8s potential produces a much better shift for the donor bound O–H stretch mode.

Three possible explanations can be given for the remaining discrepancies between theory and experiment:

1. Most of the intramolecular modes (an exception is the acceptor asymmetric stretch mode) were only observed in lower resolution molecular beam spectra with rather broad bands, in which the fine structure due to rotations and tunneling was not resolved, or in matrix spectra. The experimental frequencies extracted from these spectra mostly correspond to some maximum in the absorption bands, sometimes after subtraction of other bands. Therefore, it is uncertain what these frequencies actually correspond to, and to which of the calculated transition frequencies they should be compared.

2. The accuracy of the [6+6]D adiabatic model used for the water dimer with flexible monomers has not yet been evaluated. This is particularly relevant for the excited intramolecular vibrations since, as explained above, a formally correct treatment of the monomer excited states would require a “two-state” non-adiabatic model. The shifts computed by the more approximate adiabatic models, called G16 and G8, may be amenable to improvement.

3. One can estimate that residual inaccuracies of potential surfaces due to post-CCSD(T) excitations, basis set truncations, and fit imperfections are of the order of 1% of interaction energy, i.e., about 20 cm−1 in the minimum region. This value is larger than most discrepancies with experimental spectra, mainly due to the fact that most spectral values correspond to energy differences which are usually more accurate due to cancellations of errors. Recently, calculations beyond the CCSD(T) level improved agreement with experiment for the spectra of H2–CO by nearly an order of magnitude and similar improvements are probably possible for the water dimer.

Summarizing, we may conclude that, as reported previously, the CCpol-8s potential still constitutes the best available rigid-monomer potential for the description of the VRT levels of both (H2O)2 and (D2O)2. The VRT levels from the flexible-monomer CCpol-8s potential are, on average, equally good. Also the frequency shifts of the monomer vibrations calculated with the CCpol-8s potential agree well with the experimental data extracted from IR spectra (with some uncertainties, as indicated).

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